

The Mineral Susceptibility Database

Hey Number	Mineral Name	Other Names	Chemical Formula	Moh's Hardness	Tenacity	H&S Data	Notes
1	Elements & Alloys						
1.1	copper		Cu	2.5-3	malleable		
1.2	silver		Ag	2.5-3	malleable		
1.24	diamond		C	10	brittle		
1.29	tin		Sn	1.5-2	malleable		
1.30	lead		Pb	1.5	malleable		easily scratched
1.33	arsenic		As	3.5	brittle		
1.37	domeykite		Cu ₃ As	3-3.5			
1.38	algodonite		(Cu _{1-x} As _x); x≈0.15	4			
1.41	antimony		Sb	3-3.5	very brittle		
1.42	stibarsen	allemontite	AsSb	3-4			
1.48	dyscrasite	stibiotriargentite	Ag ₃ Sb	3.5-4	sectile, but brittle		
1.49	bismuth		Bi	2-2.5	sectile		
1.51	sulfur	α-sulfur; brimstone	S ₈	1.5-2.5	brittle		"negatively charged when rubbed" (Schumann 1992: 38)
1.54	tellurium		Te	2-2.5	brittle		
1.57	iron		Fe	4.5	malleable		
1.61	nickel		Ni	3.5			
1.62	kamacite		α-(Fe,Ni)	4			
1.63	taenite		γ-(Ni,Fe)	5-5.5			
1.82	platinum		Pt	4	malleable		Non-magnetic to distinctly magnetic when rich in iron
2	Carbides, Nitrides, Silicides and Phosphides						
2.3	osbornite		TiN	8.5	brittle		
2.14	schreibersite	rhabdite, shepardite	(Fe,Ni) ₃ P	6.5-7	very brittle		strongly magnetic
3	Sulfides, Selenides, Tellurides, Arsenides and Bismuthides						
3.1.1	chalcocite	cuprein	Cu ₂ S	2.5-3	brittle		
3.1.8	covellite		CuS	1.5-2	flexible		
3.1.9	berzelianite	selencuprite	Cu ₂ Se	2	somewhat malleable		
3.1.11	umangite		Cu ₃ Se ₂	3			
3.1.14	klockmannite		CuSe	2-2.5			
3.1.16	weissite		Cu _{2-x} Te	3			
3.1.17	rickardite		Cu ₇ Te ₅	3.5			
3.1.22	cubanite	barracanite	CuFe ₂ S ₃	3.5			
3.1.23	bornite	erubescite	Cu ₅ FeS ₄	3	brittle		
3.1.25	chalcopyrite	cupropyrrite	CuFeS ₂	3.5-4	brittle		

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3.2.1	acanthite		Ag ₂ S	2-2.5	sectile		
3.2.3	naumannite		Ag ₂ Se	2.5	sectile		
3.2.4	aguilarite		Ag ₄ SeS	2.5	sectile		
3.2.5	hessite		Ag ₂ Te	2-3	sectile		
3.2.8	cervelleite		Ag ₄ TeS	2			
3.2.10	stromeyerite		AgCuS	2.5-3	brittle		
3.2.11	mckinstryite		(Ag,Cu) ₂ S	1.5-2.5			
3.2.12	jalpaite		Ag ₃ CuS ₂	2-2.5			
3.2.13	eucairite		AgCuSe	2.5	brittle		
3.2.17	argentopyrite		AgFe ₂ S ₃	3.5-4	brittle		
3.3.1	calaverite		AuTe ₂	2.5-3	brittle		
3.3.2	krennerite	müllerine	Au ₃ AgTe ₈	2-3			
3.3.7	petzite		Ag ₃ AuTe ₂	2.5-3	brittle		
3.3.9	sylvanite		(Au,Ag) ₂ Te ₄	1.5-2	brittle		
3.4.2	oldhamite		(Ca,Mg)S	4			
3.4.4	sphalerite	blende	ZnS	3.5-4	brittle		soft
3.4.5	wurtzite	spiauterite	(Zn,Fe)S	3.5-4	brittle		
3.4.7	greenockite		CdS	3-3.5	brittle		
3.5.1	cinnabar	vermilion	HgS	2-2.5	sectile		
3.6.5	galena		PbS	2.5	brittle		
3.6.7	altaite		PbTe	2-3	sectile		
3.7.4	realgar	ruby sulfur	As ₄ S ₄	1.5-2	sectile		
3.7.8	orpiment		As ₂ S ₃	1.5-2	sectile		
3.7.14	stibnite		Sb ₂ S ₃	2	flexible		soft
3.7.17	bismuthinite		Bi ₂ S ₃	2-2.5	flexible		
3.7.29	tetradymite		Bi ₂ Te ₂ S	1.5-2			
3.7.30	joséite	oruetite	Bi ₄ TeS ₂	2	flexible		
3.8.4	daubréelite		Fe ²⁺ Cr ³⁺ ₂ S ₄	4.5-5	brittle		
3.8.6	molybdenite		MoS ₂	1-1.5	flexible		
3.8.12	alabandite		MnS	3.5-4	brittle		
3.8.13	hauerite		MnS ₂	4	brittle		
3.9.1	pyrrhotite	magnetic iron pyrites	Fe ₇ S ₈	3.5-4			strongly magnetic
3.9.3	pyrite	fool's gold	FeS ₂	6-6.5	brittle		sensitive to shock
3.9.4	marcasite		FeS ₂	6-6.5	brittle		
3.9.5	greigite	melnikovite	Fe ²⁺ Fe ³⁺ ₂ S ₄	4-4.5			
3.9.6	mackinawite	kansite	(Fe,Ni) ₉ S ₈	2.5			
3.9.7	smythite		(Fe,Ni) _{3+x} S ₄ ; x≈0-0.3	4.5	brittle		

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3.9.9	ferroselite	hastite	FeSe ₂	6-6.5	very brittle		
3.9.11	löllingite	Imohsine	FeAs ₂	5-5.5	brittle		
3.9.12	arsenopyrite		FeAsS	5.5-6	brittle		
3.9.NI	troilite		FeS	3.5-4.5			
3.10.2	linnaeite		Co ²⁺ Co ³⁺ ₂ S ₄	4.5-5.5			
3.10.10	safflorite		(Co,Ni,Fe)As ₂	4.5-5	brittle		
3.10.12	skutterudite	smaltite; modumite	CoAs ₃	5.5-6			
3.10.13	cobaltite		CoAsS	5.5	brittle		
3.10.17	glaucodot		(Co _{0.5} Fe _{0.5})AsS	5	brittle		
3.10.18	carrollite	sychnodymite	Cu(Co,Ni) ₂ S ₄	4.5-5.5	very brittle		
3.11.3	millerite		NiS	3-3.5	brittle		
3.11.4	polydymite	grünauite	Ni ²⁺ Ni ³⁺ ₂ S ₄	4.5-5.5			
3.11.5	vaesite		NiS ₂	4.5-5.5			
3.11.8	violarite	beyrichite	Fe ²⁺ Ni ³⁺ ₂ S ₄	4.5-5.5			
3.11.9	pentlandite	lillehammerite	(Fe,Ni) ₉ S ₈	3.5-4	brittle		
3.11.10	bravoite		(Fe,Ni)S ₂	6.5	brittle		
3.11.19	penrosite	blockite	(Ni,Co,Cu)Se ₂	2.5-3	brittle		
3.11.22	melonite		NiTe ₂	1-1.5			
3.11.26	maucherite	placodine, temiskamite	Ni ₁₁ As ₈	5			
3.11.27	nickeline	niccolite	NiAs	5-5.5	brittle		
3.11.28	rammelsbergite		NiAs ₂	5.5-6	brittle		
3.11.29	pararammelsbergite		NiAs ₂	5.5			
3.11.30	nickelskutterudite	chloanthite	(Ni,Co,Fe)As ₃	5.5-6			
3.11.34	gersdorffite		NiAsS	5.5	brittle		
3.11.36	breithauptite		NiSb	5.5	brittle		
3.11.38	ullmannite		NiSbS	5-5.5	brittle		
4	Oxysulfides						
4.2	kermesite		Sb ₂ S ₂ O	1-1.5			
5	Sulfosalts: Sulfarsenites & Sulfobismuthites						
5.1.4	chalcostibite	guejarite; rosite; wolfsbergite	CuSbS ₂	3-4	brittle		
5.1.7	emphletite	tannenite	CuBiS ₂	2	brittle		
5.1.9	wittichenite		Cu ₃ BiS ₃	2-3	brittle		
5.1.10	hodrušite		Cu ₈ Bi ₁₂ S ₂₂	4-5	very brittle		
5.2.1	smithite		AgAsS ₂	1.5-2	brittle		
5.2.2	trechmannite		AgAsS ₂	1.5-2			

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5.2.3	proustite	light red silver ore	Ag ₃ AsS ₃	2-2.5	brittle		
5.2.4	xanthoconite		Ag ₃ AsS ₃	2-3	brittle		
5.2.5	pearceite	arsenpolybasite	[Ag ₉ CuS ₄][(Ag,Cu) ₆ (As,Sb) ₂ S ₇]	2.5-3			2 known polymorphs, both formerly known as arsenpolybasite; forms solid solution series with polybasite-Tac
5.2.8	miargyrite	hypargyrite; kenngottite	AgSbS ₂	2.5			
5.2.9	pyrargyrite	dark red silver ore; argyrythrose	Ag ₃ SbS ₃	2.5	brittle		
5.2.10	pyrostilpnite	fireblende; pyrochrotite; pyrichrolite	Ag ₃ SbS ₃	2			
5.2.11	stephanite	goldschmidtine; psaturose; röschwächs	Ag ₅ SbS ₄	2-2.5	brittle		
5.2.12	polybasite	antimonpearcite	[(Ag,Cu) ₆ (Sb,As) ₂ S ₇][Ag ₉ CuS ₄]	2.5-3	very brittle		3 known polymorphs, including polybasite-Tac (formerly known as antimonpearcite); forms solid solution series with pearceite
5.2.16	aramayoite		Ag ₃ Sb ₂ (Bi,Sb)S ₆	2.5			
5.2.17	matildite	argentobismuthutite; morocochite; peruvite; plenargyrite	AgBiS ₂	2.5	brittle		
5.2.19	pavonite		AgBi ₃ S ₅	2			
5.2.24	volynskite		AgBiTe ₂	2.5-3			
5.2.NI	fettelite	sanguinite	[Ag ₆ As ₂ S ₇][Ag ₁₀ HgAs ₂ S ₈]	3.5	brittle		
5.3.1	nagyágite		[Pb ₃ (Pb,Sb) ₃ S ₆](Au,Te) ₃	1-1.5	flexible		
5.5.1	lorándite		TlAsS ₂	2-2.5	elastic		
5.5.7	hutchinsonite		TlPbAs ₅ S ₉	1.5-2	brittle		
5.5.10	vrbaite	urbaite	Hg ₃ Tl ₄ As ₈ Sb ₂ S ₂₀	3.5	brittle		
5.6.1	sartorite	binnite; rathite-V	PbAs ₂ S ₄	3	very brittle		
5.6.2	dufrénoysite	plumbobinnite; rathite-1a; scleroclase	Pb ₂ As ₂ S ₅	3	brittle		
5.6.4	baumhauerite		Pb ₃ As ₄ S ₉	3			
5.6.6	jordanite	reniforite; reniformite	Pb ₁₄ (As,Sb) ₆ S ₂₃	3	brittle		
5.6.7	fülöppite		Pb ₃ Sb ₈ S ₁₅	2.5	brittle		
5.6.8	zinkenite	keeleyite	Pb ₉ Sb ₂₂ S ₄₂	3-3.5			
5.6.13	semseyite		Pb ₉ Sb ₈ S ₂₁	2.5	brittle		

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5.6.15	boulangerite	embrithite; mullanite; yenerite	Pb ₅ Sb ₄ S ₁₁	2.5-3	brittle		
5.6.23	galenobismutite	bismutoplacionite	PbBi ₂ S ₄	2.5-3.5	flexible		
5.7.2	bournonite	berthonite; endellione; rädlerz; tripelglanz; wölchite	PbCuSbS ₃	2.5-3	brittle		
5.7.3	meneghinite		Pb ₁₃ CuSb ₇ S ₂₄	2.5	brittle		
5.7.9	aikinite	belonite; nadelierz; patrinite	PbCuBiS ₃	2-2.5			
5.7.20	marrite		AgPbAsS ₃	3	brittle		
5.7.22	andorite	sundtite; webnerite	AgPbSb ₃ S ₆	3.5	brittle		
5.7.23	ramdohrite		Pb _{5.9} Fe _{0.1} Mn _{0.1} In _{0.1} Cd _{0.2} Ag _{2.8} Sb _{10.8} S ₂₄	2	brittle		
5.7.27	fizélyite		Ag ₅ Pb ₁₄ Sb ₂₁ S ₄₈	2	very brittle		
5.7.28	owyheeite		Ag ₃ Pb ₁₀ Sb ₁₁ S ₂₈	2.5	brittle		
5.7.29	freieslebenite	donacargyrite	AgPbSbS ₃	2.5			
5.7.30	diaphorite	brongniardite; ultrabasite	Ag ₃ Pb ₂ Sb ₃ S ₈	2.5-3	brittle		
5.7.39	benjaminite		Ag ₃ Bi ₇ S ₁₂	3.5			
5.7.40	lengenbachite		Ag ₄ Cu ₂ Pb ₁₈ As ₁₂ S ₃₉	1-2	malleable		
5.7.47	rathite	arsenomelane; wiltshireite	Ag ₂ Pb _{12-x} Tl _{x/2} As _{18+x/2} S ₄₀	3			
5.8.1	samsonite		Ag ₄ MnSb ₂ S ₆	2.5			
5.8.4	berthierite	anglarite; chazellite; martourite	FeSb ₂ S ₄	2-3	brittle		
5.8.5	jamesonite	comuccite; pfaffite	Pb ₄ FeSb ₆ S ₁₄	2.5	brittle		
6	Sulfosalts						
6.1.6	stannite	bolivianite; volfsonite	Cu ₂ FeSnS ₄	4			
6.1.18	franckeite	Illiceria	Fe ²⁺ (Pb,Sn ²⁺) ₆ Sn ⁴⁺ ₂ Sb ₂ S ₁₄	2.5-3	malleable		
6.2.1	germanite		Cu ₁₃ Fe ₂ Ge ₂ S ₁₆	4	brittle		
6.2.4	argyrodite	plusinglanz	Ag ₈ GeS ₆	2.5-3	brittle		
6.2.5	canfieldite		Ag ₆ SnS ₆	2.5	brittle		
6.3.1	enargite	garbyite; guayacanite	Cu ₃ AsS ₄	3	brittle		
6.4.1	sulvanite		Cu ₃ VS ₄	3.5			
6.5.1	corderoite		α-Hg ²⁺ ₃ S ₂ Cl ₂	3			
6.5.NI	kenhsuite		γ-Hg ²⁺ ₃ S ₂ Cl ₂	2-3			

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7	Oxides & Hydroxides						
7.3.1	cuprite	ruberite	Cu ₂ O	3.5-4	brittle		
7.3.2	tenorite	melaconite	CuO	3.5	brittle		
7.4.7	periclase		MgO	5.5	brittle		
7.4.8	brucite	hydrate of/native magnesia, shepardite, texalite	Mg(OH) ₂	2.5-3	sectile		
7.5.1	zincite	spartalite, sterlingite (of Alger)	ZnO	4	brittle		
7.5.12	montroydite		HgO				
7.6.1	corundum	spath adamantin	Al ₂ O ₃	9	brittle		
7.6.4	gibbsite	gibbsitogelinite; hydrargillite; zirlite	Al(OH) ₃	2.5-3	brittle		hygroscopic; porous
7.8.1	quartz		SiO ₂	7	brittle		var. chalcedony "always somewhat porous" (Schumann 1974)
7.8.5	cristobalite		SiO ₂	6-7	brittle		
7.8.8	opal		SiO ₂ · nH ₂ O	5.5-6.5	brittle		porous; water content varies with conditions, composition, and porosity, but can be up to 30% H ₂ O; sensitive to pressure & stress
7.9.2	rutile	crispite; dicksbergite; titanite	TiO ₂	6-6.5	brittle		
7.10.10	thorianite		ThO ₂	6.5-7	brittle	usually radioactive	
7.11.20	massicot	chrysin	PbO	2			
7.11.21	litharge		PbO	2			
7.11.22	minium	red lead	Pb ₃ O ₄	2.5			
7.11.23	plattnerite	diplasis plumbicus	PbO ₂	5.5	brittle		
7.12.8	navajoite		(V ⁵⁺ , Fe ³⁺) ₁₀ O ₂₄ · 12H ₂ O	1.5	sectile		
7.12.10	bariandite		Al _{0.6} (V ⁵⁺ , V ⁴⁺) ₈ O ₂₀ · 9H ₂ O				
7.13.1	arsenolite		As ₂ O ₃	1.5		very poisonous	
7.13.4	valentinite	exitèlite; white antimony	Sb ₂ O ₃	2.5-3	brittle		hygroscopic; porous
7.13.6	stibiconite	arsenostibite; hydromoméite; stibanite; volgerite	Sb ³⁺ Sb ⁵⁺ ₂ O ₆ (OH)	5.5-7			
7.16.1	uraninite	ulrichite; uranatemnite	UO ₂	5-6	brittle	highly radioactive	
7.16.2	schoepite	epiianthinite	(UO ₂) ₈ O ₂ (OH) ₁₂ · 12H ₂ O	2.5	brittle		

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7.16.3	paraschoepite		$\text{UO}_3 \cdot 2\text{H}_2\text{O}$				existence questionable; probably a mixture of metaschoepite, paulscherrerite, and ianthinite
7.16.4	metaschoepite		$(\text{UO}_2)_8\text{O}_2(\text{OH})_{12} \cdot 10\text{H}_2\text{O}$				
7.16.7	ianthinite		$\text{U}^{4+}(\text{UO}_2)_5\text{O}_7 \cdot 10\text{H}_2\text{O}$	2-3			
7.16.21	davidite-(La)		$\text{La}(\text{Y},\text{U})\text{Fe}_2(\text{Ti},\text{Fe},\text{Cr},\text{V})_{18}(\text{O},\text{OH},\text{F})_{38}$	6			
7.16.22	davidite-(Ce)		$\text{Ce}(\text{Y},\text{U})\text{Fe}_2(\text{Ti},\text{Fe},\text{Cr},\text{V})_{18}(\text{O},\text{OH},\text{F})_{38}$	6			
7.16.28	vandendriesscheite		$\text{PbU}_7\text{O}_{22} \cdot 12\text{H}_2\text{O}$	3		radioactive	
7.17.2	tellurite	tellurige säure; tellurocker	TeO_2	2			
7.18.1	manganosite		MnO	5.5	brittle		
7.18.3	pyrolusite	lapis manganensis; pseudomanganite; varvicite	Mn^{4+}O_2	2; 6-6.5	brittle		hygroscopic; porous
7.18.7	manganite	newkirkite; sphenomanganite	$\text{Mn}^{3+}\text{O}(\text{OH})$	4	brittle		
7.18.10	pyrochroite	eisenpyrochroite	$\text{Mn}(\text{OH})_2$	2.5-3	elastic		
7.20.4	hematite	hematitogelite; iron glance	Fe_2O_3	5-6	brittle		
7.20.5	goethite	allcharite; brown hematite/ironstone; chileit; ehrenwerthite; hydrohematite; mesabite; przibramite	$\alpha\text{-Fe}^{3+}\text{O}(\text{OH})$	5-5.5	brittle		
7.20.6	akaganeite		$(\text{Fe}^{3+},\text{Ni}^{2+})_8(\text{OH},\text{O})_{16}\text{Cl}_{1.25} \cdot n\text{H}_2\text{O}$				hygroscopic
8	Halides						
8.1.2	villiaumite		NaF	2-2.5	brittle	highly toxic; irritant; see MSDS for further information	
8.1.3	halite	common salt; martinsite; natrikalite; rock salt; saltspar	NaCl	2.5	brittle		hygroscopic 'when contaminated' (Schumann 1992)
8.1.4	hydrohalite		$\text{NaCl} \cdot 2\text{H}_2\text{O}$				
8.1.5	carobbiite		KF	2-2.5			
8.1.6	sylvite	hoevelite	KCl	1.5-2	brittle		hygroscopic 'when contaminated' (Schumann 1992)
8.1.7	salammoniac	chlorammonium	NH_4Cl	1-2	sectile		porous
8.2.1	nantokite		CuCl	2-2.5	sectile		

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8.2.2	tolbachite		CuCl ₂				hygroscopic
8.2.3	eriochalcite	antofagasite; eriodalcite; erythrochalcite	CuCl ₂ · 2H ₂ O	2.5			
8.2.4	atacamite	marcylite, muriate of copper, remolinite	Cu ₂ (OH) ₃ Cl	3-3.5	brittle		
8.2.6	botallackite		Cu ₂ Cl(OH) ₃				
8.2.8	calumetite		CaCu ₄ (OH) ₈ Cl ₂ · 3.5H ₂ O	2	brittle		
8.2.9	anthonyite	vondechenite	Cu(OH,Cl) ₂ · 3H ₂ O	2	sectile		
8.2.11	marshite		CuI	2.5	brittle		
8.2.12	mitscherlichite		K ₂ CuCl ₄ · 2H ₂ O	2.5			
8.3.1	chlorargyrite	argyroceratite; cerargyrite; iodembolite; kerate	AgCl	1.5-2.5	sectile		
8.3.2	bromargyrite	bromyrite; chlorobromite	AgBr	2.5	sectile		
8.3.3	bromian chlorargyrite	embolite; orthobromite	Ag(Cl, Br)	1.5-2			IMA discredited species; considered variety of chlorargyrite
8.3.4	iodargyrite	iodyrite	AgI	1.5-2			
8.3.5	miersite		(Ag, Cu)I	2.5-3	brittle		
8.4.3	bischofite		MgCl ₂ · 6H ₂ O	1-2			
8.4.5	carnallite		KMgCl ₃ · 6H ₂ O	2.5			highly hygroscopic
8.4.6	tachyhydrite		CaMg ₂ Cl ₆ · 12H ₂ O	2			
8.4.7	fluorite	bruiachite; cand; lithophosphorus suhlensis	CaF ₂	4	brittle		
8.4.8	sinjarite		CaCl ₂ · 2H ₂ O	1.5			hygroscopic
8.4.9	antarcticite		CaCl ₂ · 6H ₂ O	2-3	brittle		
8.4.10	chlorocalcite	baeumlerite	KCaCl ₃	2.5-3			
8.4.11	rorisite		CaFCl	2			hygroscopic
8.4.NI	chloromagnesite		MgCl ₂				A doubtful species; natural magnesium chloride is regarded as very unlikely.
8.5.2	calomel		[Hg ₂] ²⁺ Cl ₂	1.5-2	sectile		
8.5.3	terlinguaite		Hg ₂ OCl	2-3	brittle		
8.5.6	eglestonite		([Hg ¹⁺] ₂) ₃ OCl ₃ (OH)	2.5	brittle		
8.6.2	chloraluminite		AlCl ₃ · 6H ₂ O				
8.6.3	cadwaladerite	lesukite	AlCl(OH) ₂ · 4H ₂ O				
8.6.5	cryolite		Na ₂ NaAlF ₆	2.5	brittle		

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8.6.10	koeninite	justite	$\text{Na}_4\text{Mg}_9\text{Al}_4\text{Cl}_{12}(\text{OH})_{22}$	1.5	flexible		
8.7.1	fluocerite-(Ce)		$(\text{Ce}, \text{La})\text{F}_3$	4.5-5	brittle		
8.8.4	cotunnite		PbCl_2	1.5-2	sectile		
8.8.17	bidauxite		$\text{Pb}_2\text{AgCl}_3(\text{F}, \text{OH})_2$	3	brittle		
8.8.18	boleite	argentopercylite	$\text{KPb}_{26}\text{Ag}_9\text{Cu}_{24}(\text{OH})_{48}\text{Cl}_{62}$	3-3.5			
8.10.1	scacchite		MnCl_2				hygroscopic
8.10.3	chloromanganokalite		K_4MnCl_6	2.5	brittle		
8.11.1	molysite		FeCl_3				
8.11.2	hydromolysite		$\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$				
8.11.3	rokühnite		$\text{FeCl}_2 \cdot 2\text{H}_2\text{O}$		flexible		
8.11.4	douglasite		$\text{K}_2\text{Fe}^{2+}\text{Cl}_4 \cdot 2\text{H}_2\text{O}$				
8.11.5	erythrosiderite		$\text{K}_2\text{Fe}^{3+}\text{Cl}_5 \cdot \text{H}_2\text{O}$				
8.11.6	rinneite		$\text{K}_3\text{NaFe}^{2+}\text{Cl}_6$	3			
8.11.7	kremersite	kopeiskite	$(\text{NH}_4, \text{K})_2\text{Fe}^{3+}\text{Cl}_5 \cdot \text{H}_2\text{O}$				
8.11.9	nickelbischofite		$\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$	1.5			
8.11.10	lawrencite		$(\text{Fe}^{2+}, \text{Ni})\text{Cl}_2$				
8.12.4	hieratite		K_2SiF_6	2.5			
9	Borates						
9.1.1	sassolite	sal sedativum	H_3BO_3	1			hygroscopic; porous
9.1.6	kernite	rasorite	$\text{Na}_2[\text{B}_4\text{O}_6(\text{OH})_2] \cdot 3\text{H}_2\text{O}$	2.5	flexible		
9.1.8	tincalconite		$\text{Na}_2(\text{B}_4\text{O}_7) \cdot 5\text{H}_2\text{O}$	2			
9.1.9	borax	bauracia; tincal; tinkar	$\text{Na}_2(\text{B}_4\text{O}_5)(\text{OH})_4 \cdot 8\text{H}_2\text{O}$	2-2.5	brittle		porous
9.1.14	larderellite	borocalcite	$(\text{NH}_4)_2\text{B}_5\text{O}_7(\text{OH})_2 \cdot \text{H}_2\text{O}$				
9.2.16	kaliborite	heintzite; paternoite	$\text{KMg}_2\text{H}[\text{B}_6\text{O}_8(\text{OH})_5]_2(\text{H}_2\text{O})_4$	4-4.5			
9.2.17	iquiqueite		$\text{K}_3\text{Na}_4\text{Mg}(\text{CrO}_4)\text{B}_{24}\text{O}_{39}(\text{OH}) \cdot 12\text{H}_2\text{O}$	2	brittle		somewhat hygroscopic
9.3.12	inyoite		$\text{Ca}(\text{H}_4\text{B}_3\text{O}_7)(\text{OH}) \cdot 4\text{H}_2\text{O}$	2	brittle		
9.3.14	colemanite	borspar	$\text{Ca}[\text{B}_3\text{O}_4(\text{OH})_3] \cdot \text{H}_2\text{O}$	4.5	very brittle		
9.3.20	probertite	kramerite	$\text{NaCaB}_5\text{O}_7(\text{OH})_4 \cdot 3\text{H}_2\text{O}$	3.5	brittle		
9.3.21	ulexite	borocalcite; franklandite; hayesine; tv stone	$\text{NaCa}[\text{B}_5\text{O}_6(\text{OH})_6] \cdot 5\text{H}_2\text{O}$	2.5	brittle		
10	Borates with othe anions						
10.1.2	bandylite		$\text{Cu}[\text{B}(\text{OH})_4]\text{Cl}$	2.5	elastic		
10.1.5	boracite		$\text{Mg}_3(\text{B}_7\text{O}_{13})\text{Cl}$	7-7.5			
10.1.12	hydrochlorborite		$\text{Ca}_4\text{B}_8\text{O}_{15}\text{Cl}_2 \cdot 21\text{H}_2\text{O}$	2.5			

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10.1.13	ekaterinite		$\text{Ca}_2(\text{B}_4\text{O}_7)(\text{Cl},\text{OH})_2 \cdot 2\text{H}_2\text{O}$	1-1.5			hygroscopic
10.3.1	sulfoborite		$\text{Mg}_3[\text{B}(\text{OH})_4]_2(\text{SO}_4)(\text{OH},\text{F})_2$	4-4.5	brittle		
11	Carbonates						
11.1.2	natrite		Na_2CO_3	1-1.5			
11.1.3	nahcolite	thermokalite	NaHCO_3	2.5			
11.1.5	thermonatrite		$\text{Na}_2\text{CO}_3 \cdot \text{H}_2\text{O}$	1-1.5	sectile		
11.1.6	trona		$\text{Na}_3\text{H}(\text{CO}_3)_2 \cdot 2\text{H}_2\text{O}$	2.5			
11.1.7	natron		$\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$	1-1.5	brittle		
11.1.10	nyerereite	natrofairchildite	$\text{Na}_2\text{Ca}(\text{CO}_3)_2$				
11.1.13	pirssonite		$\text{Na}_2\text{Ca}(\text{CO}_3)_2 \cdot 2\text{H}_2\text{O}$	3-3.5	brittle		
11.1.14	gaylussite		$\text{Na}_2\text{Ca}(\text{CO}_3)_2 \cdot 5\text{H}_2\text{O}$	2.5	very brittle		
11.1.16	fairchildite		$\text{K}_2\text{Ca}(\text{CO}_3)_2$	2.5			
11.1.17	bütschliite		$\text{K}_2\text{Ca}(\text{CO}_3)_2$				
11.1.18	teschemacherite		$(\text{NH}_4)\text{HCO}_3$	1.5	brittle		
11.1.NI	gregoryite		$(\text{Na}_2,\text{K}_2,\text{Ca})\text{CO}_3$				
11.2.1	malachite		$\text{Cu}_2(\text{CO}_3)(\text{OH})_2$	3.5-4	brittle		
11.2.2	azurite	chessylite; lasur	$\text{Cu}_3(\text{CO}_3)_2(\text{OH})_2$	3.5-4	brittle		
11.3.1	magnesite	baldissérite; giobertite	MgCO_3	3.5-4.5	brittle		porous
11.3.3	nesquehonite		$\text{MgCO}_3 \cdot 3\text{H}_2\text{O}$	2.5			
11.3.4	lansfordite		$\text{MgCO}_3 \cdot 5\text{H}_2\text{O}$	2.5			
11.3.7	hydromagnesite		$\text{Mg}_5(\text{CO}_3)_4(\text{OH})_2 \cdot 4\text{H}_2\text{O}$	3.5	brittle		
11.4.1	calcite		CaCO_3	3	brittle		
11.4.2	aragonite	chimborazite; conchite; ktypéite	CaCO_3	3.5-4	brittle		
11.4.4	monohydrocalcite		$\text{CaCO}_3 \cdot \text{H}_2\text{O}$	2-3			
11.4.5	ikaite	(sub-/tri-/penta-) hydrocalcite	$\text{CaCO}_3 \cdot 6\text{H}_2\text{O}$				
11.4.6	dolomite	magnesiadolomite; pearl/rhomb spar; ridolphte	$\text{CaMg}(\text{CO}_3)_2$	3.5-4	brittle		
11.6.1	smithsonite	bonamite	ZnCO_3	4-4.5	brittle		
11.6.2	hydrozincite	earthy calamine; marionite; zinc bloom	$\text{Zn}_5(\text{CO}_3)_2(\text{OH})_6$	2-2.5	very brittle		hygroscopic; porous
11.6.5	aurichalcite	buratite, messingite, risséite	$(\text{Zn},\text{Cu})_5(\text{CO}_3)_2(\text{OH})_6$	1-2	fragile		
11.7.1	scarbroite		$\text{Al}_5(\text{CO}_3)(\text{OH})_{13} \cdot 5\text{H}_2\text{O}$				hygroscopic; porous

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11.7.2	hydroscarbrite		$\text{Al}_{14}(\text{CO}_3)_3(\text{OH})_{36} \cdot n\text{H}_2\text{O}$				
11.7.13	hydrodresserite		$\text{BaAl}_2(\text{CO}_3)_2(\text{OH})_4 \cdot 3\text{H}_2\text{O}$	3-4			
11.8.3	lanthanite-(Ce)		$(\text{Ce}, \text{La}, \text{Nd})_2(\text{CO}_3)_3 \cdot 8\text{H}_2\text{O}$	2.5			
11.9.1	cerussite	white lead (ore)	PbCO_3	3-3.5	very brittle		
11.9.2	hydrocerussite		$\text{Pb}_3(\text{CO}_3)_2(\text{OH})_2$	3.5	brittle		
11.9.3	dundasite		$\text{PbAl}_2(\text{CO}_3)_2(\text{OH})_4 \cdot \text{H}_2\text{O}$	2			
11.11.3	swartzite		$\text{MgCa}(\text{UO}_2)(\text{CO}_3)_3 \cdot 12\text{H}_2\text{O}$				
11.11.4	grimselite		$\text{K}_3\text{Na}(\text{UO}_2)(\text{CO}_3)_3 \cdot \text{H}_2\text{O}$				
11.11.5	bayleyite		$\text{Mg}_2(\text{UO}_2)(\text{CO}_3)_3 \cdot 18\text{H}_2\text{O}$				
11.11.7	zellerite		$\text{Ca}(\text{UO}_2)(\text{CO}_3)_2 \cdot 5\text{H}_2\text{O}$	2		radioactive	
11.11.8	wyartite		$\text{CaU}^{5+}(\text{UO}_2)_2(\text{CO}_3)_4(\text{OH}) \cdot 7\text{H}_2\text{O}$	3-4			
11.12.1	rhodochrosite	dialogite; villaurite	MnCO_3	3.5-4	brittle		sensitive to shock
11.13.1	siderite	chalybite; junckérite; spathic/spathose iron	FeCO_3	3.5-4.5	brittle		paramagnetic
11.14.4	hellyerite		$\text{NiCO}_3 \cdot 5.5\text{H}_2\text{O}$	2.5	brittle		
12	Carbonates with other anions: with halides or sulfate						
12.1.2	barentsite		$\text{Na}_7\text{Al}(\text{HCO}_3)_2(\text{CO}_3)_2\text{F}_4$	3	brittle		
12.1.23	phosgenite	corneous lead (ore), cromfordite, galenoceratite	$\text{Pb}_2\text{CO}_3\text{Cl}_2$	2-3	sectile		
12.2.1	burkeite		$\text{Na}_6(\text{CO}_3)(\text{SO}_4)_2$	3.5	brittle		
12.2.2	hanksite		$\text{Na}_{22}\text{K}(\text{SO}_4)_9(\text{CO}_3)_2\text{Cl}$	3-3.5	brittle		
12.2.11	leadhillite	maxite	$\text{Pb}_4(\text{CO}_3)_2(\text{SO}_4)(\text{OH})_2$	2.5-3	brittle		
12.2.18	schröckingerite	dakeite	$\text{NaCa}_3(\text{UO}_2)(\text{CO}_3)_3(\text{SO}_4)\text{F} \cdot 10\text{H}_2\text{O}$	2.5			
13	Nitrates						
13.1	nitratine	natronitrite; natronsalt peter	NaNO_3	1.5-2	sectile		hygroscopic
13.2	niter	salt peter	KNO_3	2	brittle		
13.3	gerhardtite		$\text{Cu}_2(\text{NO}_3)(\text{OH})_3$	2	flexible		
13.4	nitromagnesite		$\text{Mg}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$				hygroscopic
13.5	nitrocalcite		$\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$	1-2			
13.6	nitrobarite		$\text{Ba}(\text{NO}_3)_2$	3			
13.8	sveite		$\text{KAl}_7(\text{NO}_3)_4(\text{OH})_{16}\text{Cl}_2 \cdot 8\text{H}_2\text{O}$				hygroscopic
13.9	darapskite		$\text{Na}_3(\text{SO}_4)(\text{NO}_3) \cdot \text{H}_2\text{O}$	2.5			
13.12	hydrombobomkulite		$(\text{Ni}, \text{Cu})\text{Al}_4((\text{NO}_3)_2, \text{SO}_4)(\text{OH})_{12} \cdot 13-14\text{H}_2\text{O}$				

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13.NI	gwhibaite	nitrammite	$(\text{NH}_4, \text{K})\text{NO}_3$	2			
14	Silicates not containing aluminum						
14.2.1	diopside		$\text{CaSiO}_3 \cdot \text{H}_2\text{O}$	5	brittle		
14.2.5	chrysocolla		$\text{Cu}_{2-x}\text{Al}_x(\text{H}_{2-x}\text{Si}_2\text{O}_5)(\text{OH})_4 \cdot n\text{H}_2\text{O}$, $x < 1$	2.5-3.5	brittle		hygroscopic; porous
14.3.1	phenakite		Be_2SiO_4	7.5-8	brittle		
14.4.11	sepiolite		$\text{Mg}_4(\text{Si}_6\text{O}_{15})(\text{OH})_2 \cdot 6\text{H}_2\text{O}$	2	brittle		hygroscopic; porous
14.5.1	wollastonite	donegalite, gjellebäckite	CaSiO_3	4.5-5	brittle		
14.5.6	larnite		Ca_2SiO_4	6			
14.9.6	titanite	sphene	$\text{CaTi}(\text{SiO}_4)\text{O}$	5-5.5	brittle		
14.10.1	zircon		$\text{Zr}(\text{SiO}_4)$	7.5	brittle	U- and Th-bearing zircon is radioactive	
14.12.4	pabstite		$\text{Ba}(\text{Sn}, \text{Ti})(\text{Si}_3\text{O}_9)$	6			
14.13.8	barysilite		$\text{Pb}_8\text{Mn}^{2+}[\text{Si}_2\text{O}_7]_3$	3			
14.16.10	uranophane	lambertite	$\text{Ca}(\text{UO}_2)_2(\text{SiO}_3\text{OH})_2 \cdot 5\text{H}_2\text{O}$	2-3	brittle		
14.17.1	rhodonite		$\text{Mn}^{2+}\text{SiO}_3$	5.5-6.5			very sensitive to shock
14.17.8	bementite		$\text{Mn}_7\text{Si}_6\text{O}_{15}(\text{OH})_8$	6	flexible		
14.18.21	inesite		$\text{Ca}_2(\text{Mn}, \text{Fe})_7\text{Si}_{10}\text{O}_{28}(\text{OH})_2 \cdot 5\text{H}_2\text{O}$	5.5-6	brittle		
14.19.1	fayalite		$\text{Fe}^{2+}_2\text{SiO}_4$	7			
14.19.8	hisingerite	canbyite	$\text{Fe}^{3+}_2(\text{Si}_2\text{O}_5)(\text{OH})_4 \cdot 2\text{H}_2\text{O}$	2.5-3	brittle		
15	Silicates of aluminum						
15.6	dickite		$\text{Al}_2(\text{Si}_2\text{O}_5)(\text{OH})_4$	2-2.5	flexible		hygroscopic & porous
15.7	nacrite		$\text{Al}_2(\text{Si}_2\text{O}_5)(\text{OH})_4$	2-2.5			hygroscopic & porous
15.11	allophane	ilbaite; razoumofskyn	$(\text{Al}_2\text{O}_3)(\text{SiO}_2)_{1.3-2} \cdot 2.5-3\text{H}_2\text{O}$	3	brittle		
15.NI	halloysite-10Å	endellite; (hydrated /hydro-)halloysite	$\text{Al}_2\text{Si}_2\text{O}_5(\text{OH})_4 \cdot 2\text{H}_2\text{O}$				hygroscopic; porous; Halloysite-10Å is not a polytype of halloysite. It is also a different species from halloysite-7Å, which does not contain molecular water.
16	Silicates containing aluminum & other metals						
16.1.2	spodumene	triphane	$\text{LiAlSi}_2\text{O}_6$	6.5-7	brittle		
16.2.2	analcime	eudnophite	$\text{Na}(\text{AlSi}_2\text{O}_6) \cdot \text{H}_2\text{O}$	5-5.5	brittle		
16.2.5	natrolite	apoanalcite, laubanite	$\text{Na}_2\text{Al}_2\text{Si}_3\text{O}_{10} \cdot 2\text{H}_2\text{O}$	5-5.5	brittle		
16.2.7	paranatrolite		$\text{Na}_2\text{Al}_2\text{Si}_3\text{O}_{10} \cdot 3\text{H}_2\text{O}$	5-5.5			

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16.3.8	muscovite	amphilogite; oncosine; white mica	$KAl_2(AlSi_3O_{10})(OH)_2$	2.5	elastic		variety illite = hygroscopic & porous
16.4.1	nepheline	sommitite	$Na_3K(Al_4Si_4O_{16})$	5.5-6	brittle		
16.6.1	beryl		$Be_3Al_2Si_6O_{18}$	7.5-8	brittle		
16.7.6	palygorskite		$(Mg,Al)_2Si_4O_{10}(OH) \cdot 4H_2O$	2-2.5			hygroscopic; highly porous
16.7.7	saponite	mauritzite; soapstone	$Ca_{0.25}(Mg,Fe)_3((Si,Al)_4O_{10})(OH)_2 \cdot nH_2O$	1-2			hygroscopic; porous
16.7.9	montmorillonite	amagorsite	$(Na,Ca)_{0.33}(Al,Mg)_2(Si_4O_{10})(OH)_2 \cdot nH_2O$	1-2			hygroscopic; porous
16.8.6	phlogopite		$KMg_3(AlSi_3O_{10})(OH)_2$	2-3	flexible		
16.9.2	anorthite		$Ca(Al_2Si_2O_8)$	6-6.5	brittle		
16.9.23	laumontite	schneiderite	$CaAl_2Si_4O_{12} \cdot 4H_2O$	3.5-4	brittle		
16.11.NI	chabazite		$(Ca,K_2,Na_2)_2[Al_2Si_4O_{12}]_2 \cdot 12H_2O$	4-5	brittle		Chabazite-Ca is the most common member of the Chabazite series of the Zeolite Group. Other members include chabazite-K, chabazite-Mg, chabazite-Na, and chabazite-Sr.
16.16.5	carpholite		$Mn^{2+}Al_2(Si_2O_6)(OH)_4$	5-5.5			
16.18.1	nontonite	chloropal; hydroferripyrophyllite	$Na_{0.3}Fe_2((Si,Al)_4O_{10})(OH)_2 \cdot nH_2O$	1.5-2			hygroscopic; porous
16.19.22	vermiculite	brodrickite	$Mg_{0.7}(Mg,Fe,Al)_6(Si,Al)_8O_{20}(OH)_4 \cdot 8H_2O$	1.5-2			
16.23.3	vesuvianite	idocrase	$Ca_{19}Fe^{3+}Al_4(Al_6Mg_2)(\square_4)\square[Si_2O_7]_4[(SiO_4)_{10}]O(OH)_9$	6.5	brittle		
17	Silicates containing other anions						
17.2.1	topaz	pyrophyssalite	$Al_2SiO_4(F,OH)_2$	8	brittle		
17.2.9	zinnwaldite		$KFe^{2+}Al(Al_2Si_2O_{10})(OH)_2 - KLi_2Al(Si_4O_{10})(F,OH)_2$	2-3	elastic		discredited; series of trioctahedral micas on/close to the siderophyllite-polyolithionite series; -1M, -2M, and -3T polytypes are known
17.2.NI	trilithionite		$K(Li,Al)_3(AlSi_3O_{10})(F,OH)_2$	2.5-3.5	elastic		forms polyolithionite-trilithionite series; mediaries are considered <i>lepidolite</i> , an IMA discredited species
17.3.3	sodalite	glaucolite	$Na_4(Si_3Al_3)O_{12}Cl$	5.5-6	brittle		
17.4.2	spurrite	paraspurrite	$Ca_5(SiO_4)_2(CO_3)$	5			
17.5.3	searlesite		$Na(H_2BSi_2O_7)$	3.5	brittle		hygroscopic; porous
17.10.4	haüyne		$(Na,K)_3(Ca,Na)(Al_3Si_3O_{12})(SO_4,S,Cl)$	5.5-6	brittle		

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18	Niobates & Tantalates						
18.1.4	franconite		$\text{Na}(\text{Nb}_2\text{O}_5)(\text{OH}) \cdot 3\text{H}_2\text{O}$	4			
18.1.11	microlite	haddamite; metasimpsonite; neotantalite; niobantalpyrochlor; tantalyrochlor	$(\text{Ca},\text{Na})_2\text{Ta}_2(\text{O},\text{OH},\text{F})_7$	5-5.5	brittle		Microlite is now a group name for Ta-dominant minerals within the pyrochlore supergroup (nomenclature revised: 2010)
18.2.1	fergusonite-(Y)		YNbO_4	5.5-6.5	brittle		Risörite = titaniferous variety
18.2.7	formanite-(Y)		YTao_4	5.5-6.5			
18.2.8	pyrochlore	endeiolite; fluochlor; hydrochlor; niobpyrochlor	$\text{A}_2\text{Nb}_2(\text{O},\text{OH})_6\text{Z}$; where A = Na, Ca, Sn^{2+} , Sr, Pb^{2+} , Sb^{3+} , Y, U_4^{+} , H_2O or □; & Z = OH, F, O, H_2O or □				Pyrochlore is now a group name for a number of minerals within the pyrochlore supergroup (nomenclature recently revised ~2010-2013); analysis is necessary to identify a specimen's species
18.3.3	betafite	blomstrandite; hatchettolite	$(\text{Ca},\text{Na},\text{U})_2(\text{Ti}, \text{Nb},\text{Ta})_2\text{O}_6\text{Z}(\text{OH})$	3-5.5	brittle	strogly radioactive	IMA discredited species; many historically labelled specimens are no longer part of the betafite group due to replacement and or alteration; analysis is necessary to identify a specimen's species; Betafite is also now a group name for a number of minerals within the pyrochlore supergroup (nomenclature recently revised ~2010-2013)
18.4.1	uranpyrochlore	ellsworthite; U-rich oxynatropyrochlore	$(\text{Ca},\text{U},\text{Ce})_2(\text{Nb},\text{Ti},\text{Ta})_2\text{O}_6(\text{OH},\text{F})$	4-5			IMA discredited species (due to lack of acceptable analyses); considered a variety of the pyrochlore group
18.4.5	euxenite		$(\text{Y},\text{Ca},\text{Ce},\text{U},\text{Th})(\text{Nb},\text{Ta},\text{Ti})_2\text{O}_6$	5.5-6.5		radioactive	
18.4.8	polycrase-(Y)		$\text{Y}(\text{Ti},\text{Nb})_2(\text{O},\text{OH})_6$	5-6			
18.4.12	samaraskite-(Y)	adelpholite; ampangabéite; eytlandite; nuevite; uranoniobit; uranotantal	$\text{YFe}^{3+}\text{Nb}_2\text{O}_8$	5-6	brittle	radioactive	
19	Phosphates						
19.1.2	olympite		$\text{Na}_5\text{Li}(\text{PO}_4)_2$	4	brittle		
19.1.3	nahpoite		Na_2HPO_4	1-2	brittle		
19.1.4	dorfmanite		$\text{Na}_2\text{HPO}_4 \cdot 2\text{H}_2\text{O}$	1-1.5			
19.1.NI	schertelite	muellerite	$(\text{NH}_4)_2\text{MgH}_2(\text{PO}_4)_2 \cdot 4\text{H}_2\text{O}$				

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19.2.8	turquoise	calaite; chalchihuitl; chalchuite; hydrargillite; johnite	$\text{CuAl}_6(\text{PO}_4)_4(\text{OH})_8 \cdot 4\text{H}_2\text{O}$	5-6	brittle		hygroscopic; porous
19.3.18	phosphorösslerite		$\text{Mg}(\text{HPO}_4) \cdot 7\text{H}_2\text{O}$	2.5	brittle		
19.3.23	struvite	guanite	$(\text{NH}_4)\text{Mg}(\text{PO}_4) \cdot 6\text{H}_2\text{O}$	1.5-2	brittle		
19.4.1	monetite		$\text{Ca}(\text{HPO}_4)$	3.5	brittle		
19.4.3	brushite	epiglaubite; stoffertite	$\text{Ca}(\text{HPO}_4) \cdot 2\text{H}_2\text{O}$	2.5	brittle		
19.5.6	nabaphite		$\text{NaBaPO}_4 \cdot 9\text{H}_2\text{O}$	2			
19.5.8	nastrophite		$\text{Na}(\text{Sr},\text{Ba})\text{PO}_4 \cdot 9\text{H}_2\text{O}$	2	brittle		
19.7.5	variscite	lucinite; peganite; sphaerite; utahlite	$\text{AlPO}_4 \cdot 2\text{H}_2\text{O}$	3.5-4.5			
19.8.2	brazilianite		$\text{NaAl}_3(\text{PO}_4)_2(\text{OH})_4$	5.5	brittle		
19.9.3-5	monazite	cryptolite, urite	$\text{REE}(\text{PO}_4)$; REE = Ce, La, Nd, Sm, or Gd			Monazite usually contains small amounts of Th or sometimes U, both making it radioactive.	The name monazite, without a elemental suffix, was the accepted name for all material in this group, regardless of the REE-bearing components' relative abundance. The great majority of "monazite" is monazite-(Ce); the other members are all very rare.
19.10.10	sincosite		$\text{Ca}(\text{VO})_2(\text{PO}_4)_2 \cdot 5\text{H}_2\text{O}$	1-2	brittle		
19.11.10	torbernite	uranite; uran-mica	$\text{Cu}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 12\text{H}_2\text{O}$	2-2.5	brittle	highly radioactive	
19.11.14	meta-autunite		$\text{Ca}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 6-8\text{H}_2\text{O}$	2-2.5	brittle		
19.11.15	autunite		$\text{Ca}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 11\text{H}_2\text{O}$	2-2.5	sectile	radioactive	
19.11.22	uranocircite		$\text{Ba}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 10\text{H}_2\text{O}$	2-2.5		radioactive	
19.11.32	uranospathite		$(\text{Al},\square)(\text{UO}_2)_2(\text{PO}_4)_2\text{F} \cdot 20(\text{H}_2\text{O},\text{F})$	2-2.5		radioactive	
19.11.40	lehnerite		$\text{Mn}^{2+}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 8\text{H}_2\text{O}$	2-3			
19.11.41	bassetite		$\text{Fe}^{2+}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 10\text{H}_2\text{O}$	2.5	brittle	radioactive	
19.12.10	heterosite	ferripurpurite; heterozite; pseudotriplite	$(\text{Fe}^{3+}, \text{Mn}^{3+})\text{PO}_4$	4-4.5	brittle		
19.12.11	purpurite	manganipurpurite	$\text{Mn}^{3+}\text{PO}_4$	4-4.5	brittle		
19.12.16	rockbridgeite	kobokobite	$\text{Fe}^{2+}\text{Fe}^{3+}_4(\text{PO}_4)_3(\text{OH})_5$	3.5-4.5	brittle		
19.12.22	hureaulite	baldaufite; (pseudo)palaite	$(\text{Mn},\text{Fe})_5(\text{PO}_4)_2(\text{HPO}_4)_2 \cdot 4\text{H}_2\text{O}$	5	brittle		
19.12.35	ludlamite	lehnerite	$\text{Fe}^{2+}_3(\text{PO}_4)_2 \cdot 4\text{H}_2\text{O}$	3.5	brittle		diamagnetic
19.12.36	griftonite	repossite	$\text{Fe}^{2+}\text{Fe}^{2+}_2(\text{PO}_4)_2$	5	brittle		
19.12.40	messelite		$\text{Ca}_2\text{Fe}^{2+}(\text{PO}_4)_2 \cdot 2\text{H}_2\text{O}$	3.5	brittle		
19.13.2	strengite		$\text{FePO}_4 \cdot 2\text{H}_2\text{O}$	3.5-4	brittle		

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19.13.11	vivianite	anglarite; mullicite	$\text{Fe}^{2+}_3(\text{PO}_4)_2 \cdot 8\text{H}_2\text{O}$	1.5-2	sectile		
19.12.19	phosphoferrite	reddingite	$\text{Fe}^{2+}_3(\text{PO}_4)_2 \cdot 3\text{H}_2\text{O}$	3-3.5	brittle		
19.14.14	anapaite	tamanite	$\text{Ca}_2\text{Fe}^{2+}(\text{PO}_4)_2 \cdot 4\text{H}_2\text{O}$	3.5	flexible		
19.14.28	koninckite		$\text{Fe}^{3+}\text{PO}_4 \cdot 3\text{H}_2\text{O}$	3.5-4			
19.14.29	lazulite	berkeyite; klaprothite; voraulite	$\text{MgAl}_2(\text{PO}_4)_2(\text{OH})_2$	5.5-6	brittle		
20	Arsenates						
20.1.2	olivenite		$\text{Cu}_2(\text{AsO}_4)(\text{OH})$	3	brittle		
20.1.8	euchroite		$\text{Cu}_2(\text{AsO}_4)(\text{OH}) \cdot 3\text{H}_2\text{O}$	3.5-4	brittle		
20.1.13	liroconite		$\text{Cu}_2\text{Al}(\text{AsO}_4)(\text{OH})_4 \cdot 4\text{H}_2\text{O}$	2-2.5			
20.2.3	rösslerite		$\text{Mg}(\text{HAsO}_4) \cdot 7\text{H}_2\text{O}$	2-3	brittle		
20.2.12	ferrarisite		$\text{Ca}_5(\text{AsO}_4)_2(\text{HAsO}_4)_2 \cdot 9\text{H}_2\text{O}$		brittle		
20.2.14	phaunouxite		$\text{Ca}_3(\text{AsO}_4)_2 \cdot 11\text{H}_2\text{O}$				
20.7.1	trögerite		$(\text{H}_3\text{O})(\text{UO}_2)(\text{AsO}_4) \cdot 3\text{H}_2\text{O}$	2-3	brittle		
20.7.6	metaováčekite		$\text{Mg}(\text{UO}_2)_2(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O}$	2.5			
20.7.7	nováčekite-I		$\text{Mg}(\text{UO}_2)_2(\text{AsO}_4)_2 \cdot 12\text{H}_2\text{O}$	2.5			
20.7.9	metauranospinite		$\text{Ca}(\text{UO}_2)_2(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O}$	2-3	very brittle		
20.7.10	uranospinite		$\text{Ca}(\text{UO}_2)(\text{AsO}_4)_2 \cdot 10\text{H}_2\text{O}$	2-3	flexible	radioactive	
20.7.12	heinrichite		$\text{Ba}(\text{UO}_2)_2(\text{AsO}_4)_2 \cdot 10\text{H}_2\text{O}$	2.5		radioactive	
20.7.14	arsenuranospathite		$\text{Al}(\text{UO}_2)_2(\text{AsO}_4)_2\text{F} \cdot 20\text{H}_2\text{O}$	2			
20.9.2	scorodite	loaisite	$\text{Fe}^{3+}\text{AsO}_4 \cdot 2\text{H}_2\text{O}$	3.5-4			
20.9.6	symplesite		$\text{Fe}^{2+}_3(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O}$	2.5	flexible		
20.10.1	erythrite		$\text{Co}_3(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O}$	1.5-2.5	sectile		
20.10.9	annabergite	cordillerite; nickel bloom/green/ocher	$\text{Ni}_3(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O}$	1.5-2.5	sectile		
21	Vanadates						
21.2.1	metarossite		$\text{Ca}(\text{V}_2\text{O}_6) \cdot 2\text{H}_2\text{O}$	1-2	flexible		hygroscopic; porous
21.2.2	rossite		$\text{Ca}(\text{VO}_3)_2 \cdot 4\text{H}_2\text{O}$	2-3	brittle		
21.2.3	metahewettite		$\text{CaV}_6\text{O}_{16} \cdot 3\text{H}_2\text{O}$		friable		
21.2.4	hewettite		$\text{CaV}_6\text{O}_{16} \cdot 9\text{H}_2\text{O}$				
21.2.5	pintadoite		$\text{Ca}_2(\text{V}_2\text{O}_7) \cdot 9\text{H}_2\text{O}$				
21.2.7	fernandinite		$(\text{Ca},\text{K})(\text{V}^{5+},\text{V}^{4+},\text{Fe}^{2+})_6\text{O}_{20} \cdot 10\text{H}_2\text{O}$	2-3	waxy		
21.2.8	pascoite		$\text{Ca}_3(\text{V}_{10}\text{O}_{28}) \cdot 17\text{H}_2\text{O}$	2.5			
21.2.9	sherwoodite		$\text{Ca}_{4.5}(\text{AlV}^{4+}_2\text{V}^{5+}_{12}\text{O}_{40}) \cdot 28\text{H}_2\text{O}$	2			
2.2.10	melanovanadite		$\text{Ca}(\text{V}^{5+},\text{V}^{4+})_4\text{O}_{10} \cdot 5\text{H}_2\text{O}$	2.5	brittle		

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21.2.13	metadelrioite		$\text{CaSr}(\text{V}_2\text{O}_6)(\text{OH})_2$	2			
21.2.14	delrioite		$\text{CaSr}(\text{V}_2\text{O}_6)(\text{OH})_2 \cdot 3\text{H}_2\text{O}$	2			
21.3.3	metaschoderite		$\text{Al}_2(\text{PO}_4)(\text{VO}_4) \cdot 6\text{H}_2\text{O}$	2			
21.3.4	schoderite		$\text{Al}_2(\text{PO}_4)(\text{VO}_4) \cdot 8\text{H}_2\text{O}$				
21.3.10	descloizite	eusynchite, ramirite	$\text{PbZn}(\text{VO}_4)(\text{OH})$	3-3.5	brittle		
21.4.4	carnotite		$\text{K}_2(\text{UO}_2)_2(\text{VO}_4)_2 \cdot 3\text{H}_2\text{O}$	2	fragile	radioactive	
21.4.7	metatyuyamunite		$\text{Ca}(\text{UO}_2)_2(\text{VO}_4)_2 \cdot 3\text{-}5\text{H}_2\text{O}$	2	sectile		
21.4.8	tyuyamunite	calicocarnotite	$\text{Ca}(\text{UO}_2)_2(\text{VO}_4)_2 \cdot 5\text{-}8\text{H}_2\text{O}$	1.5-2	sectile	radioactive	
21.4.9	rauvite		$\text{Ca}(\text{UO}_2)_2(\text{V}_{10}\text{O}_{28}) \cdot 16\text{H}_2\text{O}$		sectile		
21.4.11	metavanuralite		$\text{Al}(\text{UO}_2)_2(\text{VO}_4)_2(\text{OH}) \cdot 8\text{H}_2\text{O}$		brittle		
21.4.12	vanuralite		$\text{Al}(\text{UO}_2)_2(\text{V}_2\text{O}_8)(\text{OH}) \cdot 8.5\text{-}11\text{H}_2\text{O}$	2			
22	Phosphates, Arsenates, & Vanadates with other anions						
22.1.1	amblygonite		$\text{LiAl}(\text{PO}_4)\text{F}$	5.5-6	brittle		
22.1.23	triplite		$\text{Mn}^{2+}_2(\text{PO}_4)\text{F}$	5-5.5			
22.2.9	pyromorphite		$\text{Pb}_5(\text{PO}_4)_3\text{Cl}$	3.5-4	brittle		
22.2.10	mimetite	petterdite	$\text{Pb}_5(\text{AsO}_4)_3\text{Cl}$	3.5-4	brittle		
22.2.14	vanadinite		$\text{Pb}_5(\text{VO}_4)_3\text{Cl}$	2.5-3	brittle		
22.2.NI	apatite	estramadurite; kietyöite; pyroguanite; sombrierite	$\text{Ca}_5(\text{PO}_4)_3(\text{Cl}/\text{F}/\text{OH})$	5	very brittle		<i>Apatite</i> is a field term for unidentified members of the apatite group. Most apatite is fluorapatite.
22.3.6	chalcophyllite	copper mica, erinite, tamarite	$\text{Cu}_{18}\text{Al}_2(\text{AsO}_4)_4(\text{SO}_4)_3(\text{OH})_{24} \cdot 36\text{H}_2\text{O}$	2			
22.3.21	sasaite		$(\text{Al}, \text{Fe}^{3+})_{14}(\text{PO}_4)_{11}(\text{SO}_4)(\text{OH})_7 \cdot 83\text{H}_2\text{O}$				
22.4.5	tyrolite	kupaphrite, trichalcite	$\text{Ca}_2\text{Cu}_9(\text{AsO}_4)_4(\text{CO}_3)(\text{OH})_8 \cdot 11\text{H}_2\text{O}$	1.5-2	sectile		
25	Sulfates						
25.1.1	metathénardite	pyrotechnite; thénardite	Na_2SO_4	2.5-3	brittle		
25.1.2	matteuccite		$\text{NaHSO}_4 \cdot \text{H}_2\text{O}$				
25.1.3	mirabilite	exanthalose; glauber salt; reussin	$\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$	1.5-2			
25.1.4	arcanite	taylorite	K_2SO_4	2			
25.1.5	mercallite		KHSO_4				
25.1.6	misenite		$\text{K}_8\text{H}_6(\text{SO}_4)_7$				
25.1.7	aphthitalite	glaserite	$(\text{K}, \text{Na})_3\text{Na}(\text{SO}_4)_2$	3	brittle		

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25.1.8	mascagnite		$(\text{NH}_4)_2\text{SO}_4$	2.5	sectile		
25.1.NI	lecontite		$(\text{NH}_4,\text{K})\text{NaSO}_4 \cdot 2\text{H}_2\text{O}$	2-2.5			
25.2.1	chalcocyanite	hydrocyanite	CuSO_4	3.5			hygroscopic
25.2.2	bonattite		$\text{CuSO}_4 \cdot 3\text{H}_2\text{O}$				
25.2.3	chalcantite	blue/copper vitriol	$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$	2.5	brittle		
25.2.4	boothite		$\text{CuSO}_4 \cdot 7\text{H}_2\text{O}$	2-2.5	brittle		
25.2.5	dolerophanite		$\text{Cu}_2(\text{SO}_4)\text{O}$	3			
25.2.7	brochantite	brogniartine (of Huot), kamarezite, königine, waringtonite	$\text{Cu}_4(\text{SO}_4)(\text{OH})_6$	3.5-4			
25.2.11	kröhnkite	salvadorite	$\text{Na}_2\text{Cu}(\text{SO}_4)_2 \cdot 2\text{H}_2\text{O}$	2.5-3			
25.2.12	natrochalcite		$\text{NaCu}_2(\text{SO}_4)_2(\text{OH}) \cdot 2\text{H}_2\text{O}$	4.5			
25.2.13	piypite		$\text{K}_4\text{Cu}_4\text{O}_2(\text{SO}_4)_4 \cdot (\text{Na},\text{Cu})\text{Cl}$	2.5			
25.2.14	fedotovite		$\text{K}_2\text{Cu}_3(\text{SO}_4)_3\text{O}$	2.5			
25.2.15	cyanochroite		$\text{K}_2\text{Cu}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$	2-2.5			
25.2.20	cyanotrichite	lettsonite, namaqualite	$\text{Cu}_4\text{Al}_2(\text{SO}_4)(\text{OH})_{12} \cdot 2\text{H}_2\text{O}$	1-3			
25.2.24	ransomite		$\text{CuFe}_2(\text{SO}_4)_4 \cdot 6\text{H}_2\text{O}$	2.5			
25.2.27	poitevinite		$(\text{Cu},\text{Fe})\text{SO}_4 \cdot \text{H}_2\text{O}$	3.5			
25.3.1	kieserite	martinsite	$\text{MgSO}_4 \cdot \text{H}_2\text{O}$	3.5	fragile		
25.3.2	sanderite		$\text{MgSO}_4 \cdot 2\text{H}_2\text{O}$				
25.3.3	starkeyite		$\text{MgSO}_4 \cdot 4\text{H}_2\text{O}$	2-3	brittle		
25.3.4	pentahydrate		$\text{MgSO}_4 \cdot 5\text{H}_2\text{O}$	2.5			
25.3.5	hexahydrate		$\text{MgSO}_4 \cdot 6\text{H}_2\text{O}$	2-2.5	brittle		
25.3.6	epsomite	seelandite	$\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$	2-2.5	fragile		
25.3.8	vanthoffite		$\text{Na}_6\text{Mg}(\text{SO}_4)_4$	3.5	fragile		
25.3.9	blödite	astrakhanite	$\text{Na}_2\text{Mg}(\text{SO}_4)_2 \cdot 4\text{H}_2\text{O}$	2.5-3	brittle		
25.3.10	konyaite		$\text{Na}_2\text{Mg}(\text{SO}_4)_2 \cdot 5\text{H}_2\text{O}$	2.5			
25.3.11	löweite		$\text{Na}_{12}\text{Mg}_7(\text{SO}_4)_{13} \cdot 15\text{H}_2\text{O}$	2.5-3			
25.3.13	langbeinite		$\text{K}_2\text{Mg}_2(\text{SO}_4)_3$	3.5-4	brittle		
25.3.14	leonite		$\text{K}_2\text{Mg}(\text{SO}_4)_2 \cdot 4\text{H}_2\text{O}$	2.5-3			
25.3.15	picromerite	schönite	$\text{K}_2\text{Mg}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$	2.5			
25.3.17	boussingaultite		$(\text{NH}_4)_2\text{Mg}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$	2			
25.3.18	pickeringite	picralluminite; sonomaite	$\text{MgAl}_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	1.5-2	brittle		
25.3.NI.1	alpersite		$(\text{Mg},\text{Cu})(\text{SO}_4) \cdot 7\text{H}_2\text{O}$	2	brittle		
25.3.NI.2	meridianite		$\text{MgSO}_4 \cdot 11\text{H}_2\text{O}$				

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25.4.1	anhydrite		CaSO ₄	3-3.5	brittle		moderately high mechanical sensitivity
25.4.3	gypsum	marmor fugax; montmartrite	CaSO ₄ · 2H ₂ O	2	flexible		
25.4.4	glauconite		Na ₂ Ca(SO ₄) ₂	2.5-3	brittle		
25.4.8	syngenite	kalusite	K ₂ Ca(SO ₄) ₂ · H ₂ O	2.5			
25.4.10	polyhalite		K ₂ Ca ₂ Mg(SO ₄) ₄ · 2H ₂ O	2.5-3.5			
25.4.11	koktaite	ammonium-syngenite	(NH ₄) ₂ Ca(SO ₄) ₂ · H ₂ O				
25.4.12	ye'elimite	kruzhevite	Ca ₄ Al ₆ (SO ₄)O ₁₂				
25.4.13	ettringite	woodfordite	Ca ₆ Al ₂ (SO ₄) ₃ (OH) ₁₂ · 26H ₂ O	2-2.5			
25.4.15	celestine	eschwegite; sicilianite	SrSO ₄	3-3.5	brittle		very fragile
25.4.17	baryte	baroselenite; calstronbarite; dréelite; (fetid) heavy spar	BaSO ₄	3	brittle		very fragile
25.5.2	goslarite	galiznite; white/zinc vitriol	ZnSO ₄ · 7H ₂ O	2-2.5	brittle		
25.5.6	boyleite		(Zn,Mg)SO ₄ · 4H ₂ O	2			
25.5.7	serpierite		Ca(Cu,Zn) ₄ (SO ₄) ₂ (OH) ₆ · 3H ₂ O	2	brittle		
25.5.11	gunningite		ZnSO ₄ · H ₂ O	2.5			
25.5.13	bianchite		(Zn,Fe)SO ₄ · 6H ₂ O	2.5			
25.5.15	zincmelanterite		(Zn,Cu,Fe)SO ₄ · 7H ₂ O	2			
25.6.3	mendozite	natrumalaun; soda-alum	NaAl(SO ₄) ₂ · 11H ₂ O	3			
25.6.5	aluminite	hallite; native argill	Al ₂ (SO ₄)(OH) ₄ · 7H ₂ O	1-2			fragile; hygroscopic; porous
25.6.7	alunogen	davite; keramohalite; saldanite; solfatarite	Al ₂ (SO ₄) ₃ · 17H ₂ O	1.5-2			
25.6.8	alunite	aluminilite; alumite; calafatite; ignatiewite; lœwigite	KAl ₃ (SO ₄) ₂ (OH) ₆	3.5-4	brittle		
25.6.10	hydrobasaluminite		Al ₄ (SO ₄)(OH) ₁₀ · 12-36H ₂ O				
25.6.11	zaherite		Al ₁₂ (SO ₄) ₅ (OH) ₂₆ · 20H ₂ O	3.5			
25.6.12	tamarugite		NaAl(SO ₄) ₂ · 6H ₂ O	3			
25.6.14	alum-(Na)		NaAl(SO ₄) ₂ · 12H ₂ O	3			
25.6.16	kalinite	potassalunite	KAl(SO ₄) ₂ · 11H ₂ O	2			
25.6.17	alum-(K)		KAl(SO ₄) ₂ · 12H ₂ O	2			
25.6.20	tschermigite	ammonalun; ammonia alum	(NH ₄)Al(SO ₄) ₂ · 12H ₂ O	1.5-2			
25.7.3	anglesite	lead vitriol	PbSO ₄	2.5-3	brittle		very fragile

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Hey Number	Mineral Name	Other Names	Chemical Formula	Moh's Hardness	Tenacity	H&S Data	Notes
25.7.6	linarite		$\text{PbCu}(\text{SO}_4)(\text{OH})_2$	2.5	brittle		
25.8.5	minasragrite		$(\text{V}^{4+})_2(\text{SO}_4) \cdot 5\text{H}_2\text{O}$	1-2			
25.8.7	redingtonite		$(\text{Fe}^{2+}, \text{Mg}, \text{Ni})(\text{Cr}, \text{Al})_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	2			
25.8.9	uranopilite	uranocher	$(\text{UO}_2)_6(\text{SO}_4)_2(\text{OH})_6 \cdot 12-14\text{H}_2\text{O}$				
25.9.1	szmikite		$\text{MnSO}_4 \cdot \text{H}_2\text{O}$	1.5			
25.9.2	jökokuite		$\text{MnSO}_4 \cdot 5\text{H}_2\text{O}$	2.5			
25.9.3	mallardite		$\text{MnSO}_4 \cdot 7\text{H}_2\text{O}$	2			
25.9.5	chvaleticeite		$(\text{Mn}, \text{Mg})\text{SO}_4 \cdot 6\text{H}_2\text{O}$	1.5			
25.9.7	apjohnite		$\text{Mn}^{2+}\text{Al}_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	1.5-2			
25.9.9	ilesite		$(\text{Mn}, \text{Zn}, \text{Fe})\text{SO}_4 \cdot 4\text{H}_2\text{O}$	2-3			
25.9.13	dietrichite		$(\text{Zn}, \text{Fe}^{2+}, \text{Mn}^{2+})\text{Al}_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	2	brittle		
25.10.1	szomolnokite	ferropallidite	$\text{FeSO}_4 \cdot \text{H}_2\text{O}$	2.5	brittle		
25.10.2	rozenite		$\text{FeSO}_4 \cdot 4\text{H}_2\text{O}$	2-3			
25.10.3	siderotil		$\text{FeSO}_4 \cdot 5\text{H}_2\text{O}$	2.5			only forms when > 1.2% Cu present
25.10.4	ferrohexahydrate		$\text{FeSO}_4 \cdot 6\text{H}_2\text{O}$	2			
25.10.5	melanterite	green/iron vitriol	$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$	2	brittle		
25.10.6	rhomboclase		$(\text{H}_5\text{O}_2)\text{Fe}^{3+}(\text{SO}_4)_2 \cdot 2\text{H}_2\text{O}$	2	elastic		
25.10.8	kornelite		$\text{Fe}_2(\text{SO}_4)_3 \cdot 7\text{H}_2\text{O}$				
25.10.9	coquimbite	blakeite	$\text{AlFe}^{3+}_3(\text{SO}_4)_6(\text{H}_2\text{O})_{12} \cdot 6\text{H}_2\text{O}$	2.5	brittle		
25.10.10	paracoquimbite		$\text{Fe}^{3+}_4(\text{SO}_4)_6(\text{H}_2\text{O})_{12} \cdot 6\text{H}_2\text{O}$	2.5			
25.10.11	quenstedtite		$\text{Fe}^{3+}_2(\text{SO}_4)_3 \cdot 11\text{H}_2\text{O}$	2.5	brittle		
25.10.14	hohmannite	castanite	$\text{Fe}^{3+}_2(\text{SO}_4)_2\text{O} \cdot 8\text{H}_2\text{O}$	3			
25.10.17	amarantite	castanite; paposite	$\text{Fe}^{3+}_2(\text{SO}_4)_2\text{O} \cdot 7\text{H}_2\text{O}$	2.5	brittle		
25.10.18	fibroferrite		$\text{Fe}^{3+}_3(\text{SO}_4)(\text{OH}) \cdot 5\text{H}_2\text{O}$	2-2.5			
25.10.20	römerite	louderbackite	$\text{Fe}^{2+}\text{Fe}^{3+}_2(\text{SO}_4)_4 \cdot 14\text{H}_2\text{O}$	3-3.5			
25.10.21	břilinite		$\text{Fe}^{2+}\text{Fe}^{3+}_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	2			
25.10.22	copiapite	janosite	$\text{Fe}^{2+}\text{Fe}^{3+}_4(\text{SO}_4)_6(\text{OH})_2 \cdot 20\text{H}_2\text{O}$	2.5-3	fragile		
25.11.2	sideronatrite	urusite	$\text{Na}_2\text{Fe}(\text{SO}_4)_2(\text{OH}) \cdot 3\text{H}_2\text{O}$	1.5-2	elastic		hygroscopic; porous
25.11.3	metasideronatrite		$\text{Na}_2\text{Fe}(\text{SO}_4)_2(\text{OH}) \cdot \text{H}_2\text{O}$	1.5-2			
25.11.4	ferrinatrite	gordaite; leucoglaucite	$\text{Na}_3\text{Fe}(\text{SO}_4)_3 \cdot 3\text{H}_2\text{O}$	2.5	brittle		
25.11.7	krausite		$\text{KFe}(\text{SO}_4)_2 \cdot \text{H}_2\text{O}$	2.5	brittle		
25.11.9	jarosite	moronolite	$\text{KFe}^{3+}_3(\text{SO}_4)_2(\text{OH})_6$	2.5-3.5	brittle		
25.11.10	voltaite	monsmedite; pettkoite	$\text{K}_2\text{Fe}^{2+}_5\text{Fe}^{3+}_3\text{Al}(\text{SO}_4)_{12} \cdot 18\text{H}_2\text{O}$	3-3.5			

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25.11.12	metavoltine	metavoltaite	$\text{Na}_6\text{K}_2\text{Fe}^{2+}\text{Fe}^{3+}_6(\text{SO}_4)_{12}\text{O}_2 \cdot 18\text{H}_2\text{O}$	2.5			
25.11.17	lonecreekite		$(\text{NH}_4)\text{Fe}^{3+}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$	2-3			
25.11.18	godovikovite		$(\text{NH}_4)\text{Al}(\text{SO}_4)_2$	2			
25.11.19	botryogen	kubeite; palacheite; quetenite; red iron vitriol	$\text{MgFe}^{3+}(\text{SO}_4)_2(\text{OH}) \cdot 7\text{H}_2\text{O}$	2-2.5	brittle		
25.11.26	halotrichite		$\text{Fe}^{2+}\text{Al}_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	1.5-2	brittle		
25.11.NI	mikasaite		$\text{Fe}^{3+}_2(\text{SO}_4)_3$	2			
25.12.1	bieberite	cobalt/red vitriol	$\text{CoSO}_4 \cdot 7\text{H}_2\text{O}$	2			
25.12.3	aplowite		$(\text{Co,Mn,Ni})\text{SO}_4 \cdot 4\text{H}_2\text{O}$	3			
25.12.4	moorhouseite		$(\text{Co,Ni,Mn})\text{SO}_4 \cdot 6\text{H}_2\text{O}$	2.5			
25.12.5	retgersite		$\text{NiSO}_4 \cdot 6\text{H}_2\text{O}$	2.5	brittle		
25.12.6	morenosite		$\text{NiSO}_4 \cdot 7\text{H}_2\text{O}$	2-2.5			
25.12.13	nicklehexahydrite		$(\text{Ni,Mg,Fe})\text{SO}_4 \cdot 6\text{H}_2\text{O}$	2			
25.12.NI	wupatkiite		$(\text{Co,Mg,Ni})\text{Al}_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	1.5-2			
26	Sulfates with Halide						
26.2	sulphohalite		$\text{Na}_6(\text{SO}_4)_2\text{FCl}$				
26.3	schairerite		$\text{Na}_{21}(\text{SO}_4)_7\text{ClF}_6$				
26.7	kamchatkite		$\text{KCu}_3(\text{SO}_4)_2\text{OCl}$	3.5			hygroscopic
26.9	aubertite		$\text{CuAl}(\text{SO}_4)_2\text{Cl} \cdot 14\text{H}_2\text{O}$	2-3			
26.14	kainite		$\text{KMg}(\text{SO}_4)\text{Cl} \cdot 3\text{H}_2\text{O}$	2.5-3			
26.15	mosesite		$(\text{Hg}_2\text{N})(\text{Cl}, \text{SO}_4, \text{MO}_4) \cdot \text{H}_2\text{O}$				
26.16	kleinite		$(\text{Hg}_2\text{N})(\text{Cl}, \text{SO}_4) \cdot n\text{H}_2\text{O}$				
26.23	pseudograndreefite		$\text{Pb}_6(\text{SO}_4)\text{F}_{10}$	2.5	brittle		
26.25	caracolite		$\text{Na}_3\text{Pb}_2(\text{SO}_4)_3\text{Cl}$	4.5			
27	Sulfites, Chromates, Molybdates, & Tungstates						
27.2.1	tarapacáite		$\text{K}_2(\text{CrO}_4)$				
27.2.5	crocoite	beresofite; lehmannite	PbCrO_4	2.5-3	brittle		very fragile
27.2.6	phoenicochroite	subsesquichromate of lead	$\text{Pb}_2(\text{CrO}_4)\text{O}$	2.5			
27.2.NI	edoylerite		$\text{Hg}^{2+}_3(\text{CrO}_4)_2\text{S}_2$		brittle		
27.3.3	wulfenite		$\text{Pb}(\text{MoO}_4)$	2.5-3	brittle		
27.3.14	ferrimolybdite	molybdenic acid/oche; molybdite	$\text{Fe}^{3+}_2(\text{Mo}^{6+}\text{O}_4)_3 \cdot 7\text{H}_2\text{O}$	1-2			
27.4.15	wolframite		$(\text{Fe}^{2+})\text{WO}_4 - (\text{Mn}^{2+})\text{WO}_4$	4-4.5	brittle		

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28	Selenites, Selenates, Tellurites, & Tellurates						
28.1.8	marthozite		$\text{Cu}^{2+}(\text{UO}_2)_3(\text{SeO}_3)_2\text{O}_2 \cdot 8\text{H}_2\text{O}$				
29	Iodates						
29.3	lautarite		$\text{Ca}(\text{IO}_3)_2$	3.5-4			
29.5	dietzeite		$\text{Ca}_2(\text{IO}_3)_2(\text{CrO}_4) \cdot \text{H}_2\text{O}$	3.5			
30	Thiocyanates						
30.2	kafehydrocyanite		$\text{K}_4[\text{Fe}(\text{CN})_6] \cdot \text{H}_2\text{O}$	2-2.5			existence questionable
31	Oxalates, Citrates, Mellitates, & Acetates						
31.1.1	oxammitte		$(\text{NH}_4)_2(\text{C}_2\text{O}_4) \cdot \text{H}_2\text{O}$				
31.1.3	wheatleyite		$\text{Na}_2\text{Cu}(\text{C}_2\text{O}_4)_2 \cdot 2\text{H}_2\text{O}$	1-2	very brittle		
31.1.5	whewellite	oxacalcite, oxalate of lime	$\text{Ca}(\text{C}_2\text{O}_4) \cdot \text{H}_2\text{O}$	2.5-3	brittle		
31.1.NI	natroxalate		$\text{Na}_2(\text{C}_2\text{O}_4)$	3			
31.3.1	mellite	honey stone	$\text{Al}_2[\text{C}_6(\text{COO})_6] \cdot 16\text{H}_2\text{O}$	2-2.5	sectile		
32	Crystalline Organics						
32.11	acetamide		CH_3CONH_2	1-1.5			
32.13	urea		$\text{CO}(\text{NH}_2)_2$				
O	'Precious' Non-minerals						
O.1	amber	succinite	approx. $\text{C}_{10}\text{H}_{16}\text{O}$	2-2.5	brittle	flammable	"becomes electrically charged when rubbed, attractnig small particles" (Schumann 1992: 32) like dust, etc.
O.2	pearl		CaCO_3				
O.3	obsidian						
O.4	coral		CaCO_3	3-4	brittle		

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1	Elements & Alloys						
1.1	copper	Cu	light	surfacial photo-oxidation	tarnish		Nassau 1992; Howie 1984
			RT, > ~60% RH	oxidation	tarnish	to cuprite, spertiniite, & tenorite	Platzman et al. 2008
			sulfur compounds	oxidation	black spots, pitting	to copper sulfides & sulfates	Eggert et al. 2004; Weichert et al. 2004
			moisture & chlorides	oxidation	green film or encrustations	initially to nantokite, & eventually to copper trihydroxychloride species	Scott 1990, 2000; Kologo et al. 2007
			hydrochloric acid	etching, rapid tarnish	colour change: rose-red; tarnish: dull		King 1983
			acetic acid	oxidation	green encrustation	verdigris formation	Kuhn 1970
1.2	silver	Ag	moist air, especially w/ S exposure	oxidation	tarnish: black		Howie 1992; O'Donoghue 1983
			sulfur compounds	oxidation	tarnish	formation of S-bearing compounds	Howie 1992; O'Donoghue 1983
1.24	diamond	C	light	colour change	chameleon (yellow): green; chameleon (red): pink		Nassau 1992
1.29	tin	Sn	< 13.2C	structural phase change	colour change: grey; potential blistering, cracking, and disintegration	to α-tin (cubic); reaction only occurs with significant undercooling and/or after a long period of time	Ojima et al. 1993; Zeng et al. 2014
			> 13.2C	structural phase change	colour change: white (when T > ~30C)	to β-tin (tetragonal); reaction only occurs with significant overheating and/or after a long period of time	Ojima et al. 1993; Zeng et al. 2014
1.30	lead	Pb	high RH nitric acid	oxidation	tarnish	corrosion	Howie 1992 O'Donoghue 1983
			carboxylic acids	oxidation	corrosion crust	to Pb-bearing carboxylates	Tétreault et al. 1998; Raychaudhuri & Brimblecombe 2000; Tétreault et al. 2003
1.33	arsenic	As	moist air acids	oxidation	tarnish: dark grey	to arsenolite	Howie 1992; King 1985 O'Donoghue 1983
1.37	domeykite	Cu ₃ As	moist air	oxidation	tarnish: grey, yellow, or black	sequential alteration	Howie 1992
			moist air	oxidation	tarnish: brown		Howie 1992
1.38	algodonite	(Cu _{1-x} As _x); x≈0.15	light	tarnish	dull, brown		Howie 1992; O'Donoghue 1983
			moist air	oxidation	tarnish: dull		Howie 1992

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Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
1.41	antimony	Sb	moist air	oxidation	tarnish: grey		Howie 1992
			sulfuric acid			formation of S-bearing compounds	O'Donoghue 1983
1.42	stibarsen	AsSb	moist air	oxidation	tarnish: grey or brown		Howie 1992
1.48	dyscrasite	Ag ₃ Sb	moist air	oxidation	tarnish: grey, yellow, or black		Howie 1992; O'Donoghue 1983
1.49	bismuth	Bi	moist air	oxidation	tarnish: pink, brass-yellow		O'Donoghue 1983; Howie 1992; Schumann 1992
1.51	sulfur	S ₈	sudden temperature changes	cracking, spawling			O'Donoghue 1983
1.57	iron	Fe	critical : > ~65% RH	oxidation		corrosion	Howie 1992; O'Donoghue 1983
1.61	nickel	Ni	critical: > ~85% RH	oxidation	tarnish	corrosion	Howie 1992
1.62	kamacite	α-(Fe,Ni)	water, oxygen, & chloride	electrochemical oxidation	cracks, pits	akaganeite precipitation	Buchwald & Clarke 1989
1.63	taenite	γ-(Ni,Fe)	water, oxygen, & chloride	electrochemical oxidation	cracks, pits	akaganeite precipitation	Buchwald & Clarke 1989
1.82	platinum	Pt		tarnish	colour change: darker		Schumann 1992
2	Carbides, Nitrides, Silicides and Phosphides						
2.3	osbornite	TiN	moist air	oxidation			Bannister 1937
2.14	schreibersite	(Fe,Ni) ₃ P	moist air	oxidation	tarnish: brassy yellow or brown		Bannister 1937
3	Sulfides, Selenides, Tellurides, Arsenides and Bismuthides						
3.1.1	chalcocite	Cu ₂ S	light	surficial photo-oxidation	tarnish: blue or green	to digenite	Nassau 1992; King 1985
			moist air	oxidation	tarnish: dull black	to copper sulfides & ultimately to digenite	Howie 1992; King 1985
3.1.8	covellite	CuS	moist air	oxidation	tarnish: blue or purple, iridescent		Howie 1992; O'Donoghue 1983
3.1.9	berzelianite	Cu ₂ Se	light	photo-oxidation	tarnish: dull		Howie 1992; Nassau 1992
			moist air	oxidation	tarnish: dull grey		Howie 1992; O'Donoghue 1983
3.1.11	umangite	Cu ₃ Se ₂	> 50% RH	oxidation	tarnish: dull violet-blue	to chalcocite	Howie 1992
3.1.14	klockmannite	CuSe	> 50% RH	oxidation	tarnish: blue		Howie 1992
3.1.16	weissite	Cu _{2-x} Te	moist air	oxidation	tarnish: darkens		Howie 1992
3.1.17	rickardite	Cu ₇ Te ₅	light	photo-oxidation	dull		O'Donoghue 1983
			moist air	oxidation	tarnish: darkens		Howie 1992
3.1.22	cubanite	CuFe ₂ S ₃	moist air	oxidation	tarnish: brown or blue		Howie 1992

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Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
3.1.23	bornite	Cu ₅ FeS ₄	moist air	oxidation	tarnish: purple-blue, iridescent	hydrated iron oxide	Howie 1992; O'Donoghue 1983
			oxalic acid		tarnish		King 1983
3.1.25	chalcopyrite	CuFeS ₂	low temperatures, > 50% RH	oxidation	tarnish: iridescent blue or red, dull brown, or black	to hydrated ferric oxide & copper sulfate	Howie 1992
			hydrochloric acid		tarnish		O'Donoghue 1983
3.2.1	acanthite	Ag ₂ S	light	surficial photo-oxidation	darken	sulfide formation	Nassau 1992; O'Donoghue 1983
			light	photochemical oxidation	tarnish: lead-grey, eventually grey friable crust	rapid surface degradation	Howie 1992
3.2.3	naumannite	Ag ₂ Se	light	surficial photo-oxidation	tarnish: brown, iridescent		Nassau 1992
3.2.4	agularite	Ag ₄ SeS	light	surficial photo-oxidation	tarnish: darkens	to silver & cuprian stephanite	Nassau 1992; King 1985
3.2.5	hessite	Ag ₂ Te	light	surficial photo-oxidation	tarnish: dull, yellow or grey		Howie 1992; Nassau 1992
3.2.8	cervelleite	Ag ₄ TeS	light	photochemical oxidation			Howie 1992
3.2.10	stromeyerite	AgCuS	light	surficial photo-oxidation	tarnish: dark blue		Nassau 1992
			moist air	oxidation	tarnish: blue	to metallic silver & covellite	Howie 1992; O'Donoghue 1983
3.2.11	mckinstryite	(Ag,Cu) ₂ S	moist air	oxidation			Howie 1992
3.2.12	jalpaite	Ag ₃ CuS ₂	light	surficial photo-oxidation	tarnish: darkens or iridescent	to silver & silver sulfosalts	Nassau 1992; King 1985
			moist air	oxidation	tarnish: dark grey or yellow		O'Donoghue 1983
3.2.13	eucairite	AgCuSe	moist air	oxidation	tarnish: light yellow		O'Donoghue 1983
3.2.17	argentopyrite	AgFe ₂ S ₃	moist air	oxidation	tarnish: brown, blue, green, or purple		O'Donoghue 1983
3.3.1	calaverite	AuTe ₂	light	photo-oxidation	tarnish: brassy-yellow		Howie 1992
3.3.2	krennerite	Au ₃ AgTe ₈	light	photo-oxidation	tarnish: brassy-yellow		Howie 1992
3.3.7	petzite	Ag ₃ AuTe ₂	light	photo-oxidation	tarnish: yellow-gold		Howie 1992
			moist air	oxidation	tarnish: darkens		O'Donoghue 1983
			nitric acid				O'Donoghue 1983
3.3.9	sylvanite	(Au,Ag) ₂ Te ₄	high light levels	surficial photo-oxidation	tarnish: brassy-yellow, darkens	to tellurite	Howie 1992; Nassau 1992
3.4.2	oldhamite	(Ca,Mg)S	moist air	oxidation		to calcium sulfate	Howie 1992
3.4.4	sphalerite	ZnS	light	photo-oxidation	darkens; metallic tarnish		Nassau 1992
			moist air	oxidation			Howie 1984

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3.4.7	greenockite	CdS	moist air hydrochloric acid	surficial oxidation & hydration	yellow/red: white, fading	to cadmium sulfate monohydrate	Van der Snickt 2009 O'Donoghue 1983
3.5.1	cinnabar	HgS	light, moisture, & alkali halogens	surficial photochemical oxidation	darkens & blackens	to metallic mercury & other mercuric halides	Keune & Boon 2005; Radeponet et al. 2011; Anaf et al. 2013; Neiman et al. 2015
3.6.5	galena	PbS	20C, > 65% RH moist air dilute acids	oxidation oxidation etching	tarnish: blue, dark grey	surficial lead enrichment to anglesite	Howie 1992 Swartzlow 1933 King 1983
3.6.7	altaite	PbTe	moist air	oxidation	tarnish: dull brown or bronze		Howie 1992; O'Donoghue 1983
3.7.4	realgar	As ₄ S ₄	light: 500-670nm	structural alteration	red: orange	to pararealgar	Douglass et al. 1992; Kyono et al. 2005; Kyono 2007; Jovanovski & Makreski 2020
3.7.8	orpiment	As ₂ S ₃	moist air	oxidation		to arsenolite	Howie 1992
3.7.14	stibnite	Sb ₂ S ₃	light	self-limiting photochemical oxidation	tarnish: dull, blue-black &/or iridescent		Howie 1992; Nassau 1992
3.7.17	bismuthinite	Bi ₂ S ₃	moist air	oxidation	tarnish: yellow or iridescent	to hydrated bismuth oxide	Howie 1992
3.7.29	tetradymite	Bi ₂ Te ₂ S	moist air	oxidation	tarnish: dull or iridescent	to bismutite & montanite	Howie 1992
3.7.30	joséite	Bi ₄ TeS ₂	moist air	oxidation	tarnish: dull or iridescent	to montanite	Howie 1992
3.8.4	daubréelite	Fe ²⁺ Cr ³⁺ ₂ S ₄	moist air	oxidation			Howie 1984; Bannister 1937
3.8.6	molybdenite	MoS ₂	high RH	oxidation			Howie 1992
3.8.12	alabandite	MnS	light moist air	surficial photo-oxidation oxidation	tanish: brown tarnish: brown	 to hydrated manganese sulfate	Nassau 1992 Howie 1992
3.8.13	hauerite	MnS ₂	hydrochloric acid				O'Donoghue 1983
3.9.1	pyrrhotite	Fe ₇ S ₈	20C, > 50% RH	oxidation	tarnish: dull, pink-gold or iridescent	to iron sulfates & sulfuric acid; some marcasite may form	Howie 1992; Jerz & Rimstidt 2003

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Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
3.9.3	pyrite	FeS ₂	> 30% RH; critical: > 60% RH	oxidation	tarnish: dull or iridescent; efflorescence	initially to sulfate & iron (hydr)oxides; then to iron sulfates & sulfuric acid	Howie 1992; Guevremont et al. 1998; Rosso et al. 1999; Jerz & Rimstidt 2004
			Arsenic compounds, anoxia	alteration		to an FeAsS surface precipitate, potentially evolving into scorodite or orpiment	Bostick & Fendorf 2003
3.9.4	marcasite	FeS ₂	light	tarnish	darkens		O'Donoghue 1983
			> 30% RH	oxidises	tarnish: dull or iridescent; efflorescence	to iron sulfates & sulfuric acid	Howie 1992; Lowson 1982
3.9.5	greigite	Fe ²⁺ Fe ³⁺ ₂ S ₄	moisture	oxidation	tarnish		Howie 1992
3.9.6	mackinawite	(Fe,Ni) ₉ S ₈	moisture	oxidation		to sulfates, FeO, & HS	Howie 1992
3.9.7	smythite	(Fe,Ni) _{3+x} S ₄ ; x≈0-0.3	moisture	oxidation	tarnish		Howie 1992
3.9.9	ferroselite	FeSe ₂	moisture	oxidation			Howie 1984
3.9.11	löllingite	FeAs ₂	moisture	oxidation		to scorodite & symplectite	Howie 1992
3.9.12	arsenopyrite	FeAsS	> 30C, high RH	oxidation	tarnish: dull brown or grey; occasional grey/black crystalline growths at fractures	to scorodite & hydrated iron sulfates	Howie 1992
3.9.NI	troilite	FeS	RT, ambient humidity	oxidation	tarnish	initially to sulfides, then sulfate	Terranova et al. 2018
			Arsenic compounds, anoxia	alteration		to pyrite & an FeAsS surface precipitate, potentially evolving into scorodite or orpiment	Bostick & Fendorf 2003
3.10.2	linnaeite	Co ²⁺ Co ³⁺ ₂ S ₄	> 40-50% RH	oxidation	tarnish: red or violet	to hydrated cobalt sulfates	Howie 1992
3.10.10	safflorite	(Co,Ni,Fe)As ₂	> 50% RH	oxidation	tarnish: darkens	to erythrite	Howie 1992
3.10.12	skutterudite	CoAs ₃	high RH	oxidation	tarnish	to nickel & cobalt arsenates	Howie 1992
			moisture	oxidation & hydration		to erythrite	King 1985
3.10.13	cobaltite	CoAsS	moisture	oxidation	tarnish: pink-blue	to cobalt sulfates, arsenates, and oxides	Howie 1992
3.10.17	glauco-dot	(Co _{0.5} Fe _{0.5})AsS	moisture	oxidation	tarnish: brown		Howie 1992
3.10.18	carrollite	Cu(Co,Ni) ₂ S ₄	> 40-50% RH	oxidation	tarnish: red or violet	to hydrated cobalt sulfates	Howie 1992
			> 40-50% RH	oxidation	tarnish: iridescent grey		Howie 1992
3.11.3	millerite	NiS	22-24C, > 60% RH	surficial oxidation		to theophrastite & nickel sulfate	Nesbitt & Reinke 1999
			acids		matted finish		O'Donoghue 1983
3.11.4	polydymite	Ni ²⁺ Ni ³⁺ ₂ S ₄	light	surficial photo-oxidation	colour change: red or violet	to erythrite & pitticite	Nassau 1992; King 1985
			> 40-50% RH	oxidation	tarnish: red or violet		Howie 1992
3.11.5	vaesite	NiS ₂	> 40-50% RH	oxidation	tarnish		Howie 1992

Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
3.11.8	violarite	$\text{Fe}^{2+}\text{Ni}^{3+}_2\text{S}_4$	moisture	oxidation	tarnish: red		Howie 1992
3.11.9	pentlandite	$(\text{Fe},\text{Ni})_9\text{S}_8$	moisture	oxidation	tarnish		Howie 1992
3.11.10	bravoite	$(\text{Fe},\text{Ni})\text{S}_2$	> 40-50% RH	oxidation	tarnish: red		Howie 1992
3.11.19	penrosite	$(\text{Ni},\text{Co},\text{Cu})\text{Se}_2$	light	surficial photo-oxidation	tarnish: dull		Nassau 1992
			moisture	oxidation	tarnish: dull		Howie 1992
3.11.22	melonite	NiTe_2	moisture	oxidation	tarnish: brown		Howie 1992
3.11.26	maucherite	$\text{Ni}_{11}\text{As}_8$	moisture	tarnish	tarnish: darkens		Parsons 1922
			> 30% RH	oxidation	tarnish: darkens, dull	to annabergite	Howie 1992
3.11.27	nickeline	NiAs	22-24C, > 60% RH	surficial oxidation		to theophrastite, annabergite, & nickel arsenite	Nesbitt & Reinke 1999
3.11.28	rammelsbergite	NiAs_2	> 30% RH	oxidation	green coating	to annabergite	Howie 1992
3.11.29	pararammelsbergite	NiAs_2	> 30% RH	oxidation		to annabergite	Howie 1992
3.11.30	nickelskutterudite	$(\text{Ni},\text{Co},\text{Fe})\text{As}_3$	high RH	oxidation	tarnish	to nickel arsenates & cobalt arsenates	Howie 1992
3.11.34	gersdorffite	NiAsS	moisture	oxidation	tarnish: darkens, dull		Howie 1992
			moisture	oxidation	disintegrates; green mass	to annabergite	King 1983
3.11.36	breithauptite	NiSb	moisture	oxidation	tarnish: violet		Howie 1992
3.11.38	ullmannite	NiSbS		tarnish	colour change: iridescence		Schumann 1992
4	Oxysulfides						
4.2	kermesite	$\text{Sb}_2\text{S}_2\text{O}$		dehydration			O'Donoghue 1983
5	Sulfosalts: Sulfarsenites & Sulfobismuthites						
5.1.4	chalcostibite	CuSbS_2	light & moisture	photo-oxidation	tarnish: dull, pale blue		Howie 1992
5.1.7	emphletite	CuBiS_2		tarnish	colour change: yellow		Schumann 1992
5.1.9	wittichenite	Cu_3BiS_3	moisture	oxidation	tarnish: pale lead-grey; surface alteration (yellow-brown, blue, red)	to bismutite & covellite	Howie 1992
5.1.10	hodrušite	$\text{Cu}_8\text{Bi}_{12}\text{S}_{22}$	light	photo-oxidation	tarnish: brown bronze		Howie 1984
5.2.1	smithite	AgAsS_2	light	photochemical oxidation	red: orange	loss of form	Howie 1992; Nassau 1992; King 1985
5.2.2	trechmannite	AgAsS_2	light	surficial photo-oxidation			Nassau 1992
5.2.3	proustite	Ag_3AsS_3	light	photo-oxidation	colour change: darken & silvery; ultimately turns black	to silver & acanthite	Howie 1992; Nassau 1992; King 1983, 1985
5.2.4	xanthoconite	Ag_3AsS_3	light	surficial photo-oxidation	tarnish		Nassau 1992
5.2.5	pearceite	$[\text{Ag}_9\text{CuS}_4][(\text{Ag},\text{Cu})_6(\text{As},\text{Sb})_2\text{S}_7]$	light	surficial photo-oxidation	tarnish		Nassau 1992

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Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
5.2.8	miargyrite	AgSbS ₂	light	surficial photo-oxidation	darkens		Nassau 1992; O'Donoghue 1983
5.2.9	pyrargyrite	Ag ₃ SbS ₃	light	photo-oxidation	tarnish: darkens, dull	to metallic silver, acanthite, & stibnite	Howie 1992; Nassau 1992; King 1985
5.2.10	pyrostilpnite	Ag ₃ SbS ₃	light	surficial photo-oxidation	tarnish		Nassau 1992
5.2.11	stephanite	Ag ₅ SbS ₄	light	surficial photo-oxidation	tarnish: darkens	to silver & other silver sulfosalts	Nassau 1992; King 1985
5.2.12	polybasite	[(Ag,Cu) ₆ (Sb,As) ₂ S ₇][Ag ₉ CuS ₄]	light	surficial photo-oxidation	tarnish	to stephanite	Nassau 1992; King 1985
5.2.16	aramayoite	Ag ₃ Sb ₂ (Bi,Sb)S ₆	high light levels	surficial photo-oxidation	tarnish		Nassau 1992
5.2.17	matildite	AgBiS ₂	light	surficial photo-oxidation	tarnish		Nassau 1992
5.2.19	pavonite	AgBi ₃ S ₅	light	photo-oxidation	tarnish		Howie 1984
5.2.24	volynskite	AgBiTe ₂	moisture	oxidation			Howie 1984
5.2.NI	fettelite	[Ag ₆ As ₂ S ₇][Ag ₁₀ HgAs ₂ S ₈]	light	surficial photo-oxidation	tarnish		Nassau 1992
5.5.1	lorándite	TlAsS ₂	light	surficial photo-decomposition	translucent orange or yellow powdery coating		Howie 1992
5.5.7	hutchinsonite	TlPbAs ₅ S ₉	light	surficial photo-oxidation	tarnish		Nassau 1992
5.5.10	vrbaite	Hg ₃ Tl ₄ As ₈ Sb ₂ S ₂₀	light	surficial photo-oxidation			Nassau 1992
5.6.1	sartorite	PbAs ₂ S ₄	light	surficial photo-oxidation			Nassau 1992
5.6.2	dufrénoysite	Pb ₂ As ₂ S ₅	light	photo-oxidation	surficial colour change: brown		Nassau 1992
5.6.4	baumhauerite	Pb ₃ As ₄ S ₉	light	surficial photo-oxidation	tarnish: iridescent		Howie 1992; Nassau 1992
			moisture	oxidation	tarnish: iridescent		Howie 1992
5.6.6	jordanite	Pb ₁₄ (As,Sb) ₆ S ₂₃	moisture	oxidation	tarnish: iridescent		Howie 1992; Schumann 1992
5.6.7	fülöppite	Pb ₃ Sb ₈ S ₁₅	moisture	oxidation	tarnish: dull bronze		Howie 1992
5.6.8	zinkenite	Pb ₉ Sb ₂₂ S ₄₂	moisture	oxidation	tarnish: iridescent		Howie 1992
5.6.13	semseyite	Pb ₉ Sb ₈ S ₂₁	moisture	oxidation	tarnish: dull steel blue		Howie 1992
5.6.15	boulangerite	Pb ₅ Sb ₄ S ₁₁	low RH	oxidation	yellow spots	to bindheimite	Howie 1992
5.6.23	galenobismutite	PbBi ₂ S ₄	moisture	oxidation		to bismutite	Howie 1992
5.7.2	bournonite	PbCuSbS ₃	light	alteration	darkens		Howie 1992
5.7.9	aikinite	PbCuBiS ₃	moisture	oxidation	tarnish: brown or copper-red; occasional yellow coating	to bismutite	Howie 1992

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Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
5.7.20	marrite	AgPbAsS ₃	moisture	oxidation	tarnish: iridescent		Howie 1992
5.7.22	andorite	AgPbSb ₃ S ₆	light	photo-oxidation	tarnish: yellow, iridescent		Howie 1992; Nassau 1992
5.7.23	ramdohrite	Pb _{5.9} Fe _{0.1} Mn _{0.1} In _{0.1} Cd _{0.2} Ag _{2.8} Sb _{10.8} S ₂₄	light	surfical photo-oxidation			Nassau 1992
5.7.27	fizélyite	Ag ₅ Pb ₁₄ Sb ₂₁ S ₄₈	light	surfical photo-oxidation	tarnish		Nassau 1992
5.7.28	owyheeite	Ag ₃ Pb ₁₀ Sb ₁₁ S ₂₈	moisture	oxidation	tarnish: yellow		O'Donoghue 1983
5.7.29	freieslebenite	AgPbSbS ₃	light	surfical photo-oxidation	tarnish		Nassau 1992
5.7.30	diaphorite	Ag ₃ Pb ₂ Sb ₃ S ₈	light	surfical photo-oxidation	tarnish		Nassau 1992
5.7.39	benjaminite	Ag ₃ Bi ₇ S ₁₂	moisture	oxidation	tarnish: dull yellow or red	to bismutite	Howie 1992
5.7.40	lengenbachite	Ag ₄ Cu ₂ Pb ₁₈ As ₁₂ S ₃₉	light	surfical photo-oxidation	tarnish: iridescent		Nassau 1992
5.7.47	rathite	Ag ₂ Pb _{12-x} Tl _{x/2} As _{18+x/2} S ₄₀	light	surfical photo-oxidation	tarnish: iridescent	to arsenical sulfides	Nassau 1992; King 1985
5.8.1	samsonite	Ag ₄ MnSb ₂ S ₆	light	surfical photo-oxidation	tarnish		Nassau 1992
5.8.4	berthierite	FeSb ₂ S ₄	moisture	oxidation	tarnish: iridescent or dark brown		Howie 1992; O'Donoghue 1983
5.8.5	jamesonite	Pb ₄ FeSb ₆ S ₁₄	moisture	oxidation	tarnish: iridescent		Howie 1992; Schumann 1992
6	Sulfosalts						
6.1.6	stannite	Cu ₂ FeSnS ₄	moisture	oxidation	tarnish: blue		Howie 1992
6.1.18	franckeite	Fe ²⁺ (Pb,Sn ²⁺) ₆ Sn ⁴⁺ ₂ Sb ₂ S ₁₄	moisture	oxidation	tarnish: iridescent		Howie 1992
6.2.1	germanite	Cu ₁₃ Fe ₂ Ge ₂ S ₁₆		tarnish	colour change: dark violet, brown		Schumann 1992
6.2.4	argyrodite	Ag ₈ GeS ₆	light	surficial photo-oxidation	tarnish: violet		Howie 1992; Nassau 1992
6.2.5	canfieldite	Ag ₈ SnS ₆	light	surficial photo-oxidation	tarnish: blue or purple		Howie 1992; Nassau 1992
6.3.1	enargite	Cu ₃ AsS ₄	low temperatures, > 50% RH	oxidation	tarnish: dull	to tennantite	Howie 1992
6.4.1	sulvanite	Cu ₃ VS ₄	sulfur compounds	oxidation		develops black efflorescence	Howie 1992
6.5.1	corderoite	α-Hg ²⁺ ₃ S ₂ Cl ₂	moisture	oxidation	tarnish: dull	to cuprodescloizite	Howie 1992
6.5.1	corderoite	α-Hg ²⁺ ₃ S ₂ Cl ₂	light	desulfidation	orange-pink / pink-brown: grey, black	to calomel, metallic mercury, & sulfur	Keune & Boon 2005; Da Pieve et al. 2013
6.5.NI	kenhsuite	γ-Hg ²⁺ ₃ S ₂ Cl ₂	light	alteration		to corderoite	Radepon et al. 2011

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Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
7	Oxides & Hydroxides						
7.3.1	cuprite	Cu ₂ O	light	photodecomposition	darkens	copper liberated	Nassau 1992
			light	surficial photo-oxidation	tarnish		Nassau 1992; O'Donoghue 1983; Howie 1984
			chlorides	oxidation & hydrolysis	colour change: pale green encrustation	to copper trihydroxylchloride species	Sharkey & Lewin 1971; Scott 1990, 2000
7.4.7	periclase	MgO	moisture	tarnish	dull		O'Donoghue 1983
7.5.12	montroydite	HgO	light	surficial photo-oxidation	darkens		Nassau 1992
7.6.1	corundum	Al ₂ O ₃	high light levels	colour change	yellow: fades		Nassau 1992
7.8.1	quartz	SiO ₂	light, inc. sunlight	colour change (var. agate, amethyst, chrysoprase, citrine, morion, rose (crystalline), smokey)	fades		Currier 1985; Kane 1985; King 1985; Nassau 1992
7.8.8	opal	SiO ₂ · nH ₂ O	moist air	hydration			Howie 1984
			dry air &/or mild heating	dehydration	may show signs of shrinkage; cracks and loss of opalescence		Schumann 1977; King 1985
7.9.2	rutile	TiO ₂	light	colour change	darkens		Nassau 1992
7.11.20	massicot	β-PbO	high RH	carbonation		to (hydro)cerussite	Aze et al. 2007
7.11.21	litharge	PbO	sulfuric acid	sulfidation	yellow: black	to galena	Smith & Clark 2002
			sulfuric acid	sulfidation	red: black	to galena	Smith & Clark 2002
7.11.22	minium	Pb ₃ O ₄	sulfuric acid	sulfidation	red: black	to galena &/or other lead oxides	Smith & Clark 2002
			sulfuric acid	sulfidation		to plattnerite & anglesite	Aze et al. 2007
7.11.23	plattnerite	PbO ₂	acids				O'Donoghue 1983
7.12.8	navajoite	(V ⁵⁺ , Fe ³⁺) ₁₀ O ₂₄ · 12H ₂ O		dehydration		to lower hydration states	Waller 1992
7.12.10	bariandite	Al _{0.6} (V ⁵⁺ , V ⁴⁺) ₈ O ₂₀ · 9H ₂ O		dehydration		to lower hydration states	Waller 1992
7.13.1	arsenolite	As ₂ O ₃	acids				O'Donoghue 1983
7.13.6	stibiconite	Sb ³⁺ Sb ⁵⁺ ₂ O ₆ (OH)	moisture	oxidation	white coating	stibiconite	O'Donoghue 1983
			acids				O'Donoghue 1983
7.16.1	uraninite	UO ₂	radiogenic Pb ⁴⁺	auto-oxidation	colour change: black	increases U ⁶⁺ content	Finch & Murakami 2000
			oxygen & moisture	oxidation	corrosion crust	to uranyl oxyhydroxides	Finch & Murakami 2000
7.16.2	schoepite	(UO ₂) ₈ O ₂ (OH) ₁₂ · 12H ₂ O	RT, ambient humidity	dehydration	cracking along grain boundaries	to metaschoepite	Sowder et al. 1999
			RT, ambient humidity, & mechanical stress	dehydration	structural instability	irreversibly to paulscherrerite	Sowder et al. 1999
7.16.3	paraschoepite	UO ₃ · 2H ₂ O		dehydration		to anhydrous phase	Waller 1992
7.16.4	metaschoepite	(UO ₂) ₈ O ₂ (OH) ₁₂ · 10H ₂ O	22C, < 45% RH	dehydration		irreversibly to paulscherrerite	Kirkegaard et al. 2019

Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
7.16.7	ianthinite	$U^{4+}(UO_2)_5O_7 \cdot 10H_2O$	oxygen & moisture	oxidation	colour change (purple): lemon/sulfur yellow	to schoepite &/or metaschoepite	Finch & Murakami 2000; Singh et al. 2014
7.16.21	davidite-(La)	$La(Y,U)Fe_2(Ti,Fe,Cr,V)_{18}(O,OH,F)_{38}$	moisture	oxidation	tarnish		O'Donoghue 1983
7.16.22	davidite-(Ce)	$Ce(Y,U)Fe_2(Ti,Fe,Cr,V)_{18}(O,OH,F)_{38}$	moisture	oxidation	tarnish		O'Donoghue 1983
7.16.28	vandendriesscheite	$PbU_7O_{22} \cdot 12H_2O$		dehydration		to metavandendriesscheite	Waller 1992
7.17.2	tellurite	TeO_2	most acids				O'Donoghue 1983
7.18.1	manganosite	MnO	light	surficial reaction	colour change: blue		Nassau 1992
			light	surficial reaction	colour change: darkens		O'Donoghue 1983
7.18.7	manganite	$Mn^{3+}O(OH)$	moisture	oxidation	tarnish: darkens		O'Donoghue 1983
7.18.10	pyrochroite	$Mn(OH)_2$	light	surficial reaction	colour change: brown or darkens		Nassau 1992
7.20.4	hematite	Fe_2O_3	20C, > 90% RH	oxidation		to goethite	Bedarida & Pedemont 1971
			sulfuric acid	alteration		to iron sulfides	Drosdoff & Truog 1935
7.20.5	goethite	$\alpha-Fe^{3+}O(OH)$	sulfuric acid	alteration		to iron sulfides	Drosdoff & Truog 1935
7.20.6	akaganeite	$(Fe^{3+},Ni^{2+})_8(OH,O)_{16}Cl_{1.25} \cdot nH_2O$		dehydrates		loss of 5-10% wt H_2O	Buchwald & Clarke 1989
			moist air	chlorine depletion		decomposes to maghemite & goethite	Buchwald & Clarke 1989
8	Halides						
8.1.3	halite	$NaCl$	20C, > 75% RH	deliquescence			Waller 1992
			light	photodecomposition (var. huantajayite & other silver-bearing specimens)	darkens	silver liberation	Parsons 1922; King 1985
8.1.4	hydrohalite	$NaCl \cdot 2H_2O$	< -0.1C	dehydration	to saturated solution of halite		Waller 1992; Hazen & Ausubel 2016
8.1.5	carobbiite	KF	25C, > 30% RH	deliquescence			Waller 1992
				efflorescence			King 1985
8.1.6	sylvite	KCl	20C, > 85% RH	deliquescence			Waller 1992
8.1.7	salammoniac	NH_4Cl	20C, > 79% RH	deliquescence			Waller 1992
8.2.1	nantokite	$CuCl$	light	surficial photo-oxidation	green encrustation	to paratacamite	Nassau 1992; King 1985
				deliquescence			King 1985
			moist air	hydration		to eriochalcite	Scott 2000
			20C, > 70% RH	oxidation & hydrolysis	green encrustation	to copper trihydroxychloride species	Scott 1990, 2000
			moisture & sulfur compounds	alteration		to connellite &/or brochantite	Scott 2000
8.2.2	tolbachite	$CuCl_2$	moist air	hydration		to eriochalcite	Waller 1992

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Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
8.2.3	eriochalcite	$\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$	20C, > 68% RH	deliquescence dehydration	efflorescence	to lower hydration states	Waller 1992 King 1985; Howie 1984
8.2.6	botallackite	$\text{Cu}_2\text{Cl}(\text{OH})_3$	moist air	structural phase change		to other copper trihydroxychloride species	Scott 2000
8.2.8	calumetite	$\text{CaCu}_4(\text{OH})_8\text{Cl}_2 \cdot 3.5\text{H}_2\text{O}$		dehydration		to lower hydration states	Scott 2000
8.2.9	anthonyite	$\text{Cu}(\text{OH},\text{Cl})_2 \cdot 3\text{H}_2\text{O}$		dehydration		to lower hydration states	Waller 1992; Scott 2000
8.2.11	marshite	CuI	light	surficial photo-oxidation	darkens: pale pink through brownish red		Nassau 1992; O'Donoghue 1983
8.3.1	chlorargyrite	AgCl	high light levels	photodecomposition	colour change: violet-brown, grey, black	to silver; silver liberation	Nassau 1992; King 1985
8.3.2	bromargyrite	AgBr	high light levels	photodecomposition	darkens	to silver; silver liberation	Nassau 1992; King 1985
8.3.3	bromian chlorargyrite	$\text{Ag}(\text{Cl}, \text{Br})$	high light levels	photodecomposition	darkens	to silver; silver liberation	Nassau 1992; King 1985
8.3.4	iodargyrite	AgI	light	photodecomposition	darkens	to silver; silver liberation	Nassau 1992; King 1985
8.3.5	miersite	$(\text{Ag}, \text{Cu})\text{I}$	high light levels	photodecomposition	darkens	silver liberation	Nassau 1992
8.4.3	bischofite	$\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$	20C, > 33% RH 25C, < 3% RH	deliquescence dehydration	efflorescence	decomposes to lower hydration states	Waller 1992 Waller 1992
8.4.5	carnallite	$\text{KMgCl}_3 \cdot 6\text{H}_2\text{O}$	ambient conditions 32C, < 3% RH	deliquescence dehydration	dull efflorescence		Waller 1992; King 1985 Schumann 1992 Waller 1992
8.4.6	tachyhydrite	$\text{CaMg}_2\text{Cl}_6 \cdot 12\text{H}_2\text{O}$		deliquescence			Waller 1992; Hazen & Ausubel 2016
8.4.7	fluorite	CaF_2	light acids	colour change	pink: fades; green: purple; blue, purple: pink or colourless		Nassau 1992; King 1985 King 1983
8.4.8	sinjarite	$\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$	25C, > 11% RH	hydration		to higher hydration states, eventually to antarcticite	Waller 1992
8.4.9	antarcticite	$\text{CaCl}_2 \cdot 6\text{H}_2\text{O}$	> 29.9C RT, > 33% RH RT, < 21% RH	dehydration deliquescence dehydration		to saturated solution of sinjarite	Waller 1992 Waller 1992 Waller 1992
8.4.10	chlorocalcite	KCaCl_3		deliquescence			O'Donoghue 1983; Hazen & Ausubel 2016
8.4.NI	chloromagnesite	MgCl_2		deliquescence efflorescence			Howie 1984 Howie 1984
8.5.2	calomel	$[\text{Hg}_2]^{2+}\text{Cl}_2$	light	disproportionation		to metallic mercury & mercuric chloride	Neiman et al. 2015
8.5.3	terlinguaite	Hg_2OCl	light	surficial photo-oxidation	yellow: olive green		Nassau 1992; King 1985
8.5.6	eglestonite	$([\text{Hg}^{1+}]_2)_3\text{OCl}_3(\text{OH})$	light	surficial photo-oxidation	colour change: brown or darkens	mercury liberation	Nassau 1992; King 1985
8.6.2	chloraluminite	$\text{AlCl}_3 \cdot 6\text{H}_2\text{O}$	25C, > ~40% RH	deliquescence hydrolysis			Waller 1992 King 1985; Howie 1984

Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
8.6.3	cadwaladerite	$\text{AlCl}(\text{OH})_2 \cdot 4\text{H}_2\text{O}$		deliquescence			Waller 1992
8.7.1	fluocerite-(Ce)	$(\text{Ce}, \text{La})\text{F}_3$	light	colour change	yellowish- or reddish-brown		O'Donoghue 1983
8.8.17	bideauxite	$\text{Pb}_2\text{AgCl}_3(\text{F}, \text{OH})_2$	high light levels	colour change	pale lavender		Howie 1984
8.10.1	scacchite	MnCl_2	25C, > ~56% RH	deliquescence			Waller 1992
8.10.3	chloromanganokalite	K_4MnCl_6		deliquescence		decomposes easily	Waller 1992; O'Donoghue 1983
				efflorescence			Howie 1984
8.11.1	molysite	FeCl_3	25C, > ~5% RH	deliquescence			Waller 1992; Hazen & Ausubel 2016
				hydrolysis		to hydrous iron oxide	King 1985
8.11.2	hydromolysite	$\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$	25C, > ~5% RH	deliquescence			Waller 1992
8.11.3	rokühnite	$\text{FeCl}_2 \cdot 2\text{H}_2\text{O}$	moist air	hydration		to higher hydration states	Waller 1992
8.11.4	douglasite	$\text{K}_2\text{Fe}^{2+}\text{Cl}_4 \cdot 2\text{H}_2\text{O}$		deliquescence			Waller 1992; King 1985
				efflorescence			Howie 1984
8.11.5	erythrosiderite	$\text{K}_2\text{Fe}^{3+}\text{Cl}_5 \cdot \text{H}_2\text{O}$		deliquescence			Waller 1992; Hazen & Ausubel 2016
				efflorescence			Howie 1984
8.11.6	rinneite	$\text{K}_3\text{NaFe}^{2+}\text{Cl}_6$		deliquescence			Waller 1992; King 1985
8.11.7	kremersite	$(\text{NH}_4, \text{K})_2\text{Fe}^{3+}\text{Cl}_5 \cdot \text{H}_2\text{O}$		deliquescence			Waller 1992; King 1985
				efflorescence			Howie 1984
8.11.9	nickelbischofite	$\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$	20C, > ~54% RH	deliquescence			Waller 1992
			20C, < 43% RH	dehydration	efflorescence	to lower hydration states	Waller 1992
8.11.10	lawrencite	$(\text{Fe}^{2+}, \text{Ni})\text{Cl}_2$	moisture	oxidation		to ferric chloride	Bannister 1937
			25C, > 56% RH	deliquescence			Waller 1992
9	Borates						
9.1.6	kernite	$\text{Na}_2[\text{B}_4\text{O}_6(\text{OH})_2] \cdot 3\text{H}_2\text{O}$	RT, moist air	hydration		to borax	Muessig & Allen 1957
				dehydration	efflorescence	to lower hydration states	King 1985; Howie 1984
9.1.8	tincalconite	$\text{Na}_2(\text{B}_4\text{O}_7) \cdot 5\text{H}_2\text{O}$	20C, > 50% RH	hydration		to borax	Waller 1992
9.1.9	borax	$\text{Na}_2(\text{B}_4\text{O}_5)(\text{OH})_4 \cdot 8\text{H}_2\text{O}$	20C, > 99% RH	deliquescence			Waller 1992
			20C, < 50% RH	dehydration	opaque white efflorescence	to tincalconite	Waller 1992
9.1.14	larderellite	$(\text{NH}_4)\text{B}_5\text{O}_7(\text{OH})_2 \cdot \text{H}_2\text{O}$		dehydration	efflorescence		Howie 1984
9.2.17	iquiqueite	$\text{K}_3\text{Na}_4\text{Mg}(\text{CrO}_4)\text{B}_{24}\text{O}_{39}(\text{OH}) \cdot 12\text{H}_2\text{O}$	> 40C	dehydration		loss of absorbed & fluid-inclusion water	Ericksen et al. 1986
9.3.12	inoite	$\text{Ca}(\text{H}_4\text{B}_3\text{O}_7)(\text{OH}) \cdot 4\text{H}_2\text{O}$		dehydration	efflorescence	to meyerhofferite	O'Donoghue 1983; King 1985
9.3.14	colemanite	$\text{Ca}[\text{B}_3\text{O}_4(\text{OH})_3] \cdot \text{H}_2\text{O}$	< ~0C	structural phase change			Perloff & Block 1960
9.3.20	probertite	$\text{NaCaB}_5\text{O}_7(\text{OH})_4 \cdot 3\text{H}_2\text{O}$		dehydration	efflorescence	to lower hydration states	Waller 1992; King 1985

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Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
10	Borates with other anions						
10.1.2	bandylite	$\text{Cu}[\text{B}(\text{OH})_4]\text{Cl}$		deliquescence			Waller 1992
10.1.5	boracite	$\text{Mg}_3(\text{B}_7\text{O}_{13})\text{Cl}$		efflorescence			Howie 1984; King 1985
11	Carbonates						
11.1.2	natrite	Na_2CO_3	32C 25C, > 22% RH	melts hydration		to thermonatrite	Schumann 1992 Waller 1992
11.1.5	thermonatrite	$\text{Na}_2\text{CO}_3 \cdot \text{H}_2\text{O}$	25C, > 66% RH 25C, < 22% RH	deliquescence hydration dehydration		to higher hydration states to natrite	King 1985 Waller 1992 Waller 1992
11.1.6	trona	$\text{Na}_3\text{H}(\text{CO}_3)_2 \cdot 2\text{H}_2\text{O}$		deliquescence dehydration	efflorescence	to lower hydration states	Howie 1984 Waller 1992; King 1985
11.1.7	natron	$\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$	> 32C 25C, > 87% RH 25C, < 76% RH	dehydration deliquescence dehydration		to saturated solution of thermonatrite to lower hydration states, eventually to thermonatrite	Waller 1992 Waller 1992 Waller 1992
11.1.10	nyerereite	$\text{Na}_2\text{Ca}(\text{CO}_3)_2$		deliquescence			Waller 1992
11.1.13	pirssonite	$\text{Na}_2\text{Ca}(\text{CO}_3)_2 \cdot 2\text{H}_2\text{O}$		dehydration	efflorescence	to nyerereite	Waller 1992; King 1985
11.1.14	gaylussite	$\text{Na}_2\text{Ca}(\text{CO}_3)_2 \cdot 5\text{H}_2\text{O}$	> 63C	dehydration	efflorescence	to nyerereite	Waller 1992; Johnson & Robb 1973
11.1.16	fairchildite	$\text{K}_2\text{Ca}(\text{CO}_3)_2$		deliquescence hydration			Waller 1992 King 1985
11.1.17	bütschliite	$\text{K}_2\text{Ca}(\text{CO}_3)_2$	moist air	deliquescence		eventual decomposition	Waller 1992; King 1985; Pabst 1974
11.1.18	teschemacherite	$(\text{NH}_4)\text{HCO}_3$		volatization deliquescence			Parsons 1922 Howie 1984
11.2.1	malachite	$\text{Cu}_2(\text{CO}_3)(\text{OH})_2$	pH 4-6 sulfuric acid chlorides	hydrolysis alteration alteration	blackens	to paratacamite & atacamite to covellite to paratacamite or atacamite	Sharkey & Lewin 1971 Smith & Clark 2002 Sharkey & Lewin 1971
11.2.2	azurite	$\text{Cu}_3(\text{CO}_3)_2(\text{OH})_2$	light sulfuric acid	colour change alteration deliquescence	darkens blackens	to covellite	King 1985 Smith & Clark 2002 King 1985
11.3.3	nesquehonite	$\text{MgCO}_3 \cdot 3\text{H}_2\text{O}$	RT, ambient humidity	dissociation		loss of water & carbon dioxide	Robie & Hemingway 1972
11.3.4	lansfordite	$\text{MgCO}_3 \cdot 5\text{H}_2\text{O}$	< ~10C low RH	dehydration dehydration		to saturated solution of nesquehonite to nesquehonite	Waller 1992 King 1985
11.3.7	hydromagnesite	$\text{Mg}_5(\text{CO}_3)_4(\text{OH})_2 \cdot 4\text{H}_2\text{O}$		dehydration	efflorescence	to lower hydration states	Waller 1992; King 1985

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11.4.1	calcite	CaCO ₃	light	colour change	fades		Nassau 1992
			light	surfacial reaction	contains Fe/Co/Mn: darkens		Nassau 1992
			> ~10%RH	surfacial hydration		formation of a hydrated layer of strongly chemisorbed hydroxyl groups (Ca(OH)(CO ₃ H))	Stipp et al. 1996; Al-Hosney & Grassian 2004, 2005; Usher et al. 2007
			> 70%RH	hydration & surface reconstruction		to vaterite	Baltrusaitis & Grassian 2009
			carboxylic acids	alteration	efflorescence	to calcium acetate &/or calcium formate (depending on acid present)	Al-Hosney & Grassian 2004; Al-Hosney et al. 2005; Usher et al. 2007; Prince et al. 2008
			nitric acid	alteration		droplets or liquid film of calcium nitrate	Krueger 2003; Al-Hosney & Grassian 2004, 2005; Prince et al. 2008
			sulfuric acid	alteration	efflorescence	to hannebachite & eventually gypsum	Al-Hosney & Grassian 2004, 2005; Rettger 1924
11.4.4	monohydrocalcite	CaCO ₃ · H ₂ O		dehydration		to calcite	Waller 1992
11.4.5	ikaite	CaCO ₃ · 6H ₂ O		dehydration	white efflorescence	to vaterite &/or calcite	Baltrusaitis & Grassian 2009
11.7.2	hydroscarbroite	Al ₁₄ (CO ₃) ₃ (OH) ₃₆ · nH ₂ O		dehydration		to lower hydration states	Waller 1992; Hazen & Ausubel 2016
11.7.13	hydrodresserite	BaAl ₂ (CO ₃) ₂ (OH) ₄ · 3H ₂ O		dehydration		to dresserite	Waller 1992; Hazen & Ausubel 2016
11.8.3	lanthanite-(Ce)	(Ce,La,Nd) ₂ (CO ₃) ₃ · 8H ₂ O		dehydration	efflorescence	to lower hydration states	King 1985
11.9.1	cerussite	PbCO ₃		deliquescence			Howie 1984
			sulfuric acid	partial alteration	white: dark grey	to galena	Smith & Clark 2002
11.9.2	hydrocerussite	Pb ₃ (CO ₃) ₂ (OH) ₂		surfacial alteration	white: black	to galena	Smith & Clark 2002
11.11.3	swartzite	MgCa(UO ₂)(CO ₃) ₃ · 12H ₂ O		dehydration	yellowish-white	to lower hydration states	Waller 1992; O'Donoghue 1983
11.11.5	bayleyite	Mg ₂ (UO ₂)(CO ₃) ₃ · 18H ₂ O		dehydration	pales	to lower hydration states	Waller 1992; O'Donoghue 1983
11.11.7	zellerite	Ca(UO ₂)(CO ₃) ₂ · 5H ₂ O		dehydration		to metazellerite	Waller 1992
11.11.8	wyartite	CaU ⁵⁺ (UO ₂) ₂ (CO ₃)O ₄ (OH) · 7H ₂ O		oxidation		to schoepite or metaschoepite	Gauthier et al. 1989
11.12.1	rhodochrosite	MnCO ₃	light	surfacial reaction	darkens		Nassau 1992
11.13.1	siderite	FeCO ₃	moisture	oxidation	darkens: blackish-brown; occasional iridescent tarnish	to goethite & hematite	Schaller & Vlisidis 1959; Palache et al. 1951
11.14.4	hellyerite	NiCO ₃ · 5.5H ₂ O		dehydration		to lower hydration states	Waller 1992

Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
12	Carbonates with other anions: with halides or sulfate						
12.1.2	barentsite	Na ₇ Al(HCO ₃) ₂ (CO ₃) ₂ F ₄		deliquescence			Waller 1992
12.2.1	burkeite	Na ₆ (CO ₃)(SO ₄) ₂	RT, > ~75% RH	deliquescence			Waller 1992
12.2.2	hanksite	Na ₂₂ K(SO ₄) ₉ (CO ₃) ₂ Cl		deliquescence			Howie 1984
			moist air	hydration		to higher hydration states	King 1985
				efflorescence			Howie 1984; Bannister 1937
13	Nitrates						
13.1	nitratine	NaNO ₃	20C, >74% RH	deliquescence			Waller 1992
13.2	niter	KNO ₃	25C, >92% RH	deliquescence			Waller 1992
13.3	gerhardtite	Cu ₂ (NO ₃)(OH) ₃		deliquescence			King 1985
13.4	nitromagnesite	Mg(NO ₃) ₂ · 6H ₂ O	25C, >52% RH 25C, <2% RH	deliquescence dehydration	efflorescence		Waller 1992 Waller 1992
13.5	nitrocalcite	Ca(NO ₃) ₂ · 4H ₂ O	> 42.7C	dehydration		to saturate solution of Ca(NO3)2 · 3H2O	Waller 1992
			19C, > 54% RH	deliquescence			Waller 1992
			20C, < 6% RH	dehydration	efflorescence	to lower hydration states	Waller 1992
13.6	nitrobarite	Ba(NO ₃) ₂	20C, > 99% RH	deliquescence			Waller 1992
13.9	darapskite	Na ₃ (SO ₄)(NO ₃) · H ₂ O		deliquescence			Waller 1992; King 1985
13.12	hydrombobomkulite	(Ni,Cu)Al ₄ ((NO ₃) ₂ ,SO ₄)(OH) ₁₂ · 13-14H ₂ O		dehydration		to mbobomkulite	Waller 1992
13.NI	gwihabaite	(NH ₄ ,K)NO ₃		deliquescence			Hazen & Ausubel 2016
14	Silicates not containing aluminum						
14.2.5	chrysocolla	Cu _{2-x} Al _x (H _{2-x} Si ₂ O ₅)(OH) ₄ · nH ₂ O, x < 1		dehydration	may show signs of shrinkage		Howie 1984; King 1985
14.3.1	phenakite	Be ₂ SiO ₄	light	colour change	red: pale pink; yellow, orange, brown: colourless		King 1985; Nassau 1992
14.4.11	sepiolite	Mg ₄ (Si ₆ O ₁₅)(OH) ₂ · 6H ₂ O	moist air	hydration dehydration		to higher hydration states to lower hydration states	Howie 1984 Howie 1984
14.10.1	zircon	Zr(SiO ₄)	light	colour change	flesh-colour: blue or grey		King 1985; Nassau 1992
14.12.4	pabstite	Ba(Sn,Ti)(Si ₃ O ₉)	light	colour change	pink: fades		Nassau 1992
14.13.8	barysilite	Pb ₈ Mn ²⁺ [Si ₂ O ₇] ₃	moisture	oxidation	tarnish		O'Donoghue 1983
14.17.1	rhodonite	Mn ²⁺ SiO ₃	light	surficial reaction	darkens		Nassau 1992
14.17.8	bementite	Mn ₇ Si ₆ O ₁₅ (OH) ₈	moisture	oxidation	tarnish: darkens		O'Donoghue 1983
14.18.21	inesite	Ca ₂ (Mn,Fe) ₇ Si ₁₀ O ₂₈ (OH) ₂ · 5H ₂ O	moisture	oxidation	tarnish: darkens, brown		O'Donoghue 1983
14.19.1	fayalite	Fe ²⁺ ₂ SiO ₄	light	surficial reaction	green: brown		Nassau 1992
			light	colour change	green: blue		King 1985; Nassau 1992

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14.19.8	hisingerite	$\text{Fe}^{3+}_2(\text{Si}_2\text{O}_5)(\text{OH})_4 \cdot 2\text{H}_2\text{O}$	light	colour change deliquescence	red: brown		King 1985; Nassau 1992 King 1985
15	Silicates of aluminum						
15.11	allophane	$(\text{Al}_2\text{O}_3)(\text{SiO}_2)_{1.3-2} \cdot 2.5-3\text{H}_2\text{O}$	moist air > ~25C, ambient humidity	hydration dehydration		to higher hydration states to lower hydration states	Howie 1984 Kitagawa 1974; Howie 1984
16	Silicates containing aluminum & other metals						
16.1.2	spodumene	$\text{LiAlSi}_2\text{O}_6$	light	colour change (var. kunzite, hiddenite)	fades rapidly		Nassau 1992
16.2.2	analcime	$\text{Na}(\text{AlSi}_2\text{O}_6) \cdot \text{H}_2\text{O}$	moist air	hydration dehydration		to higher hydration states to lower hydration states	Howie 1984 Howie 1984
16.2.5	natrolite	$\text{Na}_2\text{Al}_2\text{Si}_3\text{O}_{10} \cdot 2\text{H}_2\text{O}$	moist air	hydration dehydration		to higher hydration states to lower hydration states	Howie 1984 Howie 1984
16.2.7	paranatrolite	$\text{Na}_2\text{Al}_2\text{Si}_3\text{O}_{10} \cdot 3\text{H}_2\text{O}$		dehydration		to natrolite	Waller 1992
16.4.1	nepheline	$\text{Na}_3\text{K}(\text{Al}_4\text{Si}_4\text{O}_{16})$	light light moist air	colour change colour change hydration	colourless: pale pink pink: fades	to nepheline hydrate I	King 1985 Nassau 1992 Edgar 1965
16.6.1	beryl	$\text{Be}_3\text{Al}_2\text{Si}_6\text{O}_{18}$	light	colour change	maxixe: fades or pink; morganite: (paler) pink		Nassau 1992
16.7.9	montmorillonite	$(\text{Na}, \text{Ca})_{0.33}(\text{Al}, \text{Mg})_2(\text{Si}_4\text{O}_{10})(\text{OH})_2 \cdot n\text{H}_2\text{O}$	moist air	hydration dehydration		to higher hydration states to lower hydration states	Howie 1984 Howie 1984
16.9.23	laumontite	$\text{CaAl}_2\text{Si}_4\text{O}_{12} \cdot 4\text{H}_2\text{O}$	moist air	hydration dehydration	efflorescence; dull, crumbly, opaque	to leonhardite	Howie 1984 Waller 1992; King 1985
			light	surfacial reaction	opaque, powdery surface		O'Donoghue 1983; Nassau 1992
16.11.NI	chabazite	$(\text{Ca}, \text{K}_2, \text{Na}_2)_2[\text{Al}_2\text{Si}_4\text{O}_{12}]_2 \cdot 12\text{H}_2\text{O}$	moist air	hydration dehydration		to higher hydration states to lower hydration states	Howie 1984 Howie 1984
16.18.1	nontonite	$\text{Na}_{0.3}\text{Fe}_2((\text{Si}, \text{Al})_4\text{O}_{10})(\text{OH})_2 \cdot n\text{H}_2\text{O}$		dehydration		to lower hydration states	Howie 1984
16.19.22	vermiculite	$\text{Mg}_{0.7}(\text{Mg}, \text{Fe}, \text{Al})_6(\text{Si}, \text{Al})_8\text{O}_{20}(\text{OH})_4 \cdot 8\text{H}_2\text{O}$	moist air	hydration dehydration		to higher hydration states to lower hydration states	Howie 1984 Howie 1984
17	Silicates containing other anions						
17.2.1	topaz	$\text{Al}_2\text{SiO}_4(\text{F}, \text{OH})_2$	high light levels	colour change	brown/orange/sherry: (pale) blue or colourless; sherry: pink; blue: colourless		King 1985; Nassau 1992
17.2.NI	trilithionite	$\text{K}(\text{Li}, \text{Al})_3(\text{AlSi}_3\text{O}_{10})(\text{F}, \text{OH})_2$	light	colour change	purple: grey		Nassau 1992

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17.3.3	sodalite	$\text{Na}_4(\text{Si}_3\text{Al}_3)\text{O}_{12}\text{Cl}$	light	colour change	colourless/pale blue: pink; hackmanite (red): green, blue, or fades; hackmanite (pink): colourless		King 1985; Nassau 1992
17.4.2	spurrite	$\text{Ca}_5(\text{SiO}_4)_2(\text{CO}_3)$		efflorescence			Howie 1984; Bannister 1937
17.10.4	haüyne	$(\text{Na},\text{K})_3(\text{Ca},\text{Na})(\text{Al}_3\text{Si}_3\text{O}_{12})(\text{SO}_4,\text{S},\text{Cl})$	light	colour change	blue: fades		Nassau 1992
18	Niobates & Tantalates						
18.1.4	franconite	$\text{Na}(\text{Nb}_2\text{O}_5)(\text{OH}) \cdot 3\text{H}_2\text{O}$		dehydration		to lower hydration states	Waller 1992
18.1.11	microlite	$(\text{Ca},\text{Na})_2\text{Ta}_2(\text{O},\text{OH},\text{F})_7$	moist air	hydration		to higher hydration states	King 1985
18.2.1	fergusonite-(Y)	YNbO_4	moist air	hydration		to higher hydration states	King 1985
18.2.7	formanite-(Y)	YTao_4	moist air	hydration		to higher hydration states	King 1985
18.2.8	pyrochlore	$\text{A}_2\text{Nb}_2(\text{O},\text{OH})_6\text{Z}$	moist air	hydration		to higher hydration states	King 1985
18.3.3	betafite	$(\text{Ca},\text{Na},\text{U})_2(\text{Ti},\text{Nb},\text{Ta})_2\text{O}_6\text{Z}(\text{OH})$	moist air	hydration		to higher hydration states	King 1985
18.4.1	uranpyrochlore	$(\text{Ca},\text{U},\text{Ce})_2(\text{Nb},\text{Ti},\text{Ta})_2\text{O}_6(\text{O},\text{H},\text{F})$	moist air	hydration		to higher hydration states	King 1985
18.4.5	euxenite	$(\text{Y},\text{Ca},\text{Ce},\text{U},\text{Th})(\text{Nb},\text{Ta},\text{Ti})_2\text{O}_6$	moisture moist air	oxidation hydration	brown coating	to higher hydration states	O'Donoghue 1983 King 1985
18.4.8	polycrase-(Y)	$\text{Y}(\text{Ti},\text{Nb})_2(\text{O},\text{OH})_6$	moisture moist air	oxidation hydration	brown coating	to higher hydration states	O'Donoghue 1983 King 1985
18.4.12	samaraskite-(Y)	$\text{YFe}^{3+}\text{Nb}_2\text{O}_8$	moisture moist air	oxidation hydration	coating	to higher hydration states	O'Donoghue 1983 King 1985
19	Phosphates						
19.1.2	olympite	$\text{Na}_5\text{Li}(\text{PO}_4)_2$	moist air	hydration		to higher hydration states	Waller 1992
19.1.3	nahpoite	Na_2HPO_4	25C, > 28% RH	hydration		to dorfmanite	Waller 1992
19.1.4	dorfmanite	$\text{Na}_2\text{HPO}_4 \cdot 2\text{H}_2\text{O}$	25C, > 61% RH 25C, < 28% RH	hydration dehydrate		to higher hydration states to nahpoite	Waller 1992 Waller 1992
19.1.NI	schertelite	$(\text{NH}_4)_2\text{MgH}_2(\text{PO}_4)_2 \cdot 4\text{H}_2\text{O}$		deliquescence			Waller 1992
19.2.8	turquoise	$\text{CuAl}_6(\text{PO}_4)_4(\text{OH})_8 \cdot 4\text{H}_2\text{O}$	light, inc. sunlight	colour change dehydration	fades may show signs of shrinkage & colour loss	loss of capillary water	Nassau 1992; King 1985 King 1985
19.3.18	phosphorösslerite	$\text{Mg}(\text{HPO}_4) \cdot 7\text{H}_2\text{O}$		dehydration	efflorescence	to newberyite	Waller 1992; King 1985
19.3.23	struvite	$(\text{NH}_4)\text{Mg}(\text{PO}_4) \cdot 6\text{H}_2\text{O}$		dehydration	efflorescence; pale, dull	to dittmarite	Waller 1992; King 1985
19.4.1	monetite	$\text{Ca}(\text{HPO}_4)$	moist air	hydration		to brushite	Waller 1992
19.4.3	brushite	$\text{Ca}(\text{HPO}_4) \cdot 2\text{H}_2\text{O}$	20C, > 95% RH	deliquescence dehydration			Waller 1992 Waller 1992; King 1985
19.5.6	nabaphite	$\text{NaBaPO}_4 \cdot 9\text{H}_2\text{O}$		dehydration	efflorescence	to monetite	Waller 1992; King 1985
						to lower hydration states	Waller 1992

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19.5.8	nastrophite	$\text{Na}(\text{Sr,Ba})\text{PO}_4 \cdot 9\text{H}_2\text{O}$		dehydration		to lower hydration states	Waller 1992
19.7.5	variscite	$\text{AlPO}_4 \cdot 2\text{H}_2\text{O}$		dehydration		to lower hydration states	Howie 1984
19.8.2	brazilianite	$\text{NaAl}_3(\text{PO}_4)_2(\text{OH})_4$	light	colour change	green: fades		Nassau 1992
19.10.10	sincosite	$\text{Ca}(\text{VO})_2(\text{PO}_4)_2 \cdot 5\text{H}_2\text{O}$		dehydration		to lower hydration states	Waller 1992
19.11.10	torbernite	$\text{Cu}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 12\text{H}_2\text{O}$		dehydration		to metatorbernite	Waller 1992
19.11.14	meta-autunite	$\text{Ca}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 6-8\text{H}_2\text{O}$	20C, >38% RH	hydration		to autunite	Waller 1992
				dehydration	may show signs of shrinkage	to lower hydrates	Howie 1984; King 1985
19.11.15	autunite	$\text{Ca}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 11\text{H}_2\text{O}$	25C, <40% RH	dehydration		to meta-autunite	Waller 1992
19.11.22	uranocircite	$\text{Ba}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 10\text{H}_2\text{O}$		dehydration		to metauranocircite-I	Waller 1992
19.11.32	uranospathite	$(\text{Al},\square)(\text{UO}_2)_2(\text{PO}_4)_2\text{F} \cdot 20(\text{H}_2\text{O},\text{F})$		dehydration		to sabugalite	Waller 1992
19.11.40	lehnerite	$\text{Mn}^{2+}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 8\text{H}_2\text{O}$	> 25C	dehydration			Vochten 1990
19.11.41	bassetite	$\text{Fe}^{2+}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 10\text{H}_2\text{O}$	> 25C, ambient humidity oxygen	dehydration		to dihydrate	Vochten et al. 1984
				surficial oxidation		iron oxidation	Vochten et al. 1984
19.12.10	heterosite	$(\text{Fe}^{3+}, \text{Mn}^{3+})\text{PO}_4$	moisture	oxidation	tarnish: darkens		O'Donoghue 1983
19.12.11	purpurite	$\text{Mn}^{3+}\text{PO}_4$	light	surficial reaction	darkens		Nassau 1992
			moisture	oxidation	tarnish: darkens		O'Donoghue 1983
19.12.16	rockbridgeite	$\text{Fe}^{2+}\text{Fe}^{3+}_4(\text{PO}_4)_3(\text{OH})_5$	moisture	oxidation	colour change: brown		O'Donoghue 1983
19.12.22	hureaulite	$(\text{Mn,Fe})_5(\text{PO}_4)_2(\text{HPO}_4)_2 \cdot 4\text{H}_2\text{O}$	light	surficial photo-oxidation			Nassau 1992
19.12.35	ludlamite	$\text{Fe}^{2+}_3(\text{PO}_4)_2 \cdot 4\text{H}_2\text{O}$	oxygen	surficial oxidation	pale green: dark green or deep blue	extensive oxidation will eventually lead to structural collapse	Moore 1971
19.12.36	grafonite	$\text{Fe}^{2+}\text{Fe}^{2+}_2(\text{PO}_4)_2$	light	surficial photo-oxidation	darkens		Nassau 1992
19.12.40	messelite	$\text{Ca}_2\text{Fe}^{2+}(\text{PO}_4)_2 \cdot 2\text{H}_2\text{O}$	light	surficial photo-oxidation			Nassau 1992
19.13.2	strengite	$\text{FePO}_4 \cdot 2\text{H}_2\text{O}$	light	surficial reaction			Nassau 1992
19.13.11	vivianite	$\text{Fe}^{2+}_3(\text{PO}_4)_2 \cdot 8\text{H}_2\text{O}$	oxygen	surficial oxidation & partial dehydration	colourless: pale green/blue to deep blue	extensive oxidation (& subsequent dehydration) will eventually lead to structural collapse & disintegration	Moore 1971; McCammon & Burns 1980; Marinacea et al. 1997
19.13.19	phosphoferrite	$\text{Fe}^{2+}_3(\text{PO}_4)_2 \cdot 3\text{H}_2\text{O}$	oxygen	surficial oxidation	colourless/pale green: reddish-brown	to kryzhanovskite	Moore 1971
19.14.14	anapaite	$\text{Ca}_2\text{Fe}^{2+}(\text{PO}_4)_2 \cdot 4\text{H}_2\text{O}$	light	surficial photo-oxidation		to mitridatite	Nassau 1992; King 1985
19.14.28	koninckite	$\text{Fe}^{3+}\text{PO}_4 \cdot 3\text{H}_2\text{O}$	light	surficial photo-oxidation			Nassau 1992

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19.14.29	lazulite	$\text{MgAl}_2(\text{PO}_4)_2(\text{OH})_2$	acids				O'Donoghue 1983
20	Arsenates						
20.2.3	rösslerite	$\text{Mg}(\text{HAsO}_4) \cdot 7\text{H}_2\text{O}$		dehydration	efflorescence	to lower hydration states	King 1985
20.2.12	ferrarisite	$\text{Ca}_5(\text{AsO}_4)_2(\text{HAsO}_4)_2 \cdot 9\text{H}_2\text{O}$		dehydration		to lower hydration states	Waller 1992
20.2.14	phaunouxite	$\text{Ca}_3(\text{AsO}_4)_2 \cdot 11\text{H}_2\text{O}$		dehydration		to rauenthalite	Waller 1992
20.7.1	trögerite	$(\text{H}_3\text{O})(\text{UO}_2)(\text{AsO}_4) \cdot 3\text{H}_2\text{O}$	18-28C	structural phase change		to low symmetry, pseudo-tetragonal form	de Benyacar & da Abeledo 1974
20.7.6	metaováčekite	$\text{Mg}(\text{UO}_2)_2(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O}$	moist air	hydration		to nováčekite-II (10 H ₂ O)	Waller 1992
20.7.7	nováčekite-I	$\text{Mg}(\text{UO}_2)_2(\text{AsO}_4)_2 \cdot 12\text{H}_2\text{O}$		dehydration		to metanováčekite	Waller 1992
20.7.9	metauranospinite	$\text{Ca}(\text{UO}_2)_2(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O}$	high humidity	hydration		to uranospinite	Mrose 1953; Frondel 1958
20.7.10	uranospinite	$\text{Ca}(\text{UO}_2)(\text{AsO}_4)_2 \cdot 10\text{H}_2\text{O}$	RT, ambient humidity	dehydration		to metauranospinite	Mrose 1953; Frondel 1958
20.7.12	heinrichite	$\text{Ba}(\text{UO}_2)_2(\text{AsO}_4)_2 \cdot 10\text{H}_2\text{O}$		dehydration		to metaheinrichite	Waller 1992
20.7.14	arsenuranospathite	$\text{Al}(\text{UO}_2)_2(\text{AsO}_4)_2\text{F} \cdot 20\text{H}_2\text{O}$		dehydration		to lower hydration states	Waller 1992
20.9.6	symplesite	$\text{Fe}^{2+}_3(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O}$	light	surficial photo-oxidation	light green: indigo	loss of form	Nassau 1992; King 1985
				dehydration	efflorescence		Howie 1984
20.10.1	erythrite	$\text{Co}_3(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O}$	high light levels	surficial photo-oxidation			Nassau 1992
				dehydration	efflorescence		Howie 1984
20.10.9	annabergite	$\text{Ni}_3(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O}$		dehydration	efflorescence		Howie 1984
21	Vanadates						
21.2.2	rossite	$\text{Ca}(\text{VO}_3)_2 \cdot 4\text{H}_2\text{O}$		dehydration	efflorescence	to metarossite	Waller 1992; King 1985
21.2.3	metahewettite	$\text{CaV}_6\text{O}_{16} \cdot 3\text{H}_2\text{O}$	RT, ambient humidity	hydration		to hewettite	Hillebrand et al. 1914
21.2.4	hewettite	$\text{CaV}_6\text{O}_{16} \cdot 9\text{H}_2\text{O}$	> 25C, dry air	dehydration	colour change: darkens	to metahewettite	Hillebrand et al. 1914
21.2.5	pintadoite	$\text{Ca}_2(\text{V}_2\text{O}_7) \cdot 9\text{H}_2\text{O}$		dehydration	efflorescence	to lower hydration states	King 1985
21.2.8	pascoite	$\text{Ca}_3(\text{V}_{10}\text{O}_{28}) \cdot 17\text{H}_2\text{O}$	RT, ~0% RH	dehydration			Hillebrand et al. 1914
21.2.9	sherwoodite	$\text{Ca}_{4.5}(\text{AlV}^{4+}_2\text{V}^{5+}_{12}\text{O}_{40}) \cdot 28\text{H}_2\text{O}$		alteration	colour change: yellow-green		O'Donoghue 1983
21.2.10	melanovanadite	$\text{Ca}(\text{V}^{5+}, \text{V}^{4+})_4\text{O}_{10} \cdot 5\text{H}_2\text{O}$	moist air	hydration		to pascoite	King 1985
21.2.13	metadelrioite	$\text{CaSr}(\text{V}_2\text{O}_6)(\text{OH})_2$	moist air	hydration		to delrioite	Waller 1992
21.2.14	delrioite	$\text{CaSr}(\text{V}_2\text{O}_6)(\text{OH})_2 \cdot 3\text{H}_2\text{O}$		dehydration		to metadelrioite	Waller 1992
21.3.3	metaschoderite	$\text{Al}_2(\text{PO}_4)(\text{VO}_4) \cdot 6\text{H}_2\text{O}$	moist air	hydration		to schoderite	Waller 1992
21.3.4	schoderite	$\text{Al}_2(\text{PO}_4)(\text{VO}_4) \cdot 8\text{H}_2\text{O}$		dehydration		to metaschoderite	Waller 1992
21.4.4	carnotite	$\text{K}_2(\text{UO}_2)_2(\text{VO}_4)_2 \cdot 3\text{H}_2\text{O}$		dehydration		to lower hydration states	Waller 1992

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21.4.7	metatyuyamunite	$\text{Ca}(\text{UO}_2)_2(\text{VO}_4)_2 \cdot 3\text{-}5\text{H}_2\text{O}$	light 24C, > 55% RH	colour change hydration	yellow: green	to tyuyamunite	Nassau 1992 Waller 1992
21.4.8	tyuyamunite	$\text{Ca}(\text{UO}_2)_2(\text{VO}_4)_2 \cdot 5\text{-}8\text{H}_2\text{O}$	high light levels 24C, <55% RH	colour change dehydration	bright yellow: dirty green-yellow efflorescence	to metatyuyamunite	King 1985 Waller 1992
21.4.9	rauvite	$\text{Ca}(\text{UO}_2)_2(\text{V}_{10}\text{O}_{28}) \cdot 16\text{H}_2\text{O}$		dehydration		to lower hydration states	Waller 1992
21.4.11	metavanuralite	$\text{Al}(\text{UO}_2)_2(\text{VO}_4)_2(\text{OH}) \cdot 8\text{H}_2\text{O}$	20C, > 47% RH	hydration dehydration		to vanuralite	Waller 1992
21.4.12	vanuralite	$\text{Al}(\text{UO}_2)_2(\text{V}_2\text{O}_8)(\text{OH}) \cdot 8.5\text{-}11\text{H}_2\text{O}$	moist air 20C, < 47% RH	hydration dehydration	efflorescence	to lower hydration states to higher hydration states to metavanuralite	King 1985; Howie 1984 Howie 1984 Waller 1992
22	Phosphates, Arsenates, & Vanadates with other anions						
22.1.23	triplite	$\text{Mn}^{2+}_2(\text{PO}_4)\text{F}$	moist air	hydration		to higher hydration states	King 1985
22.2.14	vanadinite	$\text{Pb}_5(\text{VO}_4)_3\text{Cl}$	prolonged light exposure	colour change	bright yellow: dull brown; red: brown		Nassau 1992; King 1985
22.2.NI	apatite	$\text{Ca}_5(\text{PO}_4)_3(\text{Cl}/\text{F}/\text{OH})$	light	colour change	fades		Nassau 1992; King 1985
22.3.21	sasaite	$(\text{Al}, \text{Fe}^{3+})_{14}(\text{PO}_4)_{11}(\text{SO}_4)(\text{OH})_7 \cdot 83\text{H}_2\text{O}$		dehydration		to lower hydration states	Waller 1992
25	Sulfates						
25.1.1	metathénardite	Na_2SO_4	> 32.4C 25C, > 81% RH	deliquescence hydration		to saturated solution of metathénardite to mirabilite	Waller 1992 Waller 1992
25.1.2	matteuccite	$\text{NaHSO}_4 \cdot \text{H}_2\text{O}$	20C, > 52% RH 20C, < 16% RH	deliquescence dehydration		to thénardite	Waller 1992 Waller 1992
25.1.3	mirabilite	$\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$	RT, > 93% RH 25C, < 81% RH	deliquescence dehydration	efflorescence	to lower hydration states, eventually to thénardite	Waller 1992
25.1.5	mercallite	KHSO_4	20C, > 86% RH	deliquescence			Waller 1992
25.1.8	mascagnite	$(\text{NH}_4)_2\text{SO}_4$	20C, > 81% RH	deliquescence			Waller 1992
25.2.1	chalcocyanite	CuSO_4	25C, > 3% RH	hydration		to bonattite	Waller 1992
25.2.2	bonattite	$\text{CuSO}_4 \cdot 3\text{H}_2\text{O}$	15C, > ~27% RH; 20C, > 30% RH; 25C, > 33% RH 25C, < 22% RH	hydration dehydration		to chalcantite	Chou et al. 2002
25.2.3	chalcantite	$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$	25C, > 97% RH 15C, < ~27% RH; 20C, < 30% RH; 25C, < 33% RH	deliquescence dehydration	efflorescence	to chalcocyanite to bonattite	Chou et al. 2002 Chou et al. 2002
25.2.4	boothite	$\text{CuSO}_4 \cdot 7\text{H}_2\text{O}$		dehydration	efflorescence	to chalcantite	Waller 1992; Hazen & Ausubel 2016
25.2.11	kröhnkite	$\text{Na}_2\text{Cu}(\text{SO}_4)_2 \cdot 2\text{H}_2\text{O}$	high RH	colour change	translucent sky blue: green or opaque		King 1985

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25.2.13	piypite	$K_4Cu_4O_2(SO_4)_4 \cdot (Na,Cu)Cl$		deliquescence			Waller 1992
25.2.15	cyanochroite	$K_2Cu(SO_4)_2 \cdot 6H_2O$		deliquescence			Howie 1984
			25C, < 36% RH	dehydration	efflorescence	to lower hydration states	Waller 1992
25.2.27	poitevinite	$(Cu,Fe)SO_4 \cdot H_2O$	25C, > 22% RH	hydration		to higher hydration states	Waller 1992
			25C, < 22% RH	dehydration	efflorescence	to anhydrous phase	Waller 1992
25.3.1	kieserite	$MgSO_4 \cdot H_2O$	15C, > ~41% RH; 20C, > ~43% RH; 25C, > ~45% RH	hydration		to hexahydrite	Chou & Seal 2003, 2007; Wang et al. 2009
			25C, < 4% RH	dehydration	efflorescence	to anhydrous phase	Waller 1992
25.3.2	sanderite	$MgSO_4 \cdot 2H_2O$	31C, > 22% RH	hydration		to higher hydration states	Waller 1992
			31C, < 8% RH	dehydration	efflorescence	to lower hydration states	Waller 1992
25.3.3	starkeyite	$MgSO_4 \cdot 4H_2O$	15C, > 35% RH; 20C, > ~37% RH; 25C, > ~39% RH	hydration		to hexahydrite	Chou & Seal 2003, 2007
			20C, < 19% RH	dehydration	efflorescence	to lower hydration states	Wang et al. 2009
25.3.4	pentahydrite	$MgSO_4 \cdot 5H_2O$	25C, > 41% RH	hydration		to higher hydration states	Waller 1992
			25C, < 37% RH	dehydration	efflorescence	to lower hydration states	Waller 1992
25.3.5	hexahydrite	$MgSO_4 \cdot 6H_2O$	15C, > ~43% RH ; 20C, > ~48% RH; 25C, > ~53% RH	hydration		to epsomite	Chou & Seal 2003, 2007
			15C, < ~41% RH; 20C, < ~43% RH; 25C, < ~45% RH	dehydration		to kieserite	Chou & Seal 2003, 2007; Wang et al. 2009
			15C, < 35% RH; 20C, < ~37% RH; 25C, < ~39% RH	dehydration	efflorescence	to starkeyite	Chou & Seal 2003, 2007
25.3.6	epsomite	$MgSO_4 \cdot 7H_2O$	15C, > ~94% RH; 20C, > ~93% RH; 25C, > ~91% RH	deliquescence		to saturated solution of hexahydrite	Chou & Seal 2003, 2007
			15C, < ~43% RH ; 20C, < ~48% RH; 25C, < ~53% RH	dehydration	efflorescence	to hexahydrite	Chou & Seal 2003, 2007
25.3.10	konyaite	$Na_2Mg(SO_4)_2 \cdot 5H_2O$		dehydration		to blöditite	Waller 1992
25.3.11	löweite	$Na_{12}Mg_7(SO_4)_{13} \cdot 15H_2O$		deliquescence			Waller 1992
			high RH	efflorescence	dull, white		King 1982
25.3.13	langbeinite	$K_2Mg_2(SO_4)_3$		deliquescence			Waller 1992; King 1985
25.3.14	leonite	$K_2Mg(SO_4)_2 \cdot 4H_2O$	moist air	hydration		to picromerite	Waller 1992; King 1985
25.3.15	picromerite	$K_2Mg(SO_4)_2 \cdot 6H_2O$		deliquescence			Howie 1984
				dehydration	efflorescence	to lower hydration states	King 1985; Howie 1984
25.3.17	boussingaultite	$(NH_4)_2Mg(SO_4)_2 \cdot 6H_2O$	25C, > 96% RH	deliquescence			Waller 1992
			25C, < 20% RH	dehydration	efflorescence	to lower hydration states	Waller 1992

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25.3.NI.1	alpersite	(Mg,Cu)(SO ₄) · 7H ₂ O	21C, < 46% RH	dehydration	efflorescence	to copper-bearing pentahydrate	Peterson et al. 2006
25.3.NI.2	meridianite	MgSO ₄ · 11H ₂ O	> 2C	melts		epsomite & water	Hazen & Ausubel 2016
25.4.1	anhydrite	CaSO ₄	light 15C, > ~66% RH; 20C, > 70% RH; 25C, > 75% RH	colour change hydration	blue: white	to gypsum	Nassau 1992; King 1985 Chou & Seal 2007
25.4.3	gypsum	CaSO ₄ · 2H ₂ O	light 15C, < ~66% RH; 20C, < 70% RH; 25C, < 75% RH	colour change dehydration	pink: fades	to anhydrite	Nassau 1992 Chou & Seal 2007
25.4.13	ettringite	Ca ₆ Al ₂ (SO ₄) ₃ (OH) ₁₂ · 26H ₂ O		dehydration	efflorescence	to lower hydration states	Waller 1992; Howie 1984
25.4.15	celestine	SrSO ₄	high light levels	colour change	blue: white		Nassau 1992; King 1985
25.4.17	baryte	BaSO ₄	high light levels	colour change	colourless, pale blue: chalky white or darker blue; blue: colourless; yellow/brown: green or blue		Nassau 1992; King 1985
25.5.2	goslarite	ZnSO ₄ · 7H ₂ O	> 37.9C 15C, > 90% RH; 20C, > ~92% RH; 25C, > ~89% RH 15C, < ~54% RH; 20C, < ~59% RH; 25C, < ~64% RH	deliquescence deliquescence dehydration		to saturated solution of zinkosite to aqueous solution to bianchite	Waller 1992 Chou & Seal 2005b Chou & Seal 2005b
25.5.6	boyleite	(Zn,Mg)SO ₄ · 4H ₂ O		dehydration		to lower hydration states	Waller 1992
25.5.11	gunningite	ZnSO ₄ · H ₂ O	15C, > ~52% RH; 20C, > ~55% RH; 25C, > ~58% RH 25C, < 4% RH	hydration dehydration		to bianchite to zinkosite	Chou & Seal 2005b Waller 1992
25.5.13	bianchite	(Zn,Fe)SO ₄ · 6H ₂ O	> 48.8C 15C, > ~54% RH; 20C, > ~59% RH; 25C, > ~64% RH 15C, < ~52% RH; 20C, < ~55% RH; 25C, < ~58% RH	deliquescence hydration dehydration		to saturated solution of <i>chaidamuite</i> to goslarite to gunningite	Waller 1992 Chou & Seal 2005b Chou & Seal 2005b
25.5.15	zincmelanterite	(Zn,Cu,Fe)SO ₄ · 7H ₂ O	25C, > 89% RH 25C, < 65% RH	deliquescence dehydration		to bianchite	Waller 1992 Waller 1992

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25.6.3	mendozite	$\text{NaAl}(\text{SO}_4)_2 \cdot 11\text{H}_2\text{O}$		dehydration	efflorescence	to tamarugite	Waller 1992; King 1985
25.6.5	aluminite	$\text{Al}_2(\text{SO}_4)(\text{OH})_4 \cdot 7\text{H}_2\text{O}$		dehydration		to meta-aluminite	Waller 1992
25.6.7	alunogen	$\text{Al}_2(\text{SO}_4)_3 \cdot 17\text{H}_2\text{O}$	20C, > 88% RH	deliquescence dehydration		to meta-alunogen	Waller 1992 Waller 1992
25.6.8	alunite	$\text{KAl}_3(\text{SO}_4)_2(\text{OH})_6$		dehydration	may show signs of shrinkage	to lower hydration states	King 1985
25.6.10	hydrobasaluminite	$\text{Al}_4(\text{SO}_4)(\text{OH})_{10} \cdot 12\text{-}36\text{H}_2\text{O}$	RT	dehydration		to felsöbányaite	Waller 1992; Bigham & Nordstrom 2000
25.6.11	zaherite	$\text{Al}_{12}(\text{SO}_4)_5(\text{OH})_{26} \cdot 20\text{H}_2\text{O}$		dehydration		to lower hydration states	Waller 1992; Hazen & Ausubel 2016
25.6.12	tamarugite	$\text{NaAl}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$	20C, > 86% RH	hydration		to alum-(Na)	Waller 1992
25.6.14	alum-(Na)	$\text{NaAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$	20C, < 86% RH	dehydration		to tamarugite	Waller 1992
25.6.17	alum-(K)	$\text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$	20C, < 9% RH	dehydration		to kalinite	Waller 1992
25.6.20	tschermigite	$(\text{NH}_4)\text{Al}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$	25C, < 7% RH	dehydration	efflorescence	to godovikovite	Waller 1992
25.7.3	anglesite	PbSO_4	light water & carbon dioxide; carbonic acid	colour change carbonation	yellow, brown: fades	to cerussite & sulfuric acid	Nassau 1992; King 1985 Swartzlow 1933
25.8.5	minasragrite	$(\text{V}^{4+}\text{O})(\text{SO}_4) \cdot 5\text{H}_2\text{O}$		deliquescence dehydration	efflorescence		King 1985 King 1985
25.8.9	uranopilite	$(\text{UO}_2)_6(\text{SO}_4)\text{O}_2(\text{OH})_6 \cdot 12\text{-}14\text{H}_2\text{O}$		dehydration		to meta-uranopilite	Waller 1992
25.9.1	szmikite	$\text{MnSO}_4 \cdot \text{H}_2\text{O}$	25C, > 83% RH 20C, < 17% RH	hydration dehydration	efflorescence	to higher hydration states to anhydrous state	Waller 1992 Waller 1992
25.9.2	jökokuite	$\text{MnSO}_4 \cdot 5\text{H}_2\text{O}$	15C, > 84% RH 20C, < 74% RH	hydration dehydration	efflorescence	to higher hydration states to lower hydration states	Waller 1992 Waller 1992
25.9.3	mallardite	$\text{MnSO}_4 \cdot 7\text{H}_2\text{O}$	< 9C	dehydration		to saturated solution of jökokuite	Waller 1992
25.9.5	chvaeticite	$(\text{Mn,Mg})\text{SO}_4 \cdot 6\text{H}_2\text{O}$		dehydration		to lower hydration states	Hazen & Ausubel 2016
25.9.9	ilesite	$(\text{Mn,Zn,Fe})\text{SO}_4 \cdot 4\text{H}_2\text{O}$	ambient conditions; <70% RH	dehydration		to szmikite	Helda & Bohatý 2002; Anderson et al. 2012
25.10.1	szomolnokite	$\text{FeSO}_4 \cdot \text{H}_2\text{O}$	20C, > 57% RH 20C, < 11% RH	hydration dehydration	efflorescence	to rozenite to anhydrous state	Waller 1992 Waller 1992
25.10.2	rozenite	$\text{FeSO}_4 \cdot 4\text{H}_2\text{O}$	RT, < 62% RH 15C, > ~54% RH; 20C, > ~57% RH; 25C, > ~62% RH	oxidation hydration dehydration		to copiapite to melanterite to szomolnokite	Jerz & Rimstidt 2003 Jambor et al. 2000; Chou et al. 2002 Chou et al. 2002
25.10.3	siderotil	$\text{FeSO}_4 \cdot 5\text{H}_2\text{O}$	22C, >35% RH 22C, <5% RH	hydration dehydration		to melanterite to szomolnokite	Eckel 1933; Ehlers & Stiles 1965 Eckel 1933

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25.10.4	ferrohexahydrate	$\text{FeSO}_4 \cdot 6\text{H}_2\text{O}$		dehydration		to lower hydration states	Jambor et al. 2000; Hyde et al. 2011 (Fig. 1)
			25C, > 61% RH	hydration		to higher hydration states	Waller 1992
25.10.5	melanterite	$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$	RT, 63-89% RH	oxidation		to fibroferrite (via copiapite at 63-80% RH)	Jerz & Rimstidt 2003
			15-25C, > 95% RH	deliquescence			Chou et al. 2002; Jerz & Rimstidt 2003
			15C, < ~54% RH; 20C, < ~57% RH; 25C, < ~62% RH	dehydration	efflorescence	to rozenite, when Cu < 1.2%	Chou et al. 2002
			20C, < 65% RH	dehydration	efflorescence	to siderotil, when Cu > 1.2%	Jambor et al. 2000
25.10.6	rhomboclase	$(\text{H}_5\text{O}_2)\text{Fe}^{3+}(\text{SO}_4)_2 \cdot 2\text{H}_2\text{O}$	25C, > 58-60% RH	deliquescence		to solution $\text{Fe}_2(\text{SO})_3$	Xu et al. 2009
			22-25C, 33-53% RH	alteration		to kornelite & paracoquimbite	Xu et al. 2009
				dehydration	efflorescence	to lower hydration states	Waller 1992; King 1985
25.10.8	kornelite	$\text{Fe}^{3+}_2(\text{SO}_4)_3 \cdot 7\text{H}_2\text{O}$	22-25C, > 43% RH	hydration		to paracoquimbite	Xu et al. 2009
			20C, >73% RH	deliquescence			Waller 1992
25.10.9	coquimbite	$\text{AlFe}^{3+}_3(\text{SO}_4)_6(\text{H}_2\text{O})_{12} \cdot 6\text{H}_2\text{O}$	moist air	hydration		to higher hydration states	Howie 1984
				dehydration	efflorescence	to lower hydration states	Waller 1992; King 1985
25.10.10	paracoquimbite	$\text{Fe}^{3+}_4(\text{SO}_4)_6(\text{H}_2\text{O})_{12} \cdot 6\text{H}_2\text{O}$	22-25C, < 43% RH	dehydration		to kornelite	Xu et al. 2009
25.10.11	quenstedtite	$\text{Fe}^{3+}_2(\text{SO}_4)_3 \cdot 11\text{H}_2\text{O}$		dehydration	efflorescence	to paracoquimbite	Waller 1992; King 1985
25.10.12	ferricopiapite	$\text{Fe}^{3+}_{0.67}\text{Fe}^{3+}_4(\text{SO}_4)_6(\text{OH})_2 \cdot 20\text{H}_2\text{O}$	22-25C, > 68% RH	deliquescence		to solution $\text{Fe}_2(\text{SO})_3$	Xu et al. 2009
			22-25C, 33-53% RH	alteration		to kornelite & paracoquimbite	Xu et al. 2009
25.10.14	hohmannite	$\text{Fe}^{3+}_2(\text{SO}_4)_2\text{O} \cdot 8\text{H}_2\text{O}$		dehydration		to metahohmannite	Hazen & Ausubel 2016
25.10.18	fibroferrite	$\text{Fe}^{3+}(\text{SO}_4)(\text{OH}) \cdot 5\text{H}_2\text{O}$	RT, > 88% RH	deliquescence			Jerz & Rimstidt 2003
25.10.22	copiapite	$\text{Fe}^{2+}\text{Fe}^{3+}_4(\text{SO}_4)_6(\text{OH})_2 \cdot 20\text{H}_2\text{O}$	RT, ~63-78% RH	oxidation		to fibroferrite	Jerz & Rimstidt 2003
			RT, > 80% RH	deliquescence			Jerz & Rimstidt 2003
			light	photo-sensitive		gradual decomposition	Hazen & Ausubel 2016
25.11.2	sideronatrite	$\text{Na}_2\text{Fe}(\text{SO}_4)_2(\text{OH}) \cdot 3\text{H}_2\text{O}$	>~35C	dehydration		to metasideronatrite	Ventruti et al. 2010; Ventruti et al. 2013
25.11.3	metasideronatrite	$\text{Na}_2\text{Fe}(\text{SO}_4)_2(\text{OH}) \cdot \text{H}_2\text{O}$	light <40C, >60%RH	photo-sensitive hydration		gradual decomposition to sideronatrite	Hazen & Ausubel 2016 Ventruti et al. 2010
25.11.4	ferrinatrite	$\text{Na}_3\text{Fe}(\text{SO}_4)_3 \cdot 3\text{H}_2\text{O}$		deliquescence			Waller 1992
			moist air	hydration			King 1985
25.11.10	voltaite	$\text{K}_2\text{Fe}^{2+}_5\text{Fe}^{3+}_3\text{Al}(\text{SO}_4)_{12} \cdot 18\text{H}_2\text{O}$	moist air	hydration			King 1985
				dehydration	efflorescence	to lower hydration states	Waller 1992; King 1985
25.11.17	loncreekite	$(\text{NH}_4)\text{Fe}^{3+}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$		dehydration		to sabieite	Hazen & Ausubel 2016
25.11.26	halotrichite	$\text{Fe}^{2+}\text{Al}_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	RT, > 90% RH	deliquescence			Jerz & Rimstidt 2003
				dehydration	efflorescence	to lower hydration states	Waller 1992; King 1985
25.11.NI	mikasaite	$\text{Fe}^{3+}_2(\text{SO}_4)_3$		deliquescence			Hazen & Ausubel 2016

Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
25.12.1	bieberite	CoSO ₄ · 7H ₂ O	15-25C, > ~95% RH 15C, < 60% RH; 20C, < ~65% RH; 25C, < 70% RH	deliquescence dehydration	efflorescence	to aqueous solution to moorhouseite	Chou & Seal 2005a Chou & Seal 2005a
25.12.3	aplowite	(Co,Mn,Ni)SO ₄ · 4H ₂ O		dehydration		to lower hydration states	Hazen & Ausubel 2016
25.12.4	moorhouseite	(Co,Ni,Mn)SO ₄ · 6H ₂ O	15C, > 60% RH; 20C, > ~65% RH; 25C, > 70% RH	hydration		to bieberite	Chou & Seal 2005a
25.12.5	retgersite	NiSO ₄ · 6H ₂ O	15C, > ~79% RH; 20C, > ~82%RH; 25C, > 87% RH	hydration		to morenosite	Wang et al. 2017
25.12.6	morenosite	NiSO ₄ · 7H ₂ O	15-25C, > ~95% RH 15C, < ~79% RH; 20C, < ~82%RH; 25C, < 87% RH	deliquescence dehydration	efflorescence	to retgersite	Wang et al. 2017 Wang et al. 2017
25.12.13	nicklehexahydrite	(Ni,Mg,Fe)SO ₄ · 6H ₂ O	moist air	hydration dehydration	efflorescence	to higher hydration states	Howie 1984 Howie 1984
26	Sulfates with halide						
26.14	kainite	KMg(SO ₄)Cl · 3H ₂ O	32C, < 3% RH	deliquescence dehydration		to anhydrokainite	Waller 1992; King 1985 Waller 1992
26.15	mosesite	(Hg ₂ N)(Cl, SO ₄ , MO ₄) · H ₂ O	light	colour change	yellow: green		Nassau 1992
26.16	kleinite	(Hg ₂ N)(Cl, SO ₄) · nH ₂ O	high light levels light	colour change surficial photo-oxidation	yellow: orange darkens		Nassau 1992; King 1985 Nassau 1992
27	Sulfites, Chromates, Molybdates, & Tungstates						
27.2.1	tarapacáite	K ₂ (CrO ₄)	20C, > 88% RH	deliquescence			Waller 1992
27.2.5	crocoite	PbCrO ₄	light	surficial photo-oxidation	darkens		Nassau 1992
27.2.6	phoenicochroite	Pb ₂ (CrO ₄)O	light	surficial reaction			Nassau 1992
27.2.NI	edoylerite	Hg ²⁺ ₃ (CrO ₄)S ₂	light	photo-sensitive	colour change: olive-green	gradual decomposition	Hazen & Ausubel 2016
27.3.14	ferrimolybdite	Fe ³⁺ ₂ (Mo ⁶⁺ O ₄) ₃ · 7H ₂ O	warm, dry air	dehydration	colour variation	prolonged dehydration leads to structural collapse	Golding et al. 1960
28	Selenites, Selenates, Tellurites, & Tellurates						
28.1.8	marthozite	Cu ²⁺ (UO ₂) ₃ (SeO ₃) ₂ O ₂ · 8H ₂ O		dehydration		to lower hydration states	Waller 1992; Hazen & Ausubel 2016
29	Iodates						
29.5	dietzeite	Ca ₂ (IO ₃) ₂ (CrO ₄) · H ₂ O	light	surficial photo-oxidation			Nassau 1992

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Hey Number	Mineral Name	Chemical Formula	Conditions	Response	Appearance	Alterations	References
30	Thiocyanates						
30.2	kafhydrocyanite	$K_4[Fe(CN)_6] \cdot H_2O$	20C, < 42% RH	dehydration		to lower hydration states	Waller 1992
32	Crystalline Organics						
32.11	acetamide	CH_3CONH_2	air & sunlight	volatizes deliquescence			Hazen & Ausubel 2016 Waller 1992
32.13	urea	$CO(NH_2)_2$	25C, > 77% RH	deliquescence			Waller 1992
O	'Precious' Non-minerals						
O.1	amber			dehydration	may show signs of shrinkage	volumetric change	King 1985; Howie 1984
			acids, biocides, & cleaning agents (see reference for full list)	absorbtion then desorption	may show signs of swelling/ shrinkage; possible darkening, softening, or efflorescence, depending on compound		Waddington & Fenn 1988
O.2	pearl	$CaCO_3$		dehydration			O'Donoghue 1983
O.3	obsidian			hydration	black: white; may show signs of expansion	to perlite; extensive hydration can lead to the formation of clay &/or zeolite minerals	Jezek & Noble 1978
O.4	coral	$CaCO_3$	light	colour change	red: fade		Schumann 1974

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Hey Number	Mineral Name	Chemical Formula	Water	Acids	Alterations	References
1	Elements & Alloys					
1.1	copper	Cu		easily soluble in nitric acid		O'Donoghue 1983; Schumann 1992
1.2	silver	Ag		soluble in nitric acid, hot sulfuric acid, & aqua regia		O'Donoghue 1983
1.48	dyscrasite	Ag ₃ Sb		soluble in nitric acid		O'Donoghue 1983; Schumann 1992
1.49	bismuth	Bi		soluble in nitric acid		Schumann 1992
1.54	tellurium	Te		soluble in concentrated sulfuric acid		Schumann 1992
1.57	iron	Fe		soluble in acids, especially nitric		O'Donoghue 1983
3	Sulfides, Selenides, Tellurides, Arsenides and Bismuthides					
3.1.1	chalcocite	Cu ₂ S		soluble in nitric acid		Schumann 1992
3.1.8	covellite	CuS		soluble in nitric acid		Schumann 1992
3.1.23	bornite	Cu ₅ FeS ₄		soluble in nitric acid		Schumann 1992
3.1.25	chalcopyrite	CuFeS ₂		soluble in nitric acid		Schumann 1992
3.2.1	acanthite	Ag ₂ S		soluble in acid		Schumann 1992
3.3.9	sylvanite	(Au,Ag) ₂ Te ₄		soluble in nitric acid & aqua regia		O'Donoghue 1983; Schumann 1992
3.4.4	sphalerite	ZnS		decomposes in acids	to hydrated sulphates	Howie 1992
3.4.5	wurtzite	(Zn,Fe)S		redily souble in mineral acids		King 1983
3.4.7	greenockite	CdS	soluble			King 1992
3.6.5	galena	PbS		soluble in nitric acid		Schumann 1992
3.7.8	orpiment	As ₂ S ₃	slightly soluble in hot water			King 1992
3.7.14	stibnite	Sb ₂ S ₃		soluble in most acids, inc. nitric acid & aqua regia		O'Donoghue 1983; Schumann 1992
3.8.6	molybdenite	MoS ₂		soluble in weak acids		King 1983
3.9.3	pyrite	FeS ₂		soluble in nitric acid & aqua regia		Schumann 1992
3.9.11	löllingite	FeAs ₂		soluble in nitric acid		Schumann 1992
3.9.12	arsenopyrite	FeAsS		soluble in nitric acid		Schumann 1992

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Hey Number	Mineral Name	Chemical Formula	Water	Acids	Alterations	References
3.10.2	linnaeite	$\text{Co}^{2+}\text{Co}^{3+}_2\text{S}_4$		soluble in nitric acid		Schumann 1992
3.10.12	skutterudite	CoAs_3		soluble in nitric acid		Schumann 1992
3.11.3	millerite	NiS	decomposes in hot water			King 1992
				soluble in nitric acid & aqua regia		Schumann 1992
3.11.8	violarite	$\text{Fe}^{2+}\text{Ni}^{3+}_2\text{S}_4$	decomposes in hot water			King 1982
3.11.9	pentlandite	$(\text{Fe},\text{Ni})_9\text{S}_8$		soluble in nitric acid		Schumann 1992
3.11.26	maucherite	$\text{Ni}_{11}\text{As}_8$		soluble in concentrated nitric acid		Schumann 1992
3.11.27	nickeline	NiAs		soluble in nitric acid	with green colour	Schumann 1992
3.11.30	nickelskutterudite	$(\text{Ni},\text{Co},\text{Fe})\text{As}_3$		soluble in nitric acid		Schumann 1992
3.11.34	gersdorffite	NiAsS		soluble in nitric acid		O'Donoghue 1983; Schumann 1992
3.11.36	breithauptite	NiSb		soluble in nitric acid & aqua regia		Schumann 1992
5	Sulfosalts: Sulfarsenites & Sulfobismuthites					
5.1.7	emlectite	CuBiS_2		soluble in acid		Schumann 1992
5.2.3	proustite	Ag_3AsS_3		soluble in nitric acid		Schumann 1992
5.2.9	pyrargyrite	Ag_3SbS_3		soluble in nitric acid		Schumann 1992
5.2.11	stephanite	Ag_5SbS_4		soluble in nitric acid		Schumann 1992
5.3.1	nagyágite	$[\text{Pb}_3(\text{Pb},\text{Sb})_3\text{S}_6](\text{Au},\text{Te})_3$		soluble in nitric acid & aqua regia		Schumann 1992
5.6.15	boulangerite	$\text{Pb}_5\text{Sb}_4\text{S}_{11}$		soluble in warm nitric acid		Schumann 1992
5.7.2	bournonite	PbCuSbS_3		soluble in nitric acid		Schumann 1992
5.7.3	meneghinite	$\text{Pb}_{13}\text{CuSb}_7\text{S}_{24}$		soluble in nitric acid		Schumann 1992
6	Sulfosalts					
6.1.6	stannite	$\text{Cu}_2\text{FeSnS}_4$		soluble in nitric acid		Schumann 1992
6.2.1	germanite	$\text{Cu}_{13}\text{Fe}_2\text{Ge}_2\text{S}_{16}$		soluble in nitric acid		Schumann 1992
6.3.1	enargite	Cu_3AsS_4		soluble in nitric acid & aqua regia		Schumann 1992

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Hey Number	Mineral Name	Chemical Formula	Water	Acids	Alterations	References
7	Oxides & Hydroxides					
7.3.2	tenorite	CuO		easily soluble in acids		Schumann 1992
7.4.7	periclase	MgO		soluble in acid (powder: easily; crystal: difficult)		Schumann 1992
7.4.8	brucite	Mg(OH) ₂		easily soluble in acids		Schumann 1992
7.5.1	zincite	ZnO		soluble in acids		Schumann 1992
7.8.1	quartz	SiO ₂	6ppm (@ 25C, 1 atm)			Fournier & Rowe 1962
7.8.5	cristobalite	SiO ₂	27ppm (@ 25C, 1 atm)		solution of silica	Fournier & Rowe 1962
7.10.10	thorianite	ThO ₂		soluble in nitric acid		Schumann 1992
7.13.1	arsenolite	As ₂ O ₃	soluble in hot/warm water			King 1992; Schumann 1992
7.13.4	valentinite	Sb ₂ O ₃		easily soluble in acids		Schumann 1992
7.16.1	uraninite	UO ₂		easily soluble in sulfuric acid & warm nitric acid		Schumann 1992
8	Halides					
8.1.2	villiaumite	NaF	highly soluble			King 1992
8.1.3	halite	NaCl	highly soluble			O'Donoghue 1983; King 1992; Schumann 1992
8.1.6	sylvite	KCl	highly soluble	readily soluble in acids		O'Donoghue 1983; King 1992; Schumann 1992
8.1.7	salammoniac	NH ₄ Cl	highly soluble			King 1992; Schumann 1992
8.2.4	atacamite	Cu ₂ (OH) ₃ Cl		readily soluble in ammonia		Schumann 1992
8.2.12	mitscherlichite	K ₂ CuCl ₄ · 2H ₂ O	highly soluble			King 1992
8.3.1	chlorargyrite	AgCl		soluble in ammonia		Schumann 1992
8.4.3	bischofite	MgCl ₂ · 6H ₂ O	highly soluble			King 1992
8.4.5	carnallite	KMgCl ₃ · 6H ₂ O	highly soluble		with crackling sound	King 1992; Schumann 1992
8.4.6	tachyhydrite	CaMg ₂ Cl ₆ · 12H ₂ O	highly soluble			King 1992
8.6.5	cryolite	Na ₂ NaAlF ₆	slightly soluble			King 1992
				easily soluble in sulfuric acid		Schumann 1992

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Hey Number	Mineral Name	Chemical Formula	Water	Acids	Alterations	References
8.6.10	koenenite	$\text{Na}_4\text{Mg}_9\text{Al}_4\text{Cl}_{12}(\text{OH})_{22}$	decomposes in water		to scaly precipitate	King 1992
8.8.4	cotunnite	PbCl_2	highly soluble			King 1992
8.8.18	boleite	$\text{KPb}_{26}\text{Ag}_9\text{Cu}_{24}(\text{OH})_{48}\text{Cl}_{62}$		soluble in nitric acid & ammonia		Schumann 1992
8.10.3	chloromanganokalite	K_4MnCl_6	readily soluble			Antony et al. 1997
8.11.4	douglasite	$\text{K}_2\text{Fe}^{2+}\text{Cl}_4 \cdot 2\text{H}_2\text{O}$	slightly soluble			King 1992
8.11.5	erythrosiderite	$\text{K}_2\text{Fe}^{3+}\text{Cl}_5 \cdot \text{H}_2\text{O}$	highly soluble			King 1992
8.11.6	rinneite	$\text{K}_3\text{NaFe}^{2+}\text{Cl}_6$	highly soluble			King 1992
8.11.7	kremersite	$(\text{NH}_4, \text{K})_2\text{Fe}^{3+}\text{Cl}_5 \cdot \text{H}_2\text{O}$	highly soluble			King 1992
8.11.9	nickelbischofite	$\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$	soluble			King 1992
8.12.4	hieratite	K_2SiF_6	highly soluble			King 1992
9	Borates					
9.1.1	sassolite	H_3BO_3	soluble in warm water			King 1992; Schumann 1992
9.1.6	kernite	$\text{Na}_2[\text{B}_4\text{O}_6(\text{OH})_2] \cdot 3\text{H}_2\text{O}$	moderately soluble			King 1992
9.1.8	tincalconite	$\text{Na}_2(\text{B}_4\text{O}_7) \cdot 5\text{H}_2\text{O}$	highly soluble			King 1992
9.1.9	borax	$\text{Na}_2(\text{B}_4\text{O}_5)(\text{OH})_4 \cdot 8\text{H}_2\text{O}$	highly soluble			King 1992; Schumann 1992
9.2.16	kaliborite	$\text{KMg}_2\text{H}[\text{B}_6\text{O}_8(\text{OH})_5]_2(\text{H}_2\text{O})_4$	slightly soluble			King 1992
9.2.17	iquiqueite	$\text{K}_3\text{Na}_4\text{Mg}(\text{CrO}_4)\text{B}_{24}\text{O}_{39}(\text{OH}) \cdot 12\text{H}_2\text{O}$	somewhat soluble in cold water			Ericksen et al. 1986
9.3.12	inyoite	$\text{Ca}(\text{H}_4\text{B}_3\text{O}_7)(\text{OH}) \cdot 4\text{H}_2\text{O}$	highly soluble in hot water			King 1992
9.3.14	colemanite	$\text{Ca}[\text{B}_3\text{O}_4(\text{OH})_3] \cdot \text{H}_2\text{O}$	slightly soluble			King 1992
9.3.21	ulexite	$\text{NaCa}[\text{B}_5\text{O}_6(\text{OH})_6] \cdot 5\text{H}_2\text{O}$	moderately soluble			King 1992; Schumann 1992
10	Borates with othe anions					
10.1.12	hydrochlorborite	$\text{Ca}_4\text{B}_8\text{O}_{15}\text{Cl}_2 \cdot 21\text{H}_2\text{O}$	soluble		saturated salt solution	Hurlbut et al. 1977
10.3.1	sulfoborite	$\text{Mg}_3[\text{B}(\text{OH})_4]_2(\text{SO}_4)(\text{OH}, \text{F})_2$	decomposes in water			King 1992

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Hey Number	Mineral Name	Chemical Formula	Water	Acids	Alterations	References
11	Carbonates					
11.1.2	natrite	Na_2CO_3	highly soluble			Schumann 1992
11.1.3	nahcolite	NaHCO_3	highly soluble			King 1992
11.1.5	thermonatrite	$\text{Na}_2\text{CO}_3 \cdot \text{H}_2\text{O}$	highly soluble			King 1992
11.1.6	trona	$\text{Na}_3\text{H}(\text{CO}_3)_2 \cdot 2\text{H}_2\text{O}$	highly soluble			King 1992
11.1.7	natron	$\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$	highly soluble			King 1992
11.1.14	gaylussite	$\text{Na}_2\text{Ca}(\text{CO}_3)_2 \cdot 5\text{H}_2\text{O}$	soluble			King 1992; Schumann 1992
11.1.18	teschemacherite	$(\text{NH}_4)\text{HCO}_3$	highly soluble			King 1992
11.1.NI	gregoryite	$(\text{Na}_2, \text{K}_2, \text{Ca})\text{CO}_3$	soluble			Hazen & Ausubel 2016
11.2.1	malachite	$\text{Cu}_2(\text{CO}_3)(\text{OH})_2$	decomposes in hot water			King 1992
11.2.2	azurite	$\text{Cu}_3(\text{CO}_3)_2(\text{OH})_2$	decomposes in hot water			King 1992
				soluble in nitric acid & ammonia	nitric acid: effervescence	Schumann 1992
11.3.1	magnesite	MgCO_3		readily soluble in warm dilute acids		King 1983
			4.2ppm (@ 25C & 1atm)			Morey 1962
11.4.1	calcite	CaCO_3	20ppm (@ 25C & 1atm)			Morey 1962
11.4.2	aragonite	CaCO_3	highly soluble			King 1983
11.4.6	dolomite	$\text{CaMg}(\text{CO}_3)_2$	3-4ppm (@ 25C & 1atm)			Morey 1962
11.6.1	smithsonite	ZnCO_3		easily soluble in warm acids		Schumann 1992
11.6.5	aurichalcite	$(\text{Zn}, \text{Cu})_5(\text{CO}_3)_2(\text{OH})_6$		soluble in acid & ammonia		Schumann 1992
11.8.3	lanthanite-(Ce)	$(\text{Ce}, \text{La}, \text{Nd})_2(\text{CO}_3)_3 \cdot 8\text{H}_2\text{O}$	decomposes in hot water			King 1992
			decomposes in hot water			King 1992
11.9.1	cerussite	PbCO_3		readily soluble in dilute acids		King 1983
				soluble in nitric acid	effervescence	Schumann 1992
11.9.3	dundasite	$\text{PbAl}_2(\text{CO}_3)_2(\text{OH})_4 \cdot \text{H}_2\text{O}$		soluble in acid	disintegration	King 1983
11.11.4	grimselite	$\text{K}_3\text{Na}(\text{UO}_2)(\text{CO}_3)_3 \cdot \text{H}_2\text{O}$	soluble			Hazen & Ausubel 2016
11.11.5	bayleyite	$\text{Mg}_2(\text{UO}_2)(\text{CO}_3)_3 \cdot 18\text{H}_2\text{O}$	soluble			Hazen & Ausubel 2016

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Hey Number	Mineral Name	Chemical Formula	Water	Acids	Alterations	References
11.12.1	rhodochrosite	MnCO ₃		readily soluble in warm, dilute mineral acids		King 1983
12	Carbonates with other anions: with halides or sulfate					
12.2.1	burkeite	Na ₆ (CO ₃)(SO ₄) ₂	highly soluble			King 1992
12.1.23	phosgenite	Pb ₂ CO ₃ Cl ₂		soluble in dilute nitric acid		Schumann 1992
12.2.2	hanksite	Na ₂₂ K(SO ₄) ₉ (CO ₃) ₂ Cl	highly soluble			King 1992
12.2.11	leadhillite	Pb ₄ (CO ₃) ₂ (SO ₄)(OH) ₂		soluble in dilute nitric acid		Schumann 1992
12.2.18	schröckingerite	NaCa ₃ (UO ₂)(CO ₃) ₃ (SO ₄)F · 10H ₂ O	highly soluble			King 1992
13	Nitrates					
13.1	nitratine	NaNO ₃	highly soluble			King 1992; Schumann 1992
13.2	niter	KNO ₃	highly soluble			King 1992
13.4	nitromagnesite	Mg(NO ₃) ₂ · 6H ₂ O	highly soluble			King 1992
13.5	nitrocalcite	Ca(NO ₃) ₂ · 4H ₂ O	highly soluble			King 1992
13.6	nitrobarite	Ba(NO ₃) ₂	highly soluble			King 1992
13.9	darapskite	Na ₃ (SO ₄)(NO ₃) · H ₂ O	highly soluble			King 1992
14	Silicates not containing aluminum					
14.2.1	diopase	CuSiO ₃ · H ₂ O		soluble in nitric acid & ammonia		Schumann 1992
14.4.11	sepiolite	Mg ₄ (Si ₆ O ₁₅)(OH) ₂ · 6H ₂ O		soluble in acid		Schumann 1992
14.5.1	wollastonite	CaSiO ₃		soluble in acid		Schumann 1992
14.5.6	larnite	Ca ₂ SiO ₄	moderately soluble			King 1992
14.9.6	titanite	CaTi(SiO ₄)O		soluble in sulfuric acid		Schumann 1992
14.16.10	uranophane	Ca(UO ₂) ₂ (SiO ₃ OH) ₂ · 5H ₂ O		soluble in acid		Schumann 1992
16	Silicates containing aluminum & other metals					
16.8.6	phlogopite	KMg ₃ (AlSi ₃ O ₁₀)(OH) ₂		soluble in sulfuric acid		Schumann 1992
16.9.2	anorthite	Ca(Al ₂ Si ₂ O ₈)		soluble in acid (inc. var. bytownite & labradorite)		Schumann 1992
16.16.5	carpholite	Mn ²⁺ Al ₂ (Si ₂ O ₆)(OH) ₄		easily soluble in acid		Schumann 1992

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16.23.3	vesuvianite	$\text{Ca}_{19}\text{Fe}^{3+}\text{Al}_4(\text{Al}_6\text{Mg}_2)(\square_4)\square$ $[\text{Si}_2\text{O}_7]_4[\text{SiO}_4]_{10}\text{O}(\text{OH})_9$		partially soluble in acid		Schumann 1992
17	Silicates containing other anions					
17.2.9	zinnwaldite	$\text{KFe}^{2+}\text{Al}(\text{Al}_2\text{Si}_2\text{O}_{10})(\text{OH})_2 - \text{KL}$ $\text{i}_2\text{Al}(\text{Si}_4\text{O}_{10})(\text{F}, \text{OH})_2$		soluble in acid		Schumann 1992
17.3.3	sodalite	$\text{Na}_4(\text{Si}_3\text{Al}_3)\text{O}_{12}\text{Cl}$		soluble in nitric acid		Schumann 1992
18	Niobates & Tantalates					
18.3.3	betafite	$(\text{Ca}, \text{Na}, \text{U})_2(\text{Ti}, \text{Nb}, \text{Ta})_2\text{O}_6\text{Z}(\text{OH})$		soluble in acid		Schumann 1992
18.4.5	euxenite	$(\text{Y}, \text{Ca}, \text{Ce}, \text{U}, \text{Th})(\text{Nb}, \text{Ta}, \text{Ti})_2\text{O}_6$		powder soluble in sulfuric acid		Schumann 1992
19	Phospahtes					
19.2.8	turquoise	$\text{CuAl}_6(\text{PO}_4)_4(\text{OH})_8 \cdot 4\text{H}_2\text{O}$		soluble in acids		Schumann 1992
19.8.2	brazilianite	$\text{NaAl}_3(\text{PO}_4)_2(\text{OH})_4$		soluble in acids		Schumann 1992
19.9.3-5	monazite	$\text{REE}(\text{PO}_4)$		easily soluble in sulfuric acid		Schumann 1992
19.11.10	torbernite	$\text{Cu}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 12\text{H}_2\text{O}$		soluble in nitric acid		Schumann 1992
19.11.14	meta-autunite	$\text{Ca}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 6-8\text{H}_2\text{O}$	highly soluble			King 1992
19.11.15	autunite	$\text{Ca}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 11\text{H}_2\text{O}$	highly soluble			King 1992
19.12.35	ludlamite	$\text{Fe}^{2+}_3(\text{PO}_4)_2 \cdot 4\text{H}_2\text{O}$		soluble in acids		Schumann 1992
19.13.11	vivianite	$\text{Fe}^{2+}_3(\text{PO}_4)_2 \cdot 8\text{H}_2\text{O}$		readily soluble in weak acids		King 1983
20	Arsenates					
20.1.2	olivenite	$\text{Cu}_2(\text{AsO}_4)(\text{OH})$		soluble in acid & ammonia		Schumann 1992
20.1.8	euchroite	$\text{Cu}_2(\text{AsO}_4)(\text{OH}) \cdot 3\text{H}_2\text{O}$		soluble in nitric acid		Schumann 1992
20.1.13	liroconite	$\text{Cu}_2\text{Al}(\text{AsO}_4)(\text{OH})_4 \cdot 4\text{H}_2\text{O}$		soluble in acids		Schumann 1992
20.7.10	uranospinite	$\text{Ca}(\text{UO}_2)(\text{AsO}_4)_2 \cdot 10\text{H}_2\text{O}$		easily soluble in acids		Fron del 1958
20.9.2	scorodite	$\text{Fe}^{3+}\text{AsO}_4 \cdot 2\text{H}_2\text{O}$	soluble		to iron hydroxides & aqueous arsenate	Dove & Rimstidt 1985
20.10.9	annabergite	$\text{Ni}_3(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O}$		soluble in acids		Schumann 1992

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Hey Number	Mineral Name	Chemical Formula	Water	Acids	Alterations	References
21	Vanadates					
21.2.1	metarossite	$\text{Ca}(\text{V}_2\text{O}_6) \cdot 2\text{H}_2\text{O}$	moderately soluble			King 1992
21.2.2	rossite	$\text{Ca}(\text{VO}_3)_2 \cdot 4\text{H}_2\text{O}$	slightly soluble			King 1992
21.2.3	metahewettite	$\text{CaV}_6\text{O}_{16} \cdot 3\text{H}_2\text{O}$	slightly soluble			Hillebrand et al. 1914; King 1992
21.2.4	hewettite	$\text{CaV}_6\text{O}_{16} \cdot 9\text{H}_2\text{O}$	slightly soluble			Hillebrand et al. 1914; King 1992
21.2.5	pintadoite	$\text{Ca}_2(\text{V}_2\text{O}_7) \cdot 9\text{H}_2\text{O}$	moderately soluble			King 1992
21.2.7	fernandinite	$(\text{Ca},\text{K})(\text{V}^{5+},\text{V}^{4+},\text{Fe}^{2+})_8\text{O}_{20} \cdot 10\text{H}_2\text{O}$	moderately soluble			King 1992
21.2.8	pascoite	$\text{Ca}_3(\text{V}_{10}\text{O}_{28}) \cdot 17\text{H}_2\text{O}$	easily soluble			Hillebrand et al. 1914
21.3.10	descloizite	$\text{PbZn}(\text{VO}_4)(\text{OH})$		soluble in acids		Schumann 1992
22	Phosphates, Arsenates, & Vanadates with other anions					
22.1.1	amblygonite	$\text{LiAl}(\text{PO}_4)\text{F}$		soluble in sulfuric acid		Schumann 1992
22.2.9	pyromorphite	$\text{Pb}_5(\text{PO}_4)_3\text{Cl}$		soluble in nitric acid		Schumann 1992
22.2.10	mimetite	$\text{Pb}_5(\text{AsO}_4)_3\text{Cl}$		soluble in nitric acid		Schumann 1992
22.2.14	vanadinite	$\text{Pb}_5(\text{VO}_4)_3\text{Cl}$		readily soluble in nitric acid		Schumann 1992
22.2.NI	apatite	$\text{Ca}_5(\text{PO}_4)_3(\text{Cl}/\text{F}/\text{OH})$		easily soluble in acid		Schumann 1992
22.3.6	chalcophyllite	$\text{Cu}_{18}\text{Al}_2(\text{AsO}_4)_4(\text{SO}_4)_3(\text{OH})_{24} \cdot 36\text{H}_2\text{O}$		easily soluble in acid & ammonia		Schumann 1992
22.4.5	tyrolite	$\text{Ca}_2\text{Cu}_9(\text{AsO}_4)_4(\text{CO}_3)(\text{OH})_8 \cdot 11\text{H}_2\text{O}$		easily soluble in acid & ammonia		Schumann 1992
25	Sulfates					
25.1.1	metathénardite	Na_2SO_4	soluble			King 1992; Schumann 1992
25.1.3	mirabilite	$\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$	highly soluble			King 1992
25.1.4	arcanite	K_2SO_4	highly soluble			
25.1.6	misenite	$\text{K}_8\text{H}_6(\text{SO}_4)_7$	highly soluble			King 1992
25.1.7	aphthitalite	$(\text{K},\text{Na})_3\text{Na}(\text{SO}_4)_2$	highly soluble			King 1992
25.1.8	mascagnite	$(\text{NH}_4)_2\text{SO}_4$	highly soluble			King 1992

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25.1.NI	lecontite	$(\text{NH}_4, \text{K})\text{NaSO}_4 \cdot 2\text{H}_2\text{O}$	highly soluble			King 1992; Hazen & Ausubel 2016
25.2.1	chalcocyanite	CuSO_4	highly soluble			King 1992
25.2.3	chalcantite	$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$	highly soluble			King 1992; Schumann 1992
25.2.5	dolerophanite	$\text{Cu}_2(\text{SO}_4)\text{O}$	highly soluble			King 1992
25.2.7	brochantite	$\text{Cu}_4(\text{SO}_4)(\text{OH})_6$		soluble in dilute acid		Schumann 1992
25.2.11	kröhnkite	$\text{Na}_2\text{Cu}(\text{SO}_4)_2 \cdot 2\text{H}_2\text{O}$	highly soluble			King 1992
25.2.12	natrochalcite	$\text{NaCu}_2(\text{SO}_4)_2(\text{OH}) \cdot 2\text{H}_2\text{O}$	moderately soluble			King 1992
25.2.14	fedotovite	$\text{K}_2\text{Cu}_3(\text{SO}_4)_3\text{O}$	soluble			Hazen & Ausubel 2016
25.2.15	cyanochroite	$\text{K}_2\text{Cu}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$	highly soluble			King 1992
25.2.20	cyanotrichite	$\text{Cu}_4\text{Al}_2(\text{SO}_4)(\text{OH})_{12} \cdot 2\text{H}_2\text{O}$		soluble in acid		Schumann 1992
25.2.24	ransomite	$\text{CuFe}_2(\text{SO}_4)_4 \cdot 6\text{H}_2\text{O}$	soluble			Hazen & Ausubel 2016
25.3.1	kieserite	$\text{MgSO}_4 \cdot \text{H}_2\text{O}$	moderately soluble			King 1992
25.3.5	hexahydrite	$\text{MgSO}_4 \cdot 6\text{H}_2\text{O}$	highly soluble			King 1992
25.3.6	epsomite	$\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$	highly soluble			King 1992; Schumann 1992
25.3.8	vanthoffite	$\text{Na}_6\text{Mg}(\text{SO}_4)_4$	moderately soluble			King 1992
25.3.9	blödite	$\text{Na}_2\text{Mg}(\text{SO}_4)_2 \cdot 4\text{H}_2\text{O}$	highly soluble			King 1992
25.3.11	löweite	$\text{Na}_{12}\text{Mg}_7(\text{SO}_4)_{13} \cdot 15\text{H}_2\text{O}$	highly soluble			King 1992
25.3.13	langbeinite	$\text{K}_2\text{Mg}_2(\text{SO}_4)_3$	slightly soluble			King 1992
25.3.14	leonite	$\text{K}_2\text{Mg}(\text{SO}_4)_2 \cdot 4\text{H}_2\text{O}$	highly soluble			King 1992
25.3.15	picromerite	$\text{K}_2\text{Mg}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$	highly soluble			King 1992
25.3.17	boussingaultite	$(\text{NH}_4)_2\text{Mg}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$	highly soluble			King 1992
25.3.18	pickeringite	$\text{MgAl}_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	highly soluble			King 1992
25.4.3	gypsum	$\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$	slightly soluble			King 1983; King 1992
25.4.4	glauuberite	$\text{Na}_2\text{Ca}(\text{SO}_4)_2$	moderately soluble			King 1992; Schumann 1992
25.4.8	syngenite	$\text{K}_2\text{Ca}(\text{SO}_4)_2 \cdot \text{H}_2\text{O}$	slightly soluble		separation of gypsum	King 1992
25.4.10	polyhalite	$\text{K}_2\text{Ca}_2\text{Mg}(\text{SO}_4)_4 \cdot 2\text{H}_2\text{O}$	decomposes in water		separation of gypsum	King 1992
25.4.11	koktaite	$(\text{NH}_4)_2\text{Ca}(\text{SO}_4)_2 \cdot \text{H}_2\text{O}$	soluble			Hazen & Ausubel 2016
25.4.12	ye'elimite	$\text{Ca}_4\text{Al}_6(\text{SO}_4)\text{O}_{12}$	soluble			Hazen & Ausubel 2016
25.4.13	ettringite	$\text{Ca}_6\text{Al}_2(\text{SO}_4)_3(\text{OH})_{12} \cdot 26\text{H}_2\text{O}$	moderately soluble			King 1992

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25.4.17	baryte	BaSO ₄		slowly soluble in concentrated sulfuric acid		Schumann 1992
25.5.2	goslarite	ZnSO ₄ · 7H ₂ O	highly soluble readily soluble in cold water			King 1992 King 1983
25.5.7	serpierite	Ca(Cu,Zn) ₄ (SO ₄) ₂ (OH) ₆ · 3H ₂ O		souble in acid		Schumann 1992
25.5.13	bianchite	(Zn,Fe)SO ₄ · 6H ₂ O	highly soluble			King 1992
25.5.15	zincmelanterite	(Zn,Cu,Fe)SO ₄ · 7H ₂ O	highly soluble			Waller 1992
25.6.3	mendozite	NaAl(SO ₄) ₂ · 11H ₂ O	highly soluble			King 1992
25.6.5	aluminite	Al ₂ (SO ₄)(OH) ₄ · 7H ₂ O		souble in acid		Schumann 1992
25.6.7	alunogen	Al ₂ (SO ₄) ₃ · 17H ₂ O	highly soluble			King 1992
25.6.8	alunite	KAl ₃ (SO ₄) ₂ (OH) ₆		soluble in hot sulfuric acid		Schumann 1992
25.6.14	alum-(Na)	NaAl(SO ₄) ₂ · 12H ₂ O	highly soluble			Waller 1992
25.6.16	kalinite	KAl(SO ₄) ₂ · 11H ₂ O	highly soluble			King 1992
25.6.17	alum-(K)	KAl(SO ₄) ₂ · 12H ₂ O	highly soluble			King 1992
25.6.20	tschermigite	(NH ₄)Al(SO ₄) ₂ · 12H ₂ O	highly soluble			King 1992
25.7.3	anglesite	PbSO ₄		soluble in warm sulfuric acid		Schumann 1992
25.7.6	linarite	PbCu(SO ₄)(OH) ₂		soluble in dilute nitric acid		Schumann 1992
25.8.5	minasragrite	(V ⁴⁺ O)(SO ₄) · 5H ₂ O	highly soluble			King 1992; Hazen & Ausubel 2016
25.8.7	redingtonite	(Fe ²⁺ ,Mg,Ni)(Cr,Al) ₂ (SO ₄) ₄ · 22H ₂ O	soluble			Hazen & Ausubel 2016
25.9.3	mallardite	MnSO ₄ · 7H ₂ O	highly soluble			King 1992
25.9.7	apjohnite	Mn ²⁺ Al ₂ (SO ₄) ₄ · 22H ₂ O	highly soluble			King 1992
25.9.9	ilesite	(Mn,Zn,Fe)SO ₄ · 4H ₂ O	highly soluble			King 1992
25.9.13	dietrichite	(Zn,Fe ²⁺ ,Mn ²⁺)Al ₂ (SO ₄) ₄ · 22H ₂ O	highly soluble			King 1992
25.10.1	szomolnokite	FeSO ₄ · H ₂ O	moderately soluble		oxidises to Fe3+	King 1992
25.10.3	siderotil	FeSO ₄ · 5H ₂ O	highly soluble			King 1992
25.10.5	melanterite	FeSO ₄ · 7H ₂ O	highly soluble			King 1992
25.10.6	rhomboclase	(H ₅ O ₂)Fe ³⁺ (SO ₄) ₂ · 2H ₂ O	highly soluble			Tosca et al. 2007

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25.10.8	kornelite	$\text{Fe}_2(\text{SO}_4)_3 \cdot 7\text{H}_2\text{O}$	highly soluble			King 1992
25.10.9	coquimbite	$\text{AlFe}^{3+}_3(\text{SO}_4)_6(\text{H}_2\text{O})_{12} \cdot 6\text{H}_2\text{O}$	highly soluble			King 1992
25.10.17	amarantite	$\text{Fe}^{3+}_2(\text{SO}_4)_2\text{O} \cdot 7\text{H}_2\text{O}$	decomposes in water			King 1992
25.10.20	römerite	$\text{Fe}^{2+}\text{Fe}^{3+}_2(\text{SO}_4)_4 \cdot 14\text{H}_2\text{O}$	highly soluble			King 1992
25.10.21	břilinite	$\text{Fe}^{2+}\text{Fe}^{3+}_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	highly soluble			King 1992
25.10.22	copiapite	$\text{Fe}^{2+}\text{Fe}^{3+}_4(\text{SO}_4)_6(\text{OH})_2 \cdot 20\text{H}_2\text{O}$	highly soluble			King 1992
25.11.2	sideronatrite	$\text{Na}_2\text{Fe}(\text{SO}_4)_2(\text{OH}) \cdot 3\text{H}_2\text{O}$	decomposes in hot water			King 1992
25.11.4	ferrinatrite	$\text{Na}_3\text{Fe}(\text{SO}_4)_3 \cdot 3\text{H}_2\text{O}$	slightly soluble			King 1992
25.11.7	krausite	$\text{KFe}(\text{SO}_4)_2 \cdot \text{H}_2\text{O}$	decomposes in water			King 1992
25.11.9	jarosite	$\text{KFe}^{3+}_3(\text{SO}_4)_2(\text{OH})_6$		soluble in acid		Schumann 1992
25.11.10	voltaite	$\text{K}_2\text{Fe}^{2+}_5\text{Fe}^{3+}_3\text{Al}(\text{SO}_4)_{12} \cdot 18\text{H}_2\text{O}$	decomposes in water		yellow precipitate	King 1992
25.11.12	metavoltine	$\text{Na}_6\text{K}_2\text{Fe}^{2+}\text{Fe}^{3+}_6(\text{SO}_4)_{12}\text{O}_2 \cdot 18\text{H}_2\text{O}$	slightly soluble			King 1992
25.11.18	godovikovite	$(\text{NH}_4)\text{Al}(\text{SO}_4)_2$	slowly soluble			Jambor & Grew 1990; King 1992; Antony et al. 2003
25.11.19	botryogen	$\text{MgFe}^{3+}(\text{SO}_4)_2(\text{OH}) \cdot 7\text{H}_2\text{O}$	slightly soluble in hot water			King 1992
25.11.26	halotrichite	$\text{Fe}^{2+}\text{Al}_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	highly soluble			King 1992; Schuman 1992
25.12.6	morenosite	$\text{NiSO}_4 \cdot 7\text{H}_2\text{O}$	highly soluble			King 1992
25.12.NI	wupatkiite	$(\text{Co,Mg,Ni})\text{Al}_2(\text{SO}_4)_4 \cdot 22\text{H}_2\text{O}$	soluble			Hazen & Ausubel 2016
26	Sulfates with Halide					
26.2	sulphohalite	$\text{Na}_6(\text{SO}_4)_2\text{FCl}$	slightly soluble			King 1992
26.3	schairerite	$\text{Na}_{21}(\text{SO}_4)_7\text{ClF}_6$	moderately soluble			King 1992
26.9	aubertite	$\text{CuAl}(\text{SO}_4)_2\text{Cl} \cdot 14\text{H}_2\text{O}$	soluble			Hazen & Ausubel 2016
26.14	kainite	$\text{KMg}(\text{SO}_4)\text{Cl} \cdot 3\text{H}_2\text{O}$	decomposes in water		to epsomite & sylvite	King 1992
26.23	pseudograndreefite	$\text{Pb}_6(\text{SO}_4)\text{F}_{10}$	soluble			Hazen & Ausubel 2016
26.25	caracolite	$\text{Na}_3\text{Pb}_2(\text{SO}_4)_3\text{Cl}$	soluble			Hazen & Ausubel 2016

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27	Sulfites, Chromates, Molybdates, & Tungstates					
27.2.5	crocoite	PbCrO ₄		disintegrates in acids		King 1983
27.3.3	wulfenite	Pb(MoO ₄)		decomposes in acids		King 1983
27.4.15	wolframite	(Fe ²⁺)WO ₄ – (Mn ²⁺)WO ₄		powder turns blue in concentrated sulfuric acid		Schumann 1992
29	Iodates					
29.3	lautarite	Ca(IO ₃) ₂	slightly soluble			King 1992
29.5	dietzeite	Ca ₂ (IO ₃) ₂ (CrO ₄) · H ₂ O	moderately soluble			King 1992
31	Oxalates, Citrates, Mellitates, & Acetates					
31.1.1	oxammitte	(NH ₄) ₂ (C ₂ O ₄) · H ₂ O	highly soluble			King 1992
31.1.3	wheatleyite	Na ₂ Cu(C ₂ O ₄) ₂ · 2H ₂ O	soluble			Hazen & Ausubel 2016
31.1.5	whewellite	Ca(C ₂ O ₄) · H ₂ O		easily soluble in acids		Schumann 1992
31.1.NI	natroxalate	Na ₂ (C ₂ O ₄)	soluble			Hazen & Ausubel 2016
31.3.1	mellite	Al ₂ [C ₆ (COO) ₆] · 16H ₂ O		soluble in nitric acid & potassium hydroxide		Schumann 1992
O	'Precious' Non-minerals					
O.4	coral	CaCO ₃		soluble in acid	effervescence	Schumann 1992
			soluble in hot water			Schumann 1974

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